

## Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	In the manufacture of iron from haematite, limestone is added to act as.	A. Flux B. A reducing C. Slag D. An oxidizing agent.
2	Magnesium keeps on burning in	A. $N_{sub2}2$ B. $CO_{sub2}$ C. $N_{sub2}2O$ D. $N_{sub2}$ as well as $CO_{sub2}$
3	A cell constant is generally found by measuring the conductivity of aqueous solution of	A. $BaCl_{sub2}$ B. $KCl$ C. $NaCl$ D. $MgCl_{sub2}$
4	According to MO Theory the species $O_2^+$ possesses	A. Bond order of 2.5 B. Three unpaired electrons C. Diamagnetic character D. Stability lower than $O_{sub2}$
5	Portland cement is manufactured by using	A. Limestone, clay and sand B. Limestone, gypsum and sand C. Limestone, gypsum and alumina D. Limestone, clay and gypsum
6	An electrolyte	A. Forms complex ions in solution B. Gives ions only when electricity is passed C. Possesses ions even in solid state D. Gives ions only when dissolved in water
7	Which one of the following has the lowest boiling point?	A. <span style="font-size: 14.4444465637207px;">B</span> B. <span style="font-size: 14.4444465637207px;">Al</span> C. <span style="font-size: 14.4444465637207px;">Ga</span> D. <span style="font-size: 14.4444465637207px;">Ti</span>
8	The credit of discovering neutron goes to	A. Rutherford B. Langmuir C. Chadwick D. Austen
9	Which of the following imparts violet colouration to the non-luminous flame of Bunsen burner?	A. $NaCl$ B. $BaCl_{sub2}$ C. $CaCl_{sub2}$ D. $KCl$
10	How many kinds of space lattices are possible in a crystal?	A. 23 B. 7 C. 230 D. 14
11	With increasing principle quantum number the energy difference between adjacent energy levels in H atom	A. Decreases B. Increases C. Remains constant D. Decreases for low value of Z and increases for higher value of Z.
12	Which of the following halogens does not form its oxyacids?	A. Fluorine B. Chlorine C. Bromine D. Iodine
13	Which of the following represents elements in order of increasing atomic size?	A. $I, Br, Cl$ B. $Na, Mg, C$ C. $C, N, O$ D. $Li, Na, K$
14	Which of the following cannot be produced by acidic dehydration of alcohols?	A. Ethers B. Aldehydes C. Ketones D. Carboxylic acids

		C. Alkyl Hydrogen sulphate D. Alkene
15	The total number of protons in 10 g of calcium carbonate is ( $N_0 = 6.023 \times 10^{23}$ )	A. $1.5057 \times 10^{24}$ B. <span style="font-size: 14.4444465637207px;">2.0478 <math>\times 10^{24}</math></span> C. <span style="font-size: 14.4444465637207px;">3.0115 <math>\times 10^{24}</math></span> D. <span style="font-size: 14.4444465637207px;">4.0956 <math>\times 10^{24}</math></span>
16	The carbon atoms in calcium carbide are held by	A. Ionic bonds B. 2 sigma bonds C. 2 covalent one co-ordinate bond D. $\pi$ and $\sigma$ bond
17	Which of the following belongs to the halogen family?	A. Francium B. Polonium C. Radium D. Astatine
18	Gooch crucible used to filter the solution of	A. $\text{H}_2\text{SO}_4$ B. HCl C. $\text{KMnO}_4$ D. Both B & C
19	When electrons revolve in stationary orbits	A. There is no change in energy level B. They become stationary C. They are gaining kinetic energy D. There is increase in energy
20	The last orbit of argon would have electrons	A. 8 B. 18 C. 2 D. 6