

## ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	Which of the following will affect the rate :	<ul><li>A. First step of reaction.</li><li>B. Last step of reaction.</li><li>C. Rate determining step.</li><li>D. Fastest step.</li></ul>
2	Half life period of a reaction is inversely proportion to the initial concentration of the reactant, then order of reaction is	A. Third order B. Second order C. Fist order D. Zero order
3	Question Image	
4	The unit of rate constant k is the same as that of the rate of reaction in	A. First order reaction B. Second order reaction C. Third order reaction D. Zero order reaction
5	Activation energy is the difference of energy between the energy of the reactant and	A. The product B. The activated complex C. Both a and b D. None of these
6	A white precipitate <b>of</b> silver <b>chloride</b> immediately formed on <b>addition</b> of:	A. Silver nitrate solution to sodium chloride solution.     B. Silver chloride solution to sodium nitrate solution.     C. Silver nitrate solution to potassium chloride solution     D. Silver nitrate solution to hydrogen chloride solution.
7	The rate of reaction b/w two specific time intervals is called :	A. Instantaneous rate of reaction.     B. Average rate of reaction.     C. Rate of a reaction.     D. Minimum rate of a reaction.
8	which one of the following is a heteroheneous catalysis	
9	When the rate of reaction is entirely independent of the conc. of reaction molecules then order of reaction is	A. Zero B. First C. Second D. Third
10	Which of the following reactions occur at moderate rate :	A. Rusting of iron     B. Chemical weathering of stone work of buildings by acidic gases in atmosphere.     C. Hydrolysis of an ester     D. Fermentation of sugars
11	Refrectrometric method is used when	A. Reactions involving absorption of I.R. or U. V B. Reactions involving change of refractive index C. Reactions involving ions D. Change of optical activity
12	Question Image	A. Three times B. Six times C. Nine times D. Two times
13	Decreases om concentration of reactant is denoted by	A. dc/dt Bdc/dt C. +dc/dt D. None
14	Dilatometer method is useful for the reaction that involve :	A. Small volume changes in solutions B. Change in infractive indices C. Where reactants absorb U.V, visible or infrared radiation
15	A catalyst is a substance which increase the rate of a chemical reaction, but remains	A. It increases the temperature B. It increase the surface area C. It increases the rate constant

	activation
If a reactant or product of a reaction absorbs radiation, then physical method for determining the rate of reaction is	A. Spectrometry B. Refractometry C. Conductivity measurement D. Optical method
In the reaction of oxalic acid with KMnO4and H <sub>2</sub> SO <sub>4</sub> is slow at the beginning but after sometimes the reaction becomes faster due to	A. Formation of MnSO <sub>4</sub> which acts as 'Auto catalyst B. Formation of CO <sub>2</sub> which acts as 'Auto catalyst C. Formation of K <sub>2</sub> SO <sub>4</sub> which acts as 'Auto catalyst D. Evolution of O <sub>2</sub> gas which acts as 'Auto catalyst
The number of atoms or molecules whose concentrations determine the rate of the reaction is called	A. Molecularity B. Order C. Rate of reaction D. Rate constant
The reaction rate is expressed in the units of	A. mol dm <sup>-3</sup> S <sup>-</sup> B. mol dm <sup>-3</sup> C. mol dm <sup>-3</sup> N <sup>-</sup> D. dm <sup>-3</sup> S <sup>-</sup>
In thermal decomposition of $N_2O$ the half life period for two different initial concentrations of $N_2O$ are (i) 255 second for initial $N_2O$ 290 mm Hg (ii) 212 second for initial $N_2O$ 360 mm Hg then it is	A. Zero order B. First order C. Second order D. Third order
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