

## ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	Question Image	A. Diastase B. Lipase C. Inverters D. Zymase
2	Half life period of $N_2O_5$ is 24 minutes and it remains same where we increase or decrease its initial concentration, then reactions	A. Zero order B. First order C. Second order D. Third order
3	With increases in temperature of 10 K of the reacting gases the rate of reaction is doubled because	A. Increase in number of collisions B. Number of molecules having energy more than $E_a$ is doubled C. Increase in order of reaction D. Increase in surface area
4	The effective activity of a metal catalyst is increased if it is in	A. Solid form B. Liquid state C. Gaseous state D. Finely divided form
5	In an experiment the concentration of a reactant 'A' is doubled the rate increases four times. If concentration is tripled, then rate increases nine times. Thus the rate is proportional to _____ of concentration of 'A'	A. Square root B. Square C. Twice D. Cube
6	Half life period of a reaction is inversely proportion to the initial concentration of the reactant, then order of reaction is	A. Third order B. Second order C. First order D. Zero order
7	A white precipitate of silver chloride immediately formed on addition of :	A. Silver nitrate solution to sodium chloride solution. B. Silver chloride solution to sodium nitrate solution. C. Silver nitrate solution to potassium chloride solution D. Silver nitrate solution to hydrogen chloride solution.
8	If half life period of a reaction is independent of the concentration of the reactants, then the reaction is	A. Zero order B. First order C. Second order D. Order is in fraction
9	The rate constant $k$ of a reaction activation energy $E_a$ and temperature are related by Arrhenius in the form of an equation which is	
10	The unit of rate of reaction is	A. mole $dm^{-3}$ B. mole $Kg^{-1}$ C. moles $dm^{-3} \cdot sec^{-1}$ D. grams $dm^{-3}$
11	The number of atoms or molecules whose concentrations determine the rate of the reaction is called	A. Molecularity B. Order C. Rate of reaction D. Rate constant
12	The rate of reaction :	A. Increase as the reaction proceeds. B. Decreases as the reaction proceeds. C. Remains the same as the reaction proceed. D. May decrease or increase as the reaction proceeds.
13	Question Image	A. 1st order B. 2nd order C. Zero order D. 3rd order
14	With the progressive of the reaction the slope of the curve between concentration of product and time	A. Gradually becomes more steep B. Gradually becomes less steep C. No change occurs in slope

		D. None of these occurs
15	The chemical method used for determination of rate of reaction is	A. Spectroscopic B. Conductimetric C. Refractometric D. Titration
16	Refractometric method is used when	A. Reactions involving absorption of I.R. or U. V B. Reactions involving change of refractive index C. Reactions involving ions D. Change of optical activity
17	which one of the following is a heteroheneous catalysis	
18	If the rate of decay of radioactive isotope decreases from 200 cpm to 25 cpm after 24 hours, what is its half life :	A. 8 hours B. 6 hours C. 4 hours D. 3 hours
19	Question Image <input type="text"/>	
20	Question Image <input type="text"/>	A. Homogeneous B. Heterogeneous C. Isogeneous D. None