

ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	When copper is allowed to react with HNO ₃ , the reaction is slow in the beginning, finally becomes very fast. It is due to the formation of an auto catalyst which is	A. Cu(NO ₃) ₂ B. CuO C. O ₂ D. HNO ₂
2	The number of atoms or molecules whose concentrations determine the rate of the reaction is called	A. Molecularity B. Order C. Rate of reaction D. Rate constant
3	Half life period of a reaction is inversely proportion to the initial concentration of the reactant, then order of reaction is	A. Third order B. Second order C. Fist order D. Zero order
4	Dilatometer method is useful for the reaction that involve :	A. Small volume changes in solutions B. Change in infractive indices C. Where reactants absorb U.V, visible or infrared radiation
5	In the hydrolysis of CH ₃ COOC ₂ H ₅ the acid produced is	A. Inhibitor B. Catalyst C. Auto catalyst D. None of above
6	A substance which increases the rate of a reaction without being consumed during the reaction is called	A. An autocatalyst B. A catalyst C. A negative catalyst D. All of these
7	Question Image	A. 2 B. 3 C. 4 D. 9
8	The change in concentration of reactant or product per unit time is called :	A. Rate constant. B. Rate of reaction. C. Rate equation. D. Rate law.
9	Hydrogenation of vegetable oils is accelerated by Ni catalyst. The catalytic activity of Bi is increased by a promoter of activator which is	A. Na and K B. Na and Hg C. Hg and Zn D. Cu and Te
10	Group I-A elements react with water fastly than the reaction of group II-A elements because	A. I _A elements are more soft then II _A B. I _A elements are non-metals C. I _A elements have 1 electron in their outermost s-orbital and are strongly electropositive D. I _A elements make ionic bond
11	The catalytic activity of Pt is much higher when	A. It is mixed with asbestos B. It is mixed with Pd C. It is mixed with arsenic D. In is made colloidal platinum
12	It is common observation that rates of chemical reactions differ :	A. Greatly. B. A little bit. C. Moderately.
13	The unit of rate constant K ismole $^{-1}$ dm 3 for a chemical reaction, the order of reaction is :	A. Order of reaction can determined by an experiment B. Order of reaction can determined from a balance equation only. C. Order of reaction can determined increases by increasing temperate. D. Order of reaction must be in whole number and not in fraction.
14	A catalyst is a substance which increase the rate of a chemical reaction, but remains unchanged at the end of reaction, nut remains unchanged at the end of reaction, because	A. It increases the temperature B. It increase the surface area C. It increases the rate constant D. It decrease the energy energy of activation
		

15	The unit of rate constant is the same as that of the rate of reaction in :	A. First order reaction.B. Second order reaction.C. Zero order reaction.D. Third order reaction.
16	A zero order reaction is one in which :	A. Rate is not affected by changing concentration of reactants. B. concentration of reactants do not change with the passage of time. C. Reactants do not react. D. One reactants in large excess.
17	Question Image	A. Small change in concentration of product B. Small time internal C. Co-efficient of the reactant D. Co-efficient of the product
18	The reaction rate is expressed in the units of	A. mol dm ⁻³ S ⁻ B. mol dm ⁻³ C. mol dm ⁻³ N ⁻ D. dm ⁻³ S ⁻
19	Question Image	A. Rate is independent of concentration of water since it is in excess B. Rate is independent of concentration of ester since it is in exces C. Rate depends upon the concentration of acid catalyst added D. Rate = k[CH ₃ COOC ₂ H ₅] ^{1/2} [H ₂ O] ^{1/2}
20	Optical rotation method is sued when	A. Reaction involve ions B. Change of refractive indices C. Reactions involving change of optical activity D. None of the above