

ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	London dispersion forces are also called	<p>A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces</p>
2	Which of the following is a pseudo solid?	<p>A. CaF_2 B. Glass C. NaCl D. All</p>
3	In a group on going downward, polarizability generally:	<p>A. Decreases B. Increases C. Remain constant D. Negligible</p>
4	At sea level and at 100°C the vapor pressure of water in an open system is:	<p>A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg</p>
5	Evaporation of water is possible at:	<p>A. Above 100°C B. 0°C C. 100°C D. At all temperature</p>
6	Amorphous solids:	<p>A. Have a sharp melting point B. Undergo clean cleavage when cut with knife C. Have a perfect arrangement of atoms D. Can possess small regions of orderly arrangement of atoms</p>
7	The molecules of CO_2 in dry ice form the	<p>A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystal</p>
8	Hydrogen bonding is present in one of the following pairs:	<p>A. NH_3 B. H_2 C. HF D. All of above</p>

A. Enhanced electronegative character of Nitrogen
 B. Pyramidal structure of

9	The boiling point of NH_3 is maximum among the hydrides of group V elements due to	<p>B. Trigonal structure of NH_3</p> <p>C. Very small size of Nitrogen</p> <p>D. Enhanced electropositive character of Nitrogen</p>
10	Hydrocarbon molecule with large chain lengths experience:	<p>A. Weaker attractive forces</p> <p>B. Stronger attractive forces</p> <p>C. Repulsive forces</p> <p>D. No attractive forces</p>
11	HF has exceptionally low acidic strength due to:	<p>A. Smaller size of fluorine.</p> <p>B. Stronger polar bond between H and F.</p> <p>C. Electronegativity of fluorine</p> <p>D. Strong hydrogen bonding</p>
12	A liquid on evaporation causes	<p>A. Heating effect</p> <p>B. Cooling effect</p> <p>C. Suffocation</p> <p>D. All of above</p>
13	Debye forces are present in one of the following pairs:	<p>A. Na^+ ion and water</p> <p>B. Argon and water</p> <p>C. Argon and Na^+ ion</p> <p>D. Ne and water</p>
14	The strongest forces are	<p>A. Debye forces</p> <p>B. London dispersion forces</p> <p>C. Dipole-dipole attraction</p> <p>D. Hydrogen</p>
15	Ionic Solids are characterized by	<p>A. Low melting points</p> <p>B. Good conductivity in solid state</p> <p>C. High vapour pressure</p> <p>D. Solubility in polar solvents</p>
16	London dispersion forces are also called:	<p>A. Hydrogen bonding.</p> <p>B. Debye forces</p> <p>C. Van der Waal's forces</p> <p>D. Instantaneous dipole-induced dipole forces.</p>
17	Diamond is a bad conductor because	<p>A. It has a tight structure</p> <p>B. It has a high density</p> <p>C. there is no free electron present in the crystal of diamond of conduct electricity</p> <p>D. Is transparent to light</p>
18	Rate of evaporation and rate of condensation at equilibrium	<p>A. Become very low</p> <p>B. Become very high</p> <p>C. Become equal</p> <p>D. Can never be equal</p>
19	Which of the following liquids has low vapour pressure at 25°C :	<p>A. Diethyl ether</p> <p>B. Acetone.</p> <p>C. Water.</p> <p>D. Ethyl alcohol.</p>
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