ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	If the salt bridge is not used between two half cells, then the voltage	A. Decrease rapidly B. Decrease slowly C. Does not change D. Drops to zero
2	The electrode through which the electrons enter the electrolytic solution is electrolytic solution is	A. Anode B. Cathode C. May be anode or cathode D. None of these
3	A cell in which electric current is produced as a result spontaneous redox reaction is called :	A. Dry cell B. Electrolytic cell C. Galvanic cell D. Standard cell
4	Electrochemical series is a list of element S arranged into the increasing order of their	A. Standard oxidation potential B. Standard reduction potential C. Cell voltage D. Ionization potential
5	A cell constant is generally found by measuring the conductivity of aqueous solution of	A. BaCl₂ B. KCl C. NaCl D. MgCl₂
6	In passage of electricity through aqueous solution of $AgNO_3$silver dissolves at anode to form Ag^+, the electrodes are	A. Silver metal B. Pt metal C. Graphite D. Copper metal
7	Metals like Fe, Mg, Al, Cr, Zn have more negative reduction potentials that is whey	A. These don't react with steam B. These react very slowly with steam to liberate H₂gas C. These react rapidly with steam to produce the metallic oxides and liberate H₂gas D. These react with cold water violently
8	Question Image	A. Adding H₂O and H⁺ions B. Adding OH⁻ions C. Adding O²molecules D. Adding O atoms
9	Oxidation number of oxygen in $OF_2$is	A. +1 B. -1 C. +2 D. -2
10	F_2, Cl_2, $Br_2$and I_2lie below SHE is the Electro chemical series that is why these	A. Undergo reduction B. Undergo oxidation C. Liberate H₂gas with steam D. None of these
11	Out of Cu, Ag, Fe and Zn the metal which can displace all others from theri salt solution is	A. Ag B. Cu C. Zn D. Fe
12	In $KO_2$the oxidation state of oxygen is	A. -2 B. -1 C. +1/2 D. -1/2
13	During redox reaction an oxidizing agent	A. Gains electrons B. Is oxidized C. Loses electrons D. Hydrolysed
14	In electronic watches or electronic calculators the tiny batteries used are	A. Alkaline battery B. NICAD battery C. Fuel cell

		D. Silver oxide battery
15	Prevention of corrosion of iron by Zn coating is called	A. Galvanization B. Cathodic protection C. Electrolysis D. Photoelectrolysis
16	Best way to prevent rusting of iron is by	A. Making iron cathode B. Putting it in saline water C. Both of these D. None of these
17	The function of salt bridge in the galvanic or voltaic cell is to	A. Carry out oxidation at anode B. To carry out reduction at cathode C. Carry out electrolysis D. To prevent the net charge accumulation in either of the half cells
18	The cell in which a non-spontaneous redox reaction takes place as a result electricity is known as :	A. Electrolytic cell. B. Voltaic cell. C. Daniel cell. D. Dry cell.
19	The specific conductance of 0.1 M NaCl solution is 1.06 x $10^{-2}ohm^{-1}mol^{-1}$. Its molar conductance in $ohm^{-1}cm^{2}mol^{-1}$is	A. 1.06 x 10² B. 1.06 x 10³ C. 1.06 x 10⁴ D. 53
20	The substance having highest conductivity at room temperature among the following is	A. 0.1 N HCl B. 0.1 N NaCl C. Graphite D. Glass