ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	The unit of rate constant K is $mole^{-1}$ dm^3 for a chemical reaction, the order of reaction is :	A. Order of reaction can determined by an experiment B. Order of reaction can determined from a balance equation only. C. Order of reaction can determined increases by increasing temperate. D. Order of reaction must be in whole number and not in fraction.
2	A white precipitate ofsilver chlorideimmediately formed on additionof :	A. Silver nitrate solution to sodium chloride solution. B. Silver chloride solution to sodium nitrate solution. C. Silver nitrate solution to potassium chloride solution D. Silver nitrate solution to hydrogen chloride solution.
3	The reaction rate is expressed in the units of	A. Mol dm⁻³S⁻ B. Mol dm⁻³ C. Mol dm⁻³N⁻ D. dm⁻³S⁻
4	In thermal decomposition of N_2O the half life period for two different initial concentrations of N_2O are (i) 255 second for initial N_2O 290 mm Hg (ii) 212 second for initial N_2O 360 mm Hg then it is	A. Zero order B. First order C. Second order D. Third order
5	which one of the following is a heteroheneous catalysis	
6	Which statement is not correct	A. Enzymes catalyst a specific reaction B. Enzymes show catalytic activity at a specific temperature C. The catalytic activity of enzymes is stopped if optimum pH is changed D. The catalytic activity is poisoned by a co-enzymes
7	The experimental relationship between a reaction rate and the concentration of reactants is known as	A. Order B. Molecularity C. Rate constant D. Rate law
8	The addition of a catalyst to a reaction changes the	A. Enthalpy B. Entropy C. Nature of reactants D. Energy of activation
9	When copper is allowed to react with HNO_3, the reaction is slow in the beginning, finally becomes very fast. It is due to the formation of an auto catalyst which is	A. Cu(NO₃)₂ B. CuO C. O₂ D. HNO₂
10	In zero order reaction, the rate is independent of :	A. Temperature of reaction B. Concentration ofreactants C. Concentration of products D. None of these.
11	Question Image	A. Small change in concentration of product B. Small time internal C. Co-efficient of the reactant D. Co-efficient of the product
12	The rate constant k of a reaction activation energy Ea and temperature are related by Arrhenius in the form of an equation which is	
		A. Decreases as the reaction

13	The rate of reaction :	proceeds B. Increases as the reaction proceeds C. May decrease or increase reaction proceeds D. Remains same as the reaction proceeds
14	_____ are called biocatalysts	A. Organic acids B. Organic bases C. Enzymes D. All
15	A substance which increases the rate of a reaction without being consumed during the reaction is called	A. An autocatalyst B. A catalyst C. A negative catalyst D. All of these
16	A zero order reaction is one in which :	A. Rate is not affected by changing concentration of reactants. B. concentration of reactants do not change with the passage of time. C. Reactants do not react. D. One reactants in large excess.
17	Which one of the following reaction rate is effected by the light	
18	If the rate of decay of radioactive isotope decreases from 200 cpm to 25 cpm after 24 hours, what is its half life :	A. 8 hours B. 6 hours C. 4 hours D. 3 hours
19	The sum of the exponents of the conc. terms in the rate equation is called	A. Rate of reaction B. Order of reaction C. Specific rate constant D. Average rate
20	Question Image	A. Rate = k[FeCl₃] [KI] ² B. Rate = k[Fe⁺³] [Cl⁻¹] [KI] C. Rate = k[Fe⁺³] [Cl⁻¹][K;] D. Rate = k[KI]³ [FeCl₃]°