

ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	The density of water decreases, When it is freezed at 0°C	A. Change of bond length B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of ice
2	The molecules of CO ₂ in dry ice from the	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystal
3	London dispersion forces are also called	A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces
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5	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
6	The molecules of CO ₂ in dryice from the :	A. Ionic crystals B. Coverlet crystals C. Molecular crystals D. Any type of crystals
7	At sea level and at 100°C the vapor pressure of water in an open system is:	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
8	In a group on going downward, polarizability generally	A. Decreases B. Increases C. Remains constant D. Negligible
9	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
10	Which of the following is a pseudo solid?	A. CaF_2 B. Glass C. NaCl D. All
11	Diamond is a bad conductor because:	A. It has tight structure B. It has a high density C. There is no free electron present in the crystal of diamond of conduct electricity is transparent to light
12	The gases can be converted into liquids by:	A. Increasing the pressure only. B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical points D. Lowering temperature only
13	Rate of evaporation and rate of condensation at equilibrium:	A. Become very low B. Become very high C. Become equal D. Can never be equal.

14	Chloroform and acetone are soluble in each other due to:	A. Instantaneous dipole interactions. B. Dipole-dipole interactions. C. Inter molecular hydrogen bonding. D. All of above
15	HF has exceptionally low acidic strengths due to	A. Smaller size of fluorine B. Strong polar bond between H and F C. Electronegativity of fluorine D. Strong hydrogen bonding
16	Rate of evaporation and rate of condensation at equilibrium	A. Become very low B. Become very high C. Become equal D. Can never be equal
17	Hydrocarbon molecule with large chain lengths experience:	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
18	A liquid on evaporation causes	A. Heating effect B. Cooling effect C. Suffication D. All of above
19	The inter-molecular forces in liquids are:	A. Negligible B. Very weak C. Very strong D. Reasonably strong
20	Which one of the following is weakest inter molecular force?	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatic forces b/w ions D. Dipole dipole forces