

ECAT Chemistry Chapter 9 Solutions Online Test

Sr	Questions	Answers Choice
1	When the solute is present in trace quantities the following expression is used	A. Gram per million B. Milligram percent C. Microgram percent D. Parts per million
2	Precipitation will occur until the ionic product becomes	A. Equal to K _{sp} B. Lesser than K _{sp} C. Greater than K _{sp} D. None of these
3	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
4	Number of moles of the solute dissolved per dm ³ of the solution is knows as	A. Molarity B. Formality C. %age D. None of these
5	The boiling point of an a zeotropic mixture of water and ethye alcohol is less than that of water and alcohol. The mixture shows	A. That solution is highly saturated B. No deviation from Raoult's law C. Positive deviation from Raoult's law D. Negative deviation from Raoult's law
6	A solution has 92 g of ethyl alcohol, 96 g of methyl alcohol and 90 g of water. Mole percentage of ethyl alcohol in the solution is	A. 10 B. 20 C. 25 D. 50
7	The freezing point of 1 molal NaCl solution assuming NaCl to be 100% dissociated in water in	A1.86 °C B3.72 °C C. +1.86 °C D. +3.72 °C
8	Solution may have units	A. Molarity B. Molality C. Mole fraction D. All of them
9	Solution which distill without change in composition or temperature are called	A. Amorphous B. Azeotropic mixture C. Ideal D. Super saturated
10	The osmotic pressure of a dilute solution is directly proportional to the	A. Diffusion rate of the solute B. lonic concentration C. Elevation in boiling point D. Flow of solvent from a concentrated to a dilute solution
11	An azeotropic mixture showing it's positive deviation from Raoult's law, the volume of the mixture is :	A. Slightly more than the total volum of the components. B. Slightly less than the total volume of the components. C. Equal to the total volume of the components. D. None of these.
		A. 0.1 B. 0.01

13	The amount of solute present in the given amount of solvent is called	A. Molarity B. Molality C. Concentration D. Solubility
14	A solutiion of 0.5 mole camphor in 100 grmas chloroform (K_b = 0.322) has rise in boiling point than that of chloroform by	A. 0.81°C B. 1.61°C C. 1.81°C D. 0.61°C
15	0.5 M of H ₂ SO ₄ is diluted from 1 litre to 10 litre, normality of resulting solution is	A. 1 N B. 0.1 N C. 10 N D. 11 N
16	What will be the molarity of solution if 103 g (NH ₄) ₂ SO ₄ is dissolved per 600 cm ³ of water	A. 2.32 M B. 3.32 M C. 4.32 M D. 1.30 M
17	Partial pressure of a solution component is directly proportional to its mole fraction. This statement is known as	A. Henry's law B. Raoult's law C. Distribution law D. Ostwald's dilution law
18	Equal volumes of 0.1 M AgNO $_3$ and 0.2 M NaCl are mixed. The concentration of NO $_3$ ions in the mixture will be	A. 0.1 M B. 0.05 M C. 0.2 M D. 0.15 M
19	The vant Hoff factor (1) accounts for	A. Degree of solubilisation of solute B. The extent of dissolution of solute C. The extent of dissolution of solute D. The degree of decomposition of solution
20	The solutions of NaCl and KCl are prepared separately by dissolving same amount of solute in water, which of the following statements is true fro these solutions?	A. KCl solution will have higher boiling point than NaCl solution. B. Both the solutions have same boiling points. C. KCl and NaCl solutions possess same vapour pressure. D. KCl solution possesses lower freezing point than NaCl solution.