ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	Debye forces are present in one of the following pairs	A. Na⁺ion and water B. Argon and water C. Argon and Na⁺ion D. Ne and Water
2	Amorphous solids	A. Have sharp melting point B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
3	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of solid I₂ D. Molecular of hydrochloric acid gas
4	Hydrogen bonding is present in one of the following pairs:	A. NH₃ B. H₂O C. HF D. All of above
5	Polypeptide chains are coiled about one another into a spiral by:	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
6	Amorphous solids:	A. Have a sharp melting point B. Undergo clean cleavage when cut with knife C. Have a perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
7	The density of water decreases, when it is freezed at $0^{o}C$ because of	A. Change of bond lengths B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of Ice
8	The molecules of CO_2 in dry ice from the :	A. Ionic crystals B. Coverlet crystals C. Molecular crystals D. Any type of crystals
9	Hydrocarbon molecules with large chain lengths experience	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
10	London dispersion forces are also called	A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces
11	The gases can be converted into liquids by	A. increasing the pressure only B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical point D. Lowering temperature only
12	The molecules of CO_2in dry ice from the	A. Ionic crystalls B. Covalent crystals C. Molecular crystals D. Any type of crystal
13	H_2S is a gas while H_2O is a liquid at room temperature. It is due to:	A. Less inter-molecular forces in water. B. Covalent bonding H-O in water molecule.

		C. Hydrogen bonding in water molecules D. Ionic character in water molecules
14	Force of attraction between atoms of He is	A. London dispersion forces B. Hydrogen bondign C. Coordinate covalent bond D. Covalent bond
15	Diamond is a bad conductor because	A. It has a tight structure B. It has a high density C. there is no free electron present in the crystal of diamond of conduct electicity D. Is transparent to light
16	The boiling point of $NH_3$is maximum among the hydrides of group V elements due to	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of NH₃ C. Very small size of Nitrogen D. Enhanced electropositive character of Nitrogen
17	The inter-molecular forces in liquids are:	A. Negligible B. Very weak C. Very strong D. Reasonably strong
18	HF has exceptionally low acidic strengths due to	A. Smaller size of fluorine B. Strong polar bond between H and F C. Electronegativity of fluorine D. Strong hydrogen bonding
19	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
20	Evaporation of water is possible at	A. Above 100⁰ B. 0⁰ C. 100⁰ D. At all temperature