ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	Question Image	A. Cu B. H C. N D. O
2	The amount of electricity that can deposit 108 g of silver from silver nitrate solution is	A. 1 ampere B. 1 coulomb C. 1 faraday D. 2 ampere
3	The specific conductance of 0.1 M NaCl solution is 1.06 x 10^{-2}ohm^{-1}mol^{-1}. Its molar conductance in ohm^{-1}cm^{2}mol^{-1}is	A. 1.06 x 10² B. 1.06 x 10³ C. 1.06 x 10⁴ D. 53
4	A dry alkaline cell has porous Zn anode and MnO_2as cathode the electrolyte used is	A. Ca(OH)₂ B. NaOH C. KOH D. NH₄OH
5	Ionization is the process in which ionic compounds when fused or dissolved in water split up into charged particles called :	A. Atoms. B. Electrons. C. Protons . D. Ions
6	When fused $PbBr_2$is electrolyzed	A. Bromine appears at he cathode B. Lead is deposited at the cathode C. Lead appears at the anode D. None of these happens
7	The standard e.m.f. of a galvanic cell involving cell reaction with n = 2 is found to be 0.2965 V at 25°C. The equilibrium constant of the reaction would be	A. 1.0 x 10¹⁰ B. 2.0 x 10¹¹ C. 4.0 x 10¹² D. 1.0 x 10²
8	Cell in which an electric current drives a non-spontaneous reaction is called	A. Electrolytic cell B. Voltaic cell C. Biological cell D. Electrochemical cell
9	Cell potential depends upon :	A. Concentration of ions B. Nature of electrolyte C. Temperature D. All of above
10	Which statement is incorrect for balancing of redox reactions by ion-electron method	A. The reaction is splitted into two half reactions B. H₂O and H⁺ions are added for acidic or neutral reaction to balance O and H atoms C. To balance H, HCi, is added D. To balance O and H in the alkaline reaction OH⁻added
11	Metallic conduction is also called as :	A. Ionic conduction. B. Protonic conduction. C. Electronic conduction D. Super conduction
12	Lead accumulator stops discharging current when	A. Lead at anode converted to PbO₂ B. PbO₂at cathode converted to Pb C. Both electrodes are completely covered with PbSO₄ D. Both electrodes are completely covered with Pb(OH)₂
13	The cathodic reaction in the electrolysis of dill $H_2SO_4$with Pt electrodes	A. Reduction B. Oxidation C. Both oxidation and reduction D. neither oxidation or reduction
		A. Ionization

14	Reaction takes place at anode is	B. Reduction C. Oxidation D. Hydrolysis
15	2.5 faradays of electricity is passed through solution of $CuSO_4$. The number of gram equivalents of copper deposited on the cathode would be	A. 1 B. 2 C. 2.5 D. 1.25
16	Which of the following is the use of electrolysis?	A. Eletrorefining B. Electroplating C. Both A and B D. None of the above
17	The electrode through which the electrons enter the electrolytic solution is electrolytic solution is	A. Anode B. Cathode C. May be anode or cathode D. None of these
18	In electrolytic cells, the chemical changes may be :	A. Either spontaneous or non-spontaneous B. Always spontaneous C. Always non-spontaneous D. More spontaneous and less non-spontaneous.
19	Which one of the following reaction takes place spontaneously	
20	Out of Cu, Ag, Fe and Zn the metal which can displace all others from theri salt solution is	A. Ag B. Cu C. Zn D. Fe