

Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	Which one is the property of an ideal solvent	<p>A. Should be expensive B. It should react chemically with the solute C. Impurities should crystallize along with the solute D. Should be safe to use</p>
2	The equilibrium constant in a reversible chemical reaction at a given temperature	<p>A. Depends on the initial concentration of the reactants B. Depends on the concentration of one of the products at equilibrium C. Does not depend on the initial concentrations of reactants D. Is not characteristic of the reaction</p>
3	The mass of the neutron is of the order of	<p>A. 10^{-23} kg B. 10^{-24} kg C. 10^{-26} kg D. 10^{-27} kg</p>
4	Ammonia gas used directly as a fertilizer is injected into the soil at a depth of about	<p>A. Two inches B. Three inches C. Five inches D. Six inches</p>
5	The symbol of the element whose atoms have the outer most electronic configuration $2s^2 2p^3$ is	<p>A. N B. Li C. P D. Na</p>
6	The molal elevation constant is the ratio of the elevation in boiling point to	<p>A. Molarity B. Molality C. Mole fraction of solute D. Mole fraction of solvent</p>
7	Hybridization explain the----- of orbitals	<p>A. Type of Bonding B. Shapes C. Shape and Type of bonding D. None of above</p>
8	The movement of solvent molecules through a semipermeable membrane is called	<p>A. Electrolysis B. Electrophoresis C. Osmosis D. Cataphoresis</p>
9	Isopropyl alcohol on oxidation forms	<p>A. Acetone B. Ether C. Ethylene D. Acetaldehyde</p>

10	The valence orbital configuration of an element with atomic number 23 is	<p> $1s^2 2s^2 2p^6 3s^2 3p^4$ $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$ $1s^2 2s^2 2p^6 3s^2 3p^4 3d^1$ $1s^2 2s^2 2p^6 3s^2 3p^4 3d^1 4s^1$ </p>
11	Potassium crystallizes with a	<p> A. Orthogonal lattice B. Cubic lattice C. Triclinic D. Ortho rhombic lattice </p>
12	Tollen's reagent is	<p> A. Ammonical cuprous chloride B. Ammonical cuprous oxide C. Ammonical silver bromide D. Ammonical silver nitrate </p>
13	Which of the following metal exhibits more than one oxidation?	<p> A. Na B. Mg C. Fe D. Al </p>
14	1-Chloropropane has two isomers It is an example of	<p> A. Chain isomerism B. Position isomerism C. Functional group isomerism D. Metamerism </p>
15	In the equilibrium $N_2 + 3H_2 \rightleftharpoons 2NH_3 + 22 \text{ kcal}$ the formation of ammonia is favoured by	<p> A. Increasing the pressure B. Increasing the temperature C. Decreasing the pressure D. Adding ammonia </p>
16	Atmosphere of big/metropolitan cities is polluted most by	<p> A. Automobile exhausts B. Pesticide residue C. Household waste D. Radio-active fall out </p>
17	A clathrate may be defined as a	<p> A. Cage compound B. Liquid crystal C. Mixture D. Solid solution </p>
18	In a reversible chemical reaction having two reactants in equilibrium if the concentration of the reactants are doubled then the equilibrium constant will	<p> A. Also be doubled B. Be halved C. Becomes one fourth D. Remains the same </p>
19	A zero order reaction is one whose rate is independent of	<p> A. Temperature of the reaction B. The concentration of the reactants C. The concentration of the products D. The material of the vessel in which the reaction is carried out </p>
20	Bell metal is an alloy of	<p> A. Cu, Zn, and Sn B. Cu, Zn and Ni C. Cu and Zn D. Cu and Sn </p>