

Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	Acetic anhydride is obtained from acetyl chloride by the reaction of	<p>A. P</p> <p>B. H</p> <p>C. CH_3COONa</p> <p>D. COCH_3</p>
2	Benzene is obtained by fractional distillation of	<p>A. Heavy oil</p> <p>B. Anthracene oil</p> <p>C. Middle oil</p> <p>D. Light oil</p>
3	Hybridization explain the----- of orbitals	<p>A. Type of Bonding</p> <p>B. Shapes</p> <p>C. Shape and Type of bonding</p> <p>D. None of above</p>
4	Heating a mixture of sodium benzoate and soda lime gives	<p>A. Benzene</p> <p>B. Methane</p> <p>C. Sodium benzoate</p> <p>D. Calcium benzoate</p>
5	Which one of the following has the lowest boiling point?	<p>A. B</p> <p>B. Al</p> <p>C. Ga</p> <p>D. Ti</p>
6	When KClO_3 is heated it decomposes into KCl and O_2 if some MnO_2 is added the reaction goes much faster because	<p>A. MnO_2 decomposes to give O_2</p> <p>B. MnO_2 provides heat by reacting</p> <p>C. Better contact is provided by MnO_2</p> <p>D. MnO_2 acts as a catalyst</p>
7	With increasing principle quantum number the energy difference between adjacent energy levels in H atom	<p>A. Decreases</p> <p>B. Increases</p> <p>C. Remains constant</p> <p>D. Decreases for low value of Z and increases for higher value of Z.</p>
8	Number of elements presents in the fifth period of periodic table is	<p>A. 8</p> <p>B. 10</p> <p>C. 18</p> <p>D. 32</p>
9	Which of the following has greatest reducing power?	<p>A. Hl</p> <p>B. Hbr</p> <p>C. HCl</p> <p>D. HI</p>
10	Which of the following cannot be produced by acidic dehydration of alcohols?	<p>A. Ethers</p> <p>B. Aldehyde</p> <p>C. Alkyl Hydrogen sulphate</p> <p>D. Alkene</p>

A. Calcium acetate is heated with conc. H_2SO_4

B. Calcium formate is heated with calcium

11	Formic acid is obtained when	<p>B. Calcium formate is heated with calcium acetate</p> <p>C. Glycerol is heated with oxalic acid</p> <p>D. Acetaldehyde is oxidized with K_2CrO_7 and H_2SO_4</p>
12	Vital force theory was rejected by	<p>A. Berzelius</p> <p>B. Kolbe</p> <p>C. Wholer</p> <p>D. Lavoiser</p>
13	A certain liberate 0.5 g of hydrogen in 2 h. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphate solution?	<p>A. 12.7 gm</p> <p>B. 15.9 gm</p> <p>C. 31.8 gm</p> <p>D. 63.5 gm</p>
14	The halide which is not hydrolysed is	<p>A. $SiCl_4$</p> <p>B. SiF_4</p> <p>C. CCl_4</p> <p>D. $PbCl_4$</p>
15	Inter molecular forces in solid hydrogen are	<p>A. Covalent forces</p> <p>B. Van der Waal forces or London dispersion force</p> <p>C. Hydrogen bonds</p> <p>D. All of these</p>
16	Water (H_2O) is liquid while hydrogen sulphide (H_2S) is a gas because	<p>A. Water has higher molecular weight</p> <p>B. Hydrogen sulphide is a weak acid</p> <p>C. Sulphur has high electronegativity than oxygen</p> <p>D. Water molecules associate through hydrogen bonding.</p>
17	The dimension of rate constant of a second order reaction involves	<p>A. Neither time nor concentration</p> <p>B. Only time</p> <p>C. Time and concentration</p> <p>D. Time and square of concentration</p>
18	The freezing point of 1 molal NaCl solution assuming NaCl to be 100% dissociated in water is	<p>A. $-1.86^\circ C$</p> <p>B. $-3.72^\circ C$</p> <p>C. $+1.86^\circ C$</p> <p>D. $+3.72^\circ C$</p>
19	The pore size of the filter paper depends upon	<p>A. Nature of the medium</p> <p>B. Temperature of the medium</p> <p>C. Size of the particles</p> <p>D. Mass of the particles</p>
20	Which of the following has greatest tendency to lose electron?	<p>A. F</p> <p>B. Fr</p> <p>C. S</p> <p>D. Be</p>