

ECAT Chemistry Chapter 11 Reaction Kinetics

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Sr	Questions	Answers Choice
1	In the reaction of oxalic acid with KMnO $_4$ and H $_2$ SO $_4$ is slow at the beginning but after sometimes the reaction becomes faster due to	A. Formation of MnSO ₄ which acts as 'Auto catalyst B. Formation of CO ₂ which acts as 'Auto catalyst C. Formation of K ₂ SO ₄ which acts as 'Auto catalyst D. Evolution of O ₂ gas which acts as 'Auto catalyst
2	Which of the following may affect the rate constant (k) fro a reaction :	A. Change in concentration. B. Change in pressure. C. Change in pH. D. Change in temperature.
3	Dilatometer method is useful for the reaction that involve :	A. Small volume changes in solutions B. Change in infractive indices C. Where reactants absorb U.V, visible or infrared radiation
4	The example of a photo chemical reaction is photosynthesis has order of reaction :	A. 1 B. 2 C. 0 D. 3
5	If half life period of a reaction is independent of the concentration of the reactants, then the reaction is	A. Zero order B. First order C. Second order D. Order is in fraction
6	The unit rate of constan K ismole $^{-1}$ dm 3 S $^{-1}$ for a chemical reaction, the order of reaction is :	A. 3 B. 2 C. 1 D. 0
7	The experimental relationship between a reaction rate and the concentration of reactants is called	A. Order or reaction B. Rate law C. Activated complex D. Molecularity
8	Question Image	A. 1st order B. 2nd order C. Zero order D. 3rd order
9	A white precipitate of silver chloride immediately formed on addition of:	A. Silver nitrate solution to sodium chloride solution. B. Silver chloride solution to sodium nitrate solution. C. Silver nitrate solution to potassium chloride solution D. Silver nitrate solution to hydrogen
		chloride solution.
10	When the rate of reaction is entirely independent of the conc. of reaction molecules then order of reaction is	A. Zero B. First C. Second D. Third
10		A. Zero B. First C. Second

13	Question Image	B. Pseudo first order C. Second order D. Zero order
14	The reaction rate is expressed in the units of	A. Mol dm ⁻³ B. Mol dm ⁻³ C. Mol dm ⁻³ N ^{N⁻³D. dm⁻³S⁻}
15	The rate of reaction determined at a given time is called	A. Average rate B. Instantaneous rate C. Specific rate D. Overall rate
16	If the rate equation of a reaction 2A+B>Products is , Rate = K[A] [B], and A is present in large excess, then order of reaction is :	A. 1 B. 2 C. 3 D. Above
17	A white precipitate of silver chloride immediately formed on addition of:	A. Silver nitrate solution to sodium chloride solution. B. Silver chloride solution to sodium nitrate solution. C. Silver nitrate solution to potassium chloride solution D. Silver nitrate solution to hydrogen chloride solution.
18	Rate of chemical reaction depends upon :	A. The number of total collisions per second. B. Number of molecules taking part in a chemical reaction. C. Number of fruitful collisions per second D. Number of fruitless collisions per second.
19	Question Image	A. Diastase B. Lipase C. Inverters D. Zymase
20	Question Image	A. Homogeneous B. Heterogeneous C. Isogeneous D. None