

## Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	The unit of rate constant for a zero order reaction is	<p>A. Liter sec<sup>-1</sup></p> <p>B. Liter 14.4444465637207px;"&gt;mol&lt;/span&gt; &lt;sup&gt;-1&lt;/sup&gt;&lt;span style="font-size: 14.4444465637207px;"&gt;sec&lt;/span&gt; &lt;sup&gt;-1&lt;/sup&gt;</p> <p>C. Mol liter<sup>-1</sup>&amp;nbsp;&lt;/sup&gt;&lt;span style="font-size: 14.4444465637207px;"&gt;sec&lt;/span&gt; &lt;sup&gt;-1&lt;/sup&gt;</p> <p>D. Mol&amp;nbsp;&lt;span style="font-size: 14.4444465637207px;"&gt;sec&lt;/span&gt; &lt;sup&gt;-1&lt;/sup&gt;</p>
2	The rate of a reaction that does not involve gases does not depend upon	<p>A. Pressure</p> <p>B. Temperature</p> <p>C. Concentration</p> <p>D. Catalyst</p>
3	Which of the following will not change the concentration of ammonia in the equilibrium $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ ; $\Delta H = -kj$	<p>A. Increase of pressure</p> <p>B. Increase of temperature</p> <p>C. Decrease of volume</p> <p>D. Addition of catalyst</p>
4	Which of the following elements is most electronegative?	<p>A. Oxygen</p> <p>B. Chlorine</p> <p>C. Nitrogen</p> <p>D. Fluorine</p>
5	The alkali metal which is liquid at 15°C is	<p>A. K</p> <p>B. Cs</p> <p>C. Na</p> <p>D. None</p>
6	Crystal can be classified in to basic crystal habits	<p>A. 7</p> <p>B. 4</p> <p>C. 14</p> <p>D. 3</p>
7	Which of the following value of $\Delta H_f^\circ$ represent that the product is least stable?	<p>A. -94.0 kcal mol<sup>-1</sup></p> <p>B. -231.6&amp;nbsp;&lt;span style="font-size: 14.4444465637207px;"&gt;kcal mol&lt;/span&gt;&lt;sup&gt;-1&lt;/sup&gt;</p> <p>C. +21.4&amp;nbsp;&lt;span style="font-size: 14.4444465637207px;"&gt;kcal mol&lt;/span&gt;&lt;sup&gt;-1&lt;/sup&gt;</p> <p>D. +64.8&amp;nbsp;&lt;span style="font-size: 14.4444465637207px;"&gt;kcal mol&lt;/span&gt;&lt;sup&gt;-1&lt;/sup&gt;</p>
8	The molal elevation constant is the ratio of the elevation in boiling point to	<p>A. Molarity</p> <p>B. Molality</p> <p>C. Mole fraction of solute</p> <p>D. Mole fraction of solvent</p>
9	In the equilibrium $N_2 + 3H_2 \rightleftharpoons 2NH_3 + 22 \text{ kcal}$ the formation of ammonia is favoured by	<p>A. Increasing the pressure</p> <p>B. Increasing the temperature</p> <p>C. Decreasing the pressure</p> <p>D. Adding ammonia</p>
10	Which one is the property of an ideal solvent	<p>A. Should be expensive</p> <p>B. It should react chemically with the solute</p> <p>C. Impurities should crystallize along with the solute</p> <p>D. Should be safe to use</p>
11	A certain liberate 0.5 g of hydrogen in 2 h. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphare solution?	<p>A. 12.7 gm</p> <p>B. 15.9 gm</p> <p>C. 31.8 gm</p> <p>D. 63.5 gm</p>
12	Which of the following has greatest tendency to lose electron?	<p>A. F</p> <p>B. Fr</p>

13	Alum is not used	<p>A. As a mordant in dyeing  <b>B. As an insecticide</b>            C. In purification of water            D. In tanning of leather</p>
14	Which of the following statement regarding catalyst is not true?	<p>A. A catalyst remains unchanged in composition and quantity at the end of the reaction  <b>B. A catalyst can initiate a reaction</b>            C. A catalyst dose not alter the equilibrium in a reversible reaction            D. Catalysts are sometimes very specific in respect of reaction</p>
15	Vital force theory was rejected by	<p>A. Berzelius            B. Kolbe  <b>C. Wholer</b>            D. Lavoiser</p>
16	Atmosphere of big/metropolitan cities is polluted most by	<p><b>A. Automobile exhausts</b>            B. Pesticide residue            C. Household waste            D. Radio-active fall out</p>
17	Cannizzaro reaction is not given by	<p>A. Trimethyl acetaldehyde  <b>B. Acetaldehyde</b>            C. Benzaldehyde            D. Formaldehyde</p>
18	Inert pair effect plays an important role in case of	<p>A. F            B. Al            C. Si  <b>D. Tl</b></p>
19	The rate of a reaction can be increased in general by all the factors except by	<p>A. Using a catalyst            B. Increasing temperature  <b>C. Increasing the activation energy</b>            D. Increasing the conc. of reactants</p>
20	Red lead is	<p>A. PbO  <b>B. Pb<sub>3</sub>O<sub>4</sub></b>            C. <span style="font-size: 14.4444465637207px;">Pb</span>  <span style="font-size: 14.4444465637207px;">O</span>  <span style="font-size: 14.4444465637207px;">Pb</span>  <span style="font-size: 14.4444465637207px;">O</span>            D. <span style="font-size: 14.4444465637207px;">Pb</span>  <span style="font-size: 14.4444465637207px;">O</span>  <span style="font-size: 14.4444465637207px;">Pb</span>  <span style="font-size: 14.4444465637207px;">O</span></p>