

ECAT Chemistry Chapter 9 Solutions Online Test

Sr	Questions	Answers Choice
1	A solution of glucose is 10% The volume to which 1g mole of it dissolved will be :	A. 8.1 dm ³ class="MsoNormal"> <o:p></o:p> class="MsoNormal"> <o:p></o:p> B. 1.7 dm ³ class="MsoNormal"> <o:p></o:p> C. 1.8 dm ³ class="MsoNormal"> <o:p></o:p> C. 1.8 dm ³ class="MsoNormal"> <o:p></o:p> D. 6.1 dm ³ cp class="MsoNormal"> <o:p></o:p> D. 6.1 dm ³ cp class="MsoNormal"> <o:p></o:p>
2	Which one of the following is used as antifreeze in the radiator	A. Methanol B. Ethanol C. Ethylene glycol D. Glycerin
3	Solubility cure of Na ₂ SO ₄ 10H ₂ shows	A. Constant increase of solubility B. Constant decreases of solubility C. Discontinuously solubility wit temperature D. None of above
4	Solution which distill without change in composition or temperature are called	A. Amorphous B. Azeotropic mixture C. Ideal D. Super saturated
5	A solution containing maximum amount of solute dissolved at a given temperature is called	A. Saturated solution B. Unsaturated solution C. Supersaturated solution D. Impure solution
6	Which of the following liquid pairs shown a positive deviation from Raoult's law	A. CH ₃ COOH ₃ CH ₃ CI B. C ₆ H ₆ + CH ₄ OH C. H ₂ O + HCI D. H ₂ O + HNO ₃ O
7	Units of molarity are	A. gm/lit B. mol/lit C. kg/lit D. None of these
8	If 18 g glucose (C ₆ H ₁₂ C ₆) is present in 1000 g of an aqueous of glucose it is said to be	A. 1 molal B. 1.1 molal C. 0.5 molal D. 0.1 molal
9	The movement of solvent molecules through a semipermeable membrane is called	A. Electrolysis B. Electrophoresis C. Osmosis D. Cataphoresis
10	Maximum freezing point falls in	A. Camphor B. Naphthalene C. Benzene D. Water
11	What is the molarity of H ₂ SO ₄ solution that has density of 1.84 gm/cc at 35°C and contains 98% by weight?	A. 4.18 M B. 8.14 M C. 18.4 M D. 18 M
12	The process in which solvent particles surround solute particles is called	A. Hydrolysis B. Hydration C. Solvation D. Dissolution
13	Which statement is incorrect for and ideal solution	A. The forces of attractions between solute and solvent molecules are same B. There is no evolution or absorption of heat C. Volume of the solution is less than

		sum of volumes of individual components D. Vapour pressure of solution is directly proportional to the mole fraction of solvent
14	The weight of pure NaOH required to prepare 250 cm ³ of 0.1 N solution is	A. 4 g B. 1 g C. 2 g D. 5 g
15	The molal depression constant depends upon	A. Nature of solute B. Nature of solvent C. \Delta H<\sub>\solution<!--\sub--> D. Vapour pressure of solution
16	(A) is one molar NaCl solution and (b) is 1 molal NaCl solutin :	A. A and B are of same strength. B. A is more Concentrate than B. C. b is more Concentrate than A. D. None of above.
17	Which of the following mixture of liquids show negative deviation form Raults law	A. Ethyl alcohol and ether B. HCl and water C. Phenol- water D. Chlorobenene-bromobenene
18	The boiling point of an a zeotropic mixture of water and ethye alcohol is less than that of water and alcohol. The mixture shows	A. That solution is highly saturated B. No deviation from Raoult's law C. Positive deviation from Raoult's law D. Negative deviation from Raoult's law
19	0.5 M of H ₂ SO ₄ is diluted from 1 litre to 10 litre, normality of resulting solution is	A. 1 N B. 0.1 N C. 10 N D. 11 N
20	Which is independent of temperature	A. Molarity B. Molality C. Normality D. Mole fraction