

ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	Hydrocarbon molecules with large chain lengths experience	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
2	Heat of vapourization for liquids with strong dipole-dipole forces will have	A. Negligible Values B. Reasonably high values C. Very high values D. very low values
3	Escape of high energy molecules from the surface of liquid is called:	A. Sublimation. B. Distillation. C. Condensation. D. Evaporation.
4	Which of the following is a pseudo solid?	A. CaF_2 B. Glass C. NaCl D. All
5	Diamond is a bad conductor because:	A. It has tight structure B. It has a high density C. There is no free electron present in the crystal of diamond of conduct electricity is transparent to light
6	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
7	Rate of evaporation and rate of condensation at equilibrium	A. Become very low B. Become very high C. Become equal D. Can never be equal
8	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
9	Diamond is a bad conductor because	A. It has a tight structure B. It has a high density C. there is no free electron present in the crystal of diamond of conduct electricity D. Is transparent to light
10	Hydrogen bonding is present between the molecules of	A. NH_3 B. H_2O C. HF D. All of above
11	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of solid I_2 D. Molecular of hydrochloric acid gas
12	The gases can be converted into liquids by:	A. Increasing the pressure only. B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical points D. Lowering temperature only
13	At sea level and at 100°C the vapor pressure of water in an open system is:	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg

14	H ₂ S is a gas while H ₂ O is a liquid at room temperature. It is due to:	A. Less inter-molecular forces in water. B. Covalent bonding H-O in water molecule. C. Hydrogen bonding in water molecules D. Ionic character in water molecules
15	Ionic Solids are characterized by	A. Low melting points B. Good conductivity in solid state C. High vapour pressure D. Solubility in polar solvents
16	The gases can be converted into liquids by	A. increasing the pressure only B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical point D. Lowering temperature only
17	London dispersion forces are also called:	A. Hydrogen bonding. B. Debye forces C. Van der Waal's forces D. Instantaneous dipole-induced dipole forces.
18	The molecules of CO ₂ in dry ice form the :	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystals
19	In a group on going downward, polarizability generally	A. Decreases B. Increases C. Remains constant D. Negligible
20	The molecules of CO ₂ in dry ice form the	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystal