ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	In a compound an atom has negative oxidation state because	A. Atom is negatively charged B. Atom acts as cathode C. Atom is more electronegative D. Atom has lowest ionization energy
2	Calculate the amount of charge flowing in 2 minute in a wire of resistance $10\Omega$when a potential difference of 20 V is applied	A. 120 C B. 240 C C. 20 C D. 4 C
3	Which of the following cannot conduct electricity :	A. NaCl fused. B. NaCl(aq) _{<o:p></o:p>} C. NaCl(Solid) _{<o:p></o:p>} <o:p></o:p> D. Both (b) and (c)
4	What are the products electrolysis of aqueous sodium chloride at two electrodes	A. Chlorine at anode and oxygen at cathode. B. Hydrogen at anode and chlorine at cathode. C. Chlorine at anode and hydrogen at cathode. D. Chlorine at anode and sodium at cathode.
5	In the presence of copper electrodes the electrolysis of aqueous $CuSO_4$produces at cathode	A. H₂gas B. O₂gas C. SO₂gas D. Cu metal
6	A cell in which electric current is produced as a result spontaneous redox reaction is called :	A. Dry cell B. Electrolytic cell C. Galvanic cell D. Standard cell
7	Corrosion reaction are	A. Spontaneous redox reaction B. Non-spontaneous acid-base reactions C. Spontaneous acid-base reactions D. None of these
8	Question Image	A. Iodine can oxidise iron B. Bromine can oxidise iron C. Iodine can oxidise bromine D. Iron can oxidise bromine
9	Strong oxidizing agents have	A. Greater positive value of standard reduction potential B. Lesser positive value of standard C. Greater negative value of standard D. None of these
10	Which statement is incorrect for NICAD battery	A. The electrolyte is alkali B. Cd acts as anode C. MnO₂acts as electrolyte D. NiO₂acts as cathode
11	Which of the following statement is incorrect about SHE(Standard hydrogen electrode):	A. Reduction potential of Cu⁺² is smaller than H⁺ions whenit is coupled with copper electrode. B. gas is passed in it at 1 atm pressure. C. Its oxidation potential and reduction potential is zero. D. It is made of platinum wire dipped in HCl solution
12	Which statement is incorrect for balancing of redox reactions by ion-electron method	A. The reaction is splitted into two half reactions B. H₂O and H⁺ions are added for acidic or neutral reaction to balance O and H atoms C. To balance H, HCi, is added D. To balance O and H in the alkaline reaction OH⁻added
		A. Hydrolysis

13	When electric current is used to carry out non-spontaneous redox, the process is called	B. Electrolysis C. Decomposition D. Neutralization
14	Fe can displace Cu form CuSO4solution because	A. Fe is ferromagnetic B. Fe is below Cu in electrochemical series C. Fe is above Cu in electrochemical series D. Fe exists in divalent oxidation state
15	During electrolysis, the reaction that takes place at cathode is :	A. Reduction B. Both (a) and (c) C. Oxidation D. No reaction occurs
16	A standard hydrogen electrode is used as standard electrode of which electrode potential is arbitrarily taken as	A. +1 B. -1 C. 0.1 D. Zero
17	Metallic conduction is due to the	A. Movement of electrons B. Movement of ions C. Both a and b D. None of these
18	The best electrode used in salt bridge is KCl. Which other electrolyte an also be used for the purpose :	A. NaCl B. NH₄NO₃<o:p></o:p> C. KNO₃<o:p></o:p> D. NaNO₃<o:p></o:p>
19	Matals usually conduct electricity because	A. There are mobile electrons in the metallic structure B. Metals are decomposed by current C. Metals have high resistance D. In metals the ions are free to move
20	In passage of electricity through aqueous solution of $AgNO_3$silver dissolves at anode to form Ag^+, the electrodes are	A. Silver metal B. Pt metal C. Graphite D. Copper metal