

## ECAT Chemistry Chapter 12 Periodic Classification of Elements and Periodicity Online Test

Sr	Questions	Answers Choice
1	NaH is	A. Ionic hydride B. Complex hydride C. Covalent hydride D. Interstitial hydride
2	Ionic hydrides are generally	A. Liquid at room temperature B. Good electrical conductors C. Good reducing agents D. Easily reduced
3	Li, Be, B, C, O, F, Ne are elements of	A. Second period B. First period C. Third period D. Fourth period
4	The elements of sub-group A are called	A. Transition elements B. Main elements C. Typical elements D. Rare earth elements
5	All the elements belongs to the 2nd period are	A. Normal elements B. Transition elements C. Stable elements D. Halogens
6	According to Mendeleev, the properties of the elements are periodic function of their	A. Atomic number B. Atomic volumes C. Atomic masses D. Atomic densities
7	Who gave the concept of atomic number	A. Newton B. Mosley C. Dalton D. Newland
8	The classify the elements, Newland gave the idea of	A. Octaves B. Triads C. Atomic volume D. Atomic mass
9	The statement that the properties of every eight elements are similar to the first is the law of	A. Dobereiner B. Newland C. Mendeleev D. L. Meyer
10	From 39Y to 48Cd are called	A. Transition elements B. Outer transition elements C. Inner transition elements D. 2nd transition series
11	Which of the following is not true for metalloids	A. They are borderline elements that exhibit both metallic and non-metallic properties to some extent B. They usually act as electron donors with non-metals and as electron acceptors with metals C. Some of these elements are boron, silicon and germanium D. They are good conductors of heat and electricity
12	Indicate the correct statement	A. All lanthanidees are present in the same group B. All halogens are present in the same period C. All the alkali metals are present in the same group D. All the noble gases are present in the same period
13	From left to right, atomic radii of transition elements	A. Increases B. Decreases C. Remain same D. None of the above

14	The coinage metals are	A. Ni, Pd, Pt B. Cu, Ag, Au C. Zn, Al, Pb D. Fe, Si, Sn
15	Which is the most volatile compound	A. HI B. HCl C. HBr D. HF
16	Mark the correct statement	A. Na <sup>+</sup> is smaller than Na atom B. Na <sup>+</sup> is larger than Na atom C. Cl <sup>-</sup> is the smaller than Cl atom D. Cl <sup>-</sup> (ion) and Cl (atom) are equal in size
17	Keeping in view the size of atoms, which order is the correct one	A. Mg > Sr B. Ba > Mg C. Lu > Ce D. Cl > I
18	The valency, ionization energy and electronegativity of elements are related to its	A. Atomic number B. Properties C. Atomic weight D. Family group
19	Which of the following has highest oxidation potential	A. Be B. Li C. Na D. Ca
20	The valence shell of hydrogen is half filled like those of	A. IV - A B. VIA C. V - A D. VIIA
21	Which of the following statement about electron affinity of two elements is correct	A. Carbon has greater than oxygen B. Sulphur has less than oxygen C. Iodine has greater than bromine D. Bromine has less than chlorine
22	The structure of complex hydrides is	A. Tetrahedral B. Trigonal C. Octahedral D. Square planar
23	For the representative elements from left to right across a period in the periodic table, the electron affinity of the atom generally	A. Increases B. Remains constant C. Decreases D. Not clear
24	The elements of f-block are also known as	A. Inner-transition B. Outer transition C. Normal elements D. Alkaline earth metals
25	Elements in the same family have	A. Same atomic number B. Molecular wt same C. Same chemical properties D. Same electronic configuration
26	According to the periodic law, the chemical properties of the elements are periodic functions of their	A. Density B. Atomic number C. Atomic mass D. Mass number
27	Which is not interstitial hydride	A. LaH B. VH C. TaH D. None
28	In a group from top to bottom, the hardness of alkali metals	A. Remains unchanged B. Increases C. Decreases D. None
29	Which has highest 1st I.E.	A. Br B. Cl C. F D. I
30	Each vertical column of the periodic table includes elements with chemical characteristics that are in general	A. Identical B. Similar C. Different D. Similar as well as different

