

Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	The total number of protons in 10 g of calcium carbonate is ($N_0 = 6.023 \times 10^{23}$)	A. 1.5057×10^{24} B. 2.0478×10^{24} C. 3.0115×10^{24} D. 4.0956×10^{24}
2	Evaporation of water is	A. An exothermic change B. An endothermic change C. A process where no heat changes occur D. A process accompanied by chemical
3	Formic acid is obtained when	A. Calcium acetate is heated with conc. H_2SO_4 B. Calcium formate is heated with calcium acetate C. Glycerol is heated with oxalic acid D. Acetaldehyde is oxidized with K_2CrO_7 and H_2SO_4
4	The formula of nitre is	A. KNO_3 B. $NaNO_3$ C. $NaCl$ D. Na_2CO_3
5	The bond angle H - O - H in ice is closest to	A. 120° , 28° B. 60° C. 90° D. 109°
6	The formula of calcium cyanamide is	A. $Ca(CN)_2$ B. CaC_2N C. $CaNCN$ D. $CaCHNH_2$
7	Which of the following fluorides of xenon is impossible?	A. XeF_2 B. XeF_3 C. XeF_4 D. XeF_6
8	When electrons revolve in stationary orbits	A. There is no change in energy level B. They become stationary C. They are gaining kinetic energy D. There is increase in energy
9	The pore size of the filter paper depends upon	A. Nature of the medium B. Temperature of the medium C. Size of the particles D. Mass of the particles
10	The last orbit of argon would have electrons	A. 8 B. 18 C. 2 D. 6
11	Which of the following metal exhibits more than one oxidation?	A. Na B. Mg C. Fe D. Al
12	Carbon atom holding halogen in aryl halides is	A. sp^2 -hybridised B. sp^3 C. sp D. sp^2

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hybridised
D. <span style="font-size:
14.44444465637207px;">Sp
³</sup><span style="font-
size: 14.44444465637207px;">d -
hybridised

13 $\Delta H_{\text{Neutralisation}}$ is always

- A. Positive
- B. Negative
- C. Zero
- D. Positive or negative

14 In cold countries ethylene glycol is added to water in radiators of cars during winter. It results in

- A. Lowering in b.pt.
- B. Reducing viscosity
- C. Reducing specific heat
- D. Lowering in freezing pt.

15 Bragg's law is given by equation

- A. $n \lambda \sin \theta = d \sin \theta$
- B. $n \lambda = 2 d \sin \theta$
- C. $2n\lambda = d \sin \theta$
- D. $n\lambda = 1/2 d \sin \theta$

16 Which of the following has the greatest tendency to lose an electron?

- A. F
- B. Fr
- C. S
- D. Be

17 The equilibrium $\text{Cr}_2\text{O}_7^{2-} \rightleftharpoons 2\text{CrO}_4^{2-}$ is shifted to the right in

- A. An acidic medium
- B. A basic medium
- C. A neutral medium
- D. It does not exist

18 Which of the following alcohols cannot be produced by the treatment of aldehydes or ketones with NaBH_4 or LiAlH_4 ?

- A. 1-Propanol
- B. 2-Propanol
- C. 2-Methyl-2-propanol
- D. Ethanol

19 Ionic solids with defects contain

- A. Equal number of cation and anion vacancies
- B. Interstitial anions and anion vacancies
- C. Cation vacancies only
- D. Cation vacancies and interstitial cations

20 The rate at which a substance reacts depends on its

- A. Atomic weight
- B. Equivalent weight
- C. Molecular weight
- D. Active mass