ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	Out of Cu, Ag, Fe and Zn the metal which can displace all others from theri salt solution is	A. Ag B. Cu C. Zn D. Fe
2	The standard e.m.f. of a galvanic cell involving cell reaction with n = 2 is found to be 0.2965 V at 25°C. The equilibrium constant of the reaction would be	A. 1.0 x 10¹⁰ B. 2.0 x 10¹¹ C. 4.0 x 10¹² D. 1.0 x 10²
3	Which statement is incorrect about standard hydrogen electrode	A. Its oxidation and reduction potential is zero B. It consists of Pt wire dipped on 1 molar HCl solution C. The electrolyte is 1 molar NaOH solution D. H₂gas is passes in it at 1 atmospheric pressure
4	In a solution of $CuSO_4$how much time will be required to precipitate 2g copper by 0.5 ampere current?	A. 12157.48 sec B. 102 sec C. 510 sec D. 642 sec
5	F_2, Cl_2, $Br_2$and I_2lie below SHE is the Electro chemical series that is why these	A. Undergo reduction B. Undergo oxidation C. Liberate H₂gas with steam D. None of these
6	Sodium can be obtained by :	A. Electrolysis of acidified water. B. By heating NaCl and water at 100° _{<o:p></o:p>} C. Electrolysis of molten sodium chloride. D. Electrolysis of aqueous sodium chloride.
7	Which of the following cannot conduct electricity :	A. NaCl fused. B. NaCl(aq) _{<o:p></o:p>} C. NaCl(Solid) _{<o:p></o:p>} <o:p></o:p> D. Both (b) and (c)
8	The conductivity of strong electrolyte	A. Increases on dilution slightly B. Does not change on dilution C. Decreases on dilution D. Depends on density of electrolyte it self
9	Time required to deposit one millimole of aluminium metal by the passage of 9.65 amperes through molten electrolyte containing aluminium ion is	A. 30 s B. 10 s C. 30,000 s D. 10,000 s
10	The oxidation number of free element is always taken to be	A. 0 B. 1 C. 2 D. -1
		A. 4.5 V

11	Alkaline battery has a voltage of	B. 3.5 V C. 2.5 V D. 1.5 V
12	Corrosion is basically a	A. Altered reaction in presence of H₂O B. Electrochemical phenomenon C. Interaction D. Union between two light metals and a heavy metal
13	Coupling of Pb with its Pb^{++}/Pb = -0.13 V and Ag with Ag^{+}/ Ag = +0.80 V, the cell reaction	
14	Metals which are above SHE in electrochemical series	A. Can liberate H₂from acid B. Cannot liberate H₂from acid C. Cannot always liberate H₂from acid D. None of these
15	Metallic conduction is due to the	A. Movement of electrons B. Movement of ions C. Both a and b D. None of these
16	Reduction or oxidation potential of standard hydrogen electrode is :	A. 0.0 Volt B. 0.8Volt C. 1.0Volt D. 1.8Volt
17	When fused PbBr2 is electrolyzed :	A. Lead appears at anode. B. Lead appears at cathode. C. Bromine appears at cathode. D. Lead appears at both electrodes.
18	Purification of an impure copper is made by electrolytic cell, in which impure copper is anode and pure copper is cathode, and the electrolyte used is	A. H₂SO₄ B. CuSO₄ C. ZnSO₄ D. Na₂SO₄
19	In the electrolysis of aqueous solution of sodium nitrate, the ions which are reduced at the cathode are	A. H₃O⁺ B. Na⁺ C. OH⁻ D. NO³⁻
20	Electrochemical series is a list of element S arranged into the increasing order of their	A. Standard oxidation potential B. Standard reduction potential C. Cell voltage D. Ionization potential