ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	Hydrogenation of vegetable oils is accelerated by Ni catalyst. The catalytic activity of Bi is increased by a promoter of activator which is	A. Na and K B. Na and Hg C. Hg and Zn D. Cu and Te
2	Which technique is used to determine the absorption of radiations?	A. Dilatometer method<o:p></o:p> B. Optical rotation method<o:p></o:p> C. Spectrometry<o:p></o:p> D. Refractometric method<o:p></o:p>
3	Which of the following is not affected by light	
4	The rate of reaction :	A. Decreases as the reaction proceeds B. Increases as the reaction proceeds C. May decrease or increase reaction proceeds D. Remains same as the reaction proceeds
5	The change in concentration of reactant or product per unit time is called :	A. Rate constant. B. Rate of reaction. C. Rate equation. D. Rate law.
6	Which of the following factors does not influenced the rate of reaction	A. Concentration of the reaction B. Nature of the reactants C. Molecularity of the reaction D. Temperature
7	Question Image	A. Diastase B. Lipase C. Inverters D. Zymase
8	In the reaction of oxalic acid with $KMnO_4$and $H_2SO_4$is slow at the beginning but after sometimes the reaction becomes faster due to	A. Formation of MnSO₄which acts as 'Auto catalyst B. Formation of CO₂which acts as 'Auto catalyst C. Formation of K₂SO₄which acts as 'Auto catalyst D. Evolution of O₂gas which acts as 'Auto catalyst
9	By the use of catalysis the energy of activation is	A. Decreased B. Increased C. Not affected D. None
10	Question Image	
11	The unit of rate constant is the same as that of the rate of reaction in :	A. First order reaction. B. Second order reaction. C. Zero order reaction. D. Third order reaction.
12	Value of rate constant k is specific for a reaction, and varies from reaction to reaction. The value of k of a reaction changes with	A. Time B. Temperature C. Concentration of reactants D. Order of reaction
13	Which of the following may affect the rate constant (k) fro a reaction :	A. Change in concentration. B. Change in pressure. C. Change in pH. D. Change in temperature.
		A. k

14	To determine the rate of reaction chemically a graphical method is applied. A graph is plotted between the amount or reactant decomposed or product formed against the time. The rate d_x/d_t at any time is equal to	B. Tangentθ C. 1/a, a is initial conc. D. 1/a²
15	The addition of a catalyst to a reaction changes the	A. Enthalpy B. Entropy C. Nature of reactants D. Energy of activation
16	The rate of reaction determined at a given time is called	A. Average rate B. Instantaneous rate C. Specific rate D. Overall rate
17	A catalyst is a substance which increase the rate of a chemical reaction, but remains unchanged at the end of reaction, nut remains unchanged at the end of reaction, because	A. It increases the temperature B. It increase the surface area C. It increases the rate constant D. It decrease the energy energy of activation
18	The rate equation for a reaction is Rate =k[A]. what are unit of K ?	A. Mole-1 dm³ S⁻¹ <o:p></o:p> B. Mole dm³ S⁻¹ <o:p></o:p> C. Moledm³<o:p></o:p> D. S⁻¹<o:p></o:p>
19	When copper is allowed to react with HNO_3, the reaction is slow in the beginning, finally becomes very fast. It is due to the formation of an auto catalyst which is	A. Cu(NO₃)₂ B. CuO C. O₂ D. HNO₂
20	The unit of the rate constant is the same as that of rate of reaction in :	A. Third order reaction B. Second order reaction C. First order reaction D. Zero order reaction