

## ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	The cathodic reaction in the electrolysis of dill H <sub>2</sub> SO <sub>4</sub> with Pt electrodes	A. Reduction B. Oxidation C. Both oxidation and reduction D. neither oxidation or reduction
2	In KO <sub>2</sub> the oxidation state of oxygen is	A2 B1 C. +1/2 D1/2
3	Zn does not displace Mg from MgSO <sub>4</sub> solution because	A. Zn is more electropositive than Mg     B. Zn is below Mg in electropositive     series     C. Zn is above Mg in electrochemical     series     D. Zn is trivalent Mg is divalent
4	$F_2$ , $Cl_2$ , $Br_2$ and $l_2$ lie below SHE is the Electro chemical series that is why these	A. Undergo reduction B. Undergo oxidation C. Liberate H <sub>2</sub> gas with steam D. None of these
5	Time required to deposit one millimole of aluminium metal by the passage of 9.65 amperes through molten electrolyte containing aluminium ion is	A. 30 s B. 10 s C. 30,000 s D. 10,000 s
6	lonization is the process in which ionic compounds when fused or dissolved in water split up into charged particles called :	A. Atoms. B. Electrons. C. Protons . D. lons
7	Purification of an impure copper is made by electrolytic cell, in which impure copper is anode and pure copper is cathode, and the electrolyte used is	A. H <sub>2</sub> SO <sub>4</sub> B. CuSO <sub>4</sub> C. ZnSO <sub>4</sub> D. Na <sub>2</sub> SO <sub>4</sub>
8	Which one of the following reaction takes place spontaneously	
9	Which statement is incorrect for NICAD battery	A. The electrolyte is alkali B. Cd acts as anode C. MnO <sub>2</sub> acts as electrolyte D. NiO <sub>2</sub> acts as cathode
10	When fused PbBr2 is electrolyzed :	A. Lead appears at anode.     B. Lead appears at cathode.     C. Bromine appears at cathode.     D. Lead appears at both electrodes.
11	Metals like Fe, Mg, Al, Cr, Zn have more negative reduction potentials that is whey	A. These don't react with steam B. These react very slowly with steam to liberate H <sub>2</sub> gas C. These react rapidly with steam to produce the metallic oxides and liberate H <sub>2</sub> gas D. These react with cold water violently
12	A standard hydrogen electrode (S.H.E) consists of a platinized platinum electrode dipped in 1 molar solution of H <sup>+</sup> ions and hydrogen gas is passed at a pressure of	A. One pascal B. One kilo pascal C. One atmoshpere D. Then atmoshpere
13	96500 C electricity is passed through CuSO <sub>4</sub> . The amount of copper precipitated is	A. 0.25 mole B. 0.5 mole C. 1.0 mole D. 2.00 mole

15	Metals which are above SHE in electrochemical series	A. Can liberate H <sub>2</sub> fror acid B. Cannot liberate H <sub>2</sub> from acid C. Cannot always liberate H <sub>2</sub> from acid D. None of these
16	The oxidation number of H is -1 in the compound	A. H <sub>2</sub> O B. H <sub>3</sub> BO <sub>3</sub>
17	The number of coulombs required for the deposition of 107.870 g of silver is	A. 96500 B. 48250 C. 193000 D. 10000
18	Question Image	A. Adding H <sub>2</sub> O and H <sup>+</sup> ions B. Adding OH <sup>-</sup> ions C. Adding O <sup>2</sup> molecule D. Adding O atoms
19	During electrolysis, the reaction that takes place at cathode is :	A. Reduction B. Both (a) and (c) C. Oxidation D. No reaction occurs
20	The process of electrical coating of one metal on another to protect, decorate or to have greater resistance to corrosion is called	A. Electroplating B. Electrolysis C. Conduction D. Induction