ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	Which of the following is a strong electrolyte?	A. Ca(NO₃)₂ B. HCN C. CH₃COOH D. NH₄OH
2	The unit of specific conductivity is	A. Ohm cm⁻¹ B. Ohm cm⁻² C. Ohm⁻¹cm D. Ohm⁻¹cm⁻¹
3	When fused PbBr2 is electrolyzed :	A. Lead appears at anode. B. Lead appears at cathode. C. Bromine appears at cathode. D. Lead appears at both electrodes.
4	If the standard electrode potential of CU^{2+}/Cu electrode is 0.34 V, what is the electrode potential of 0.01 M concentration of Cu^{2+}? (T=298)	A. 0.399 V B. 0.281 V C. 0.222 V D. 0.176 V
5	Which of the following statement is incorrect about SHE(Standard hydrogen electrode):	A. Reduction potential of Cu⁺² is smaller than H⁺ions whenit is coupled with copper electrode. B. gas is passed in it at 1 atm pressure. C. Its oxidation potential and reduction potential is zero. D. It is made of platinum wire dipped in HCl solution
6	The number of coulombs required for the deposition of 107.870 g of silver is	A. 96500 B. 48250 C. 193000 D. 10000
7	A standard hydrogen electrode is used as standard electrode of which electrode potential is arbitrarily taken as	A. +1 B. -1 C. 0.1 D. Zero
8	A standard hydrogen electrode (S.H.E) consists of a platinized platinum electrode dipped in 1 molar solution of H^+ions and hydrogen gas is passed at a pressure of	A. One pascal B. One kilo pascal C. One atmoshpere D. Then atmoshpere
9	K,Ca and Li metals may be arranged in decreasing order of their reduction potential as :	A. Li, k, Ca B. Ca, K, Li C. Li, Ca, K D. K, Ca, Li
10	Which of the following statements is true about Galvanic cell	A. Anode is negatively charged B. Reduction occurs at anode C. Cathode is positively charge D. Reduction occurs at cathode
11	The art of electroplating was given by	A. Faraday B. Edison C. Thomas Gradam D. Brugan
12	Best way to prevent rusting of iron is by	A. Making iron cathode B. Putting it in saline water C. Both of these D. None of these
13	Strong reducing agents gave	A. Greater positive value of standard reduction potential B. Greater negative value of standard reduction potential C. Lesser positive value of standard reaction potential D. None of these

14	A cell in which electric current is produced as a result spontaneous redox reaction is called :	A. Dry cell B. Electrolytic cell C. Galvanic cell D. Standard cell
15	A solution of sodium sulphate was electrolysed using some insert electrodes. The products at the electrodes are	A. O₂, H₂ B. O₂, Na C. O₂, SO₂ D. O₂, S₂O₈²
16	Which statement is incorrect for NICAD battery	A. The electrolyte is alkali B. Cd acts as anode C. MnO₂acts as electrolyte D. NiO₂acts as cathode
17	Electrochemical series is a list of element S arranged into the increasing order of their	A. Standard oxidation potential B. Standard reduction potential C. Cell voltage D. Ionization potential
18	The cell which generates electricity as a result of spontaneous oxidation-reduction reaction is called	A. Electrolytic cell B. Nelson's cell C. Galvanic cell D. Down's cell
19	The two half cells of a galvanic cell are connected by	A. Ammeter B. Salt bridge C. Hydrogen electrode D. Copper electrode
20	The value of SHE is cathode and anode is always taken to be	A. One B. Zero C. Different D. Same