

## Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	The symbol of the element whose atoms have the outer most electronic configuration $2s^2 2p^3$ is	A. N B. Li C. P D. Na
2	Which of the following elements is most electronegative?	A. Oxygen B. Chlorine C. Nitrogen D. Fluorine
3	In cold countries ethylene glycol is added to water in radiators of cars during winter. It results in	A. Lowering in b.pt. B. Reducing viscosity C. Reducing specific heat D. Lowering in freezing pt.
4	Acetic anhydride is obtained from acetyl chloride by the reaction of	A. $\text{P}_2\text{O}_5$ B. $\text{HCl}$ C. $\text{CH}_3\text{COONa}$ D. $\text{CH}_3\text{COCH}_3$
5	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
6	In a reversible chemical reaction having two reactants in equilibrium, if the concentration of the reactants are doubled, then the equilibrium constant will	A. Also be doubled B. Be halved C. Become one fourth D. Remain the same
7	Benzene + Ozone $\rightarrow$ Y. In this sequence, Y is	A. Benzene monoozonide B. Benzene diozonide C. Benzene triozone D. Succinic acid
8	The order of reactivity of halogens in aliphatic substitution reactions is	A. $\text{Br}_2 > \text{Cl}_2 > \text{F}_2$ B. $\text{Cl}_2 > \text{Br}_2 > \text{F}_2$ C. $\text{Cl}_2 > \text{Br}_2$ D. $\text{F}_2 > \text{Br}_2 > \text{Cl}_2$
9	Which of the following minerals does not contain Al?	A. Cryolite B. Mica C. Feldspar D. Fluorspar
10	Which of the following is different from the other three oxides?	A. $\text{MgO}$ B. $\text{SnO}$ C. $\text{ZnO}$ D. $\text{Cr}_2\text{O}_3$
11	Which is the most volatile compound?	A. HI B. HCl C. HBr D. HF
		A. Ice bath

12	If the solvent is inflammable for heating purpose we use	B. Water bath C. Wire gauze D. Thermostat
13	At 500 K the equilibrium constant for reaction $\text{cis-C}_2\text{H}_2\text{Cl}_2 \rightleftharpoons \text{trans-C}_2\text{H}_2\text{Cl}_2$ is 0.6. At the same temperature the equilibrium constant for the reaction $\text{trans-C}_2\text{H}_2\text{Cl}_2 \rightleftharpoons \text{cis-C}_2\text{H}_2\text{Cl}_2$ will be	A. 0.60 B. 1.67 C. 0.66 D. 2.6
14	For the reaction $2\text{A(g)} + \text{B(g)} \rightleftharpoons 3\text{C(g)} + \text{D(g)}$ two moles each of A and B were taken into a flask. The following must always be true when the system attained equilibrium	A. $[\text{A}] = [\text{B}]$ B. $[\text{A}] < [\text{B}]$ C. $[\text{B}] = [\text{C}]$ D. $[\text{A}] > [\text{B}]$
15	In the manufacture of iron from haematite, limestone is added to act as.	A. Flux B. A reducing C. Slag D. An oxidizing agent.
16	Chile salt petre is	A. $\text{NaNO}_3$ B. $\text{Na}_2\text{SO}_4$ C. $\text{KNO}_3$ D. $\text{Na}_2\text{S}_2\text{O}_3$
17	Which of the following cannot be produced by acidic dehydration of alcohols?	A. Ethers B. Aldehyde C. Alkyl Hydrogen sulphate D. Alkene
18	Which of the following has greatest tendency to lose electron?	A. F B. Fr C. S D. Be
19	Saturated solution of NaCl on heating becomes	A. Super saturated B. Unsaturated C. Remains saturated D. None
20	Which of the following belongs to the halogen family?	A. Francium B. Polonium C. Radium D. Astatine