

ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	London dispersion forces are also called:	A. Hydrogen bonding. B. Debye forces C. Van der Waal's forces D. Instantaneous dipole-induced dipole forces.
2	London dispersion forces are also called	A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces
3	Ionic Solids are characterized by	A. Low melting points B. Good conductivity in solid state C. High vapour pressure D. Solubility in polar solvents
4]Which of the following is a pseudo solid?	A. Atoms of He is gaseous stat at high temperate. B. Molecules of water in liquid state. C. Molecules of solid $I_{2(s)}$ D. Molecular of hydrochloric acid gas
5	Debye forces are present in one of the following pairs	A. Na^{+} ion and water B. Argon and water C. Argon and Na^{+} ion D. Ne and Water
6	Force of attraction between atoms of He is	A. London dispersion forces B. Hydrogen bondign C. Coordinate covalent bond D. Covalent bond
7	Diamond is a bad conductor because:	A. It has tight structure B. It has a high density C. There is no free electron present in the crystal of diamond of conduct electricity is transparent to light
8	The strongest forces are:	A. Debye froces B. London dispersion C. Dipole-dipole attraction D. Hydrogen bonding
9	The density of water decreases, when it is freezed at $0^{\circ}C$ because of	A. Change of bond lengths B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of Ice
10	Debye forces are present in on of the following pairs:	A. Na^{+} ion and water B. Argon and water C. Argon and Na^{+} ion D. Ne and water
11	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
12	Polypeptide chains are coiled about one another into a spiral by	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
13	At sea level and at $100^{\circ}C$ the vapour pressure of water in an open system is	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
14	The gases can be converted into liquids by:	A. Increasing the pressure only. B. Lowering temperature and increasing pressure C. Increasing pressure and bringing

temperature below critical points
D. Lowering temperature only

15	Hydrogen bonding is present in one of the following pairs:	A. NH_3 B. H_2O C. HF D. All of above
16	In a group on going downward, polarizability generally	A. Decreases B. Increases C. Remains constant D. Negligible
17	Heat of vapourization for liquids with strong dipole-dipole forces will have	A. Negligible Values B. Reasonably high values C. Very high values D. very low values
18	Trend of boiling point of halogens from fluorine to iodine is that it	A. Decreases B. Is negligible C. Increases D. Remains constant
19	Which one of the following is the weakest intermolecular force	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatics forces between ions D. Dipole-dipole forces
20	Which one of the following is weakest inter molecular force?	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatic forces b/w ions D. Dipole dipole forces