

Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	Which is not a colligative property?	A. Osmotic pressure B. Lowering of vapour pressure C. Depression of freezing point D. Elevation of boiling point
2	In cold countries ethylene glycol is added to water in radiators of cars during winter. It results in	A. Lowering in b.pt. B. Reducing viscosity C. Reducing specific heat D. Lowering in freezing pt.
3	Among alkali metal salts, the lithium salts are the poorest conductors of electricity in aqueous solution because of	A. Easy diffusion of Li^{+} ions B. Lower ability of Li^{+} ions to polarize water molecules C. Lowest charge to radius ratio D. Higher degree of hydration of Li^{+} ions.
4	Wt. of 112 ml of oxygen at NTP on liquefaction would be	A. 0.32 g B. 0.64 g C. 0.16 g D. 0.96 g
5	The following has zero valency	A. Na B. Be C. Al D. Kr
6	The essential component of organic compound is	A. O B. C C. P D. N
7	A certain liberate 0.5 g of hydrogen in 2 h. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphate solution?	A. 12.7 gm B. 15.9 gm C. 31.8 gm D. 63.5 gm
8	Setting of cement is an	A. Exothermic reaction B. Endothermic reaction C. Neither exothermic nor endothermic D. None
9	1 mole of CH_4 contains	A. 6.02×10^{23} atoms of H B. 4 g-atom of hydrogen C. 1.81×10^{23} molecules of CH_4 D. 3.0 g of carbon
10	Which of the following belongs to the halogen family?	A. Francium B. Polonium C. Radium D. Astatine
11	Sulphur dioxide affects	A. Cell wall B. Plasmodesmata C. All membrane systems D. Nucleus
12	Which of the following with aqueous KOH will give acetaldehyde?	A. 1,2-Dichloroethane B. 1, 1-Dichloroethane C. Chloroacetic acid D. Ethyl chloride
13	A current of 9.65 ampere flowing for 10 minutes deposits 3.0 g of the metal which is monovalent. The atomic mass of the metal is	A. 10 B. 50 C. 30 D. 96.5
14	Octane number can be changed by	A. Isomerisation B. Alkylation C. Cyclisation D. All of these

15	The freezing point of 1 molal NaCl solution assuming NaCl to be 100% dissociated in water in	A. -1.86°C B. -3.72°C C. +1.86°C D. +3.72°C
16	At 500 K the equilibrium constant for reaction $\text{cis-C}_2\text{H}_2\text{Cl}_2 \rightleftharpoons \text{trans-C}_2\text{H}_2\text{Cl}_2$ is 0.6. At the same temperature the equilibrium constant for the reaction $\text{trans-C}_2\text{H}_2\text{Cl}_2 \rightleftharpoons \text{cis-C}_2\text{H}_2\text{Cl}_2$ will be	A. 0.60 B. 1.67 C. 0.66 D. 2.6
17	1-Chloropropane has two isomers It is an example of	A. Chain isomerism B. Position isomerism C. Functional group isomerism D. Metamerism
18	Which one is the property of an ideal solvent	A. Should be expensive B. It should react chemically with the solute C. Impurities should crystallize along with the solute D. Should be safe to use
19	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
20	In Friedal-Craft's alkylation besides AlCl_3 the other reactants are	A. $\text{C}_6\text{H}_6 + \text{NH}_3$ B. $\text{C}_6\text{H}_6 + \text{CH}_4$ C. $\text{C}_6\text{H}_6 + \text{CH}_3\text{Cl}$ D. $\text{C}_6\text{H}_6 + \text{CH}_3\text{COCl}$