

Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	If the solvent is inflammable for heating purpose we use	A. Ice bath B. Water bath C. Wire gauze D. Thermostat
2	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
3	Ethyl chloride on treatment with aqueous alkali gives	A. Ethane B. Ethene C. Ethanal D. Ethanol
4	Which has largest first ionization energy?	A. Li B. Na C. K D. Rb
5	Which of the following will not change the concentration of ammonia in the equilibrium $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$; $\Delta H = -kJ$	A. Increase of pressure B. Increase of temperature C. Decrease of volume D. Addition of catalyst
6	Hybridization explain the----- of orbitals	A. Type of Bonding B. Shapes C. Shape and Type of bonding D. None of above
7	Bell metal is an alloy of	A. Cu,Zn, and Sn B. Cu,Zn and Ni C. Cu and Zn D. Cu and Sn
8	According to MO Theory the species O_2^+ possesses	A. Bond order of 2.5 B. Three unpaired electrons C. Diamagnetic character D. Stability lower then O_2
9	Atmosphere of big/metropolitan cities is polluted most by	A. Automobile exhausts B. Pesticide residue C. Household waste D. Radio-active fall out
10	An electrolyte	A. Forms complex ions in solution B. Gives ions only when electricity is passed C. Possesses ions even in solid state D. Gives ions only when dissolved in water
11	Which one of the following compounds does not exist?	A. NCl_5 B. AsF_5 C. $SbCl_5$ D. PF_5
12	For the reaction $2A(g) + B(g) \rightleftharpoons 3C(g) + D(g)$ two moles each of A and B were taken into a flask The following must always be true when the system attained equilibrium	A. $[A] = [B]$ B. $[A] \text{ \&t; } [B]$ C. $[B] = [C]$ D. $[A] \text{ \&t; } [B]$
13	The weight of 11.2 liters of CO_2 at S.T.P. would be	A. 88 g B. 44 g C. 32 g

D. 22 g

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| 14 | The rotation of two carbon atoms joined by double bond would happen only if | A. Pi bond is broken
B. Sigma bond is broken
C. Both bonds are broken
D. None of above |
| 15 | Which of the substances Na, Hg, S, Pt and graphite can be used as electrodes in electrolytic cells having aqueous solution? | A. Na, Pt and graphite
B. Na and Hg
C. Pt and graphite only
D. Na and S only |
| 16 | The percentage of oxygen in NaOH is | A. 40
B. 60
C. 8
D. 10 |
| 17 | Which of the following is acidic? | A. SO_3
B. N_2O
C. BeO
D. HgO |
| 18 | The movement of solvent molecules through a semipermeable membrane is called | A. Electrolysis
B. Electrophoresis
C. Osmosis
D. Cataphoresis |
| 19 | Benzene + Ozone \rightarrow Y. In this sequence Y is | A. Benzene monoozonide
B. Benzene diozonide
C. Benzene triozonide
D. Succinic acid |
| 20 | Bleaching action of bleaching powder is due to the liberation of | A. O_2
B. OCl^-
C. Cl_2
D. Cl^- |