ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	Factor which slows down the rate of reaction is	A. Small size of the particles of the reactant B. High temperature of reaction C. More concentration of reactant D. Lowering the temperature
2	Decreases om concentration of reactant is denoted by	A. dc/dt B. -dc/dt C. +dc/dt D. None
3	Value of rate constant k is specific for a reaction, and varies from reaction to reaction. The value of k of a reaction changes with	A. Time B. Temperature C. Concentration of reactants D. Order of reaction
4	A substance which increases the rate of a reaction without being consumed during the reaction is called	A. An autocatalyst B. A catalyst C. A negative catalyst D. All of these
5	Question Image	A. Zero B. 1 C. 2 D. 1.5
6	Which of the following is not affected by light	
7	If the rate equation of a reaction 2A+B--------->Products is , Rate = K[A] [B], and A is present in large excess, then order of reaction is :	A. 1 B. 2 C. 3 D. Above
8	The factor which effect the rate of reaction	A. Nature of reactants B. Surface area C. Light D. All of the above
9	If initial concentration of the reactants and half life period of the reaction is known, then we can determine	A. Average rate of reaction B. Order of reaction C. Rate constant k D. Instantaneous rate
10	The rate of reaction determined at a given time is called	A. Average rate B. Instantaneous rate C. Specific rate D. Overall rate
11	The actual number of atoms or molecules taking part in rate determining step is	A. Rate of reaction B. Velocity of reaction C. Order of reaction D. Molecularly
12	Which technique is used to determine the absorption of radiations?	A. Dilatometer method<o:p></o:p> B. Optical rotation method<o:p></o:p> C. Spectrometry<o:p></o:p> D. Refractometric method<o:p></o:p>
13	Refrectrometric method is used when	A. Reactions involving absorption of I.R. or U. V B. Reactions involving change of refractive index C. Reactions involving ions D. Change of optical activity
14	The rate determining step is the :	A. Slowest step. B. Fastest step. C. Moderate step.

		D. Both (a) and (b).
15	If the rate of reaction is independent of the concentration of the reactant, the reaction is of	A. Zero order B. First order C. Second order D. Third order
16	In thermal decomposition of N_2O the half life period for two different initial concentrations of N_2O are (i) 255 second for initial N_2O 290 mm Hg (ii) 212 second for initial N_2O 360 mm Hg then it is	A. Zero order B. First order C. Second order D. Third order
17	The value of activation energy Ea of a reaction can be determined from the value of slope of the straight line obtained by plotting a graph between 1/T and log k. the value of Ea is equal to	A. Slope B. 1/Slope C. Slope x R D. Slope x 2.303 R
18	Which statement is true about order of reaction :	A. Order of reaction can only be determined by an experiment. B. Order of reaction can be determined from a balance equation only. C. Order of reaction increase by increasing temperature. D. Order of reaction must be in whole number and not in fraction.
19	The addition of a catalyst to a reaction changes the	A. Enthalpy B. Entropy C. Nature of reactants D. Energy of activation
20	If half life period of a reaction is independent of the concentration of the reactants, then the reaction is	A. Zero order B. First order C. Second order D. Order is in fraction