

ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	Half life period of N_2O_5 is 24 minutes and it remains same where we increase or decrease its initial concentration, then reactions	A. Zero order B. First order C. Second order D. Third order
2	The sum of the exponents of the conc. terms in the rate equation is called	A. Rate of reaction B. Order of reaction C. Specific rate constant D. Average rate
3	The chemical method used for determination of rate of reaction is	A. Spectroscopic B. Conductiometric C. Refractometric D. Titration
4	Question Image	A. Diastase B. Lipase C. Inverters D. Zymase
5	Which of the following may affect the rate constant (k) fro a reaction :	A. Change in concentration.B. Change in pressure.C. Change in pH.D. Change in temperature.
6	When the rate of reaction is entirely independent of the conc. of reaction molecules then order of reaction is	A. Zero B. First C. Second D. Third
7	The rate of reaction determined at a given time is called	A. Average rate B. Instantaneous rate C. Specific rate D. Overall rate
8	To determine the rate of reaction chemically a graphical method is applied. A graph is plotted between the amount or reactant decomposed or product formed against the time. The rate $d_\chi d_t$ at any time is equal to	A. k B. Tangent θ C. 1/a, a is initial conc. D. 1/a ²
		A. Small change in concentration of
9	Question Image	product B. Small time internal C. Co-efficient of the reactant D. Co-efficient of the product
10	The rate of reaction b/w two specific time intervals is called :	A. Instantaneous rate of reaction. B. Average rate of reaction. C. Rate of a reaction. D. Minimum rate of a reaction.
11	Question Image	A. Zero B. 1 C. 2 D. 1.5
12	In the reaction of oxalic acid with KMnO4and H_2SO_4 is slow at the beginning but after sometimes the reaction becomes faster due to	A. Formation of MnSO ₄ which acts as 'Auto catalyst B. Formation of CO ₂ which acts as 'Auto catalyst C. Formation of K ₂ SO ₄ which acts as 'Auto catalyst D. Evolution of O ₂ gas which acts as 'Auto catalyst
13	In an experiment the concentration of a reactant 'A' is doubled the rate increases four times. If concentration in tripled, then rate increases nint times. Thus the rate is proportional to of concentration of 'A'	A. Square root B. Square C. Twice

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14	are called biocatalysts	A. Organic acids B. Organic bases C. Enzymes D. All
15	Value of rate constant k is specific for a reaction, and varies from reaction to reaction. The value of k of a reaction changes with	A. Time B. Temperature C. Concentration of reactants D. Order of reaction
16	The experimental relationship between a reaction rate and the concentration of reactants is known as	A. Order B. Molecularity C. Rate constant D. Rate law
17	If the rate equation of a reaction 2A+B>Products is , Rate = K[A] [B], and A is present in large excess, then order of reaction is :	A. 1 B. 2 C. 3 D. Above
18	If initial concentration of the reactants and half life period of the reaction is known, then we can determine	A. Average rate of reaction B. Order of reaction C. Rate constant k D. Instantaneous rate
19	Hydrogenation of vegetable oils is accelerated by Ni catalyst. The catalytic activity of Bi is increased by a promoter of activator which is	A. Na and K B. Na and Hg C. Hg and Zn D. Cu and Te
20	All reactions occur in :	A. A single step. B. A series of steps C. Two steps. D. Both (a) and (b)