

Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	Which of the following is an example of body centred cube?	A. Magnesium B. Zinc C. Copper D. Sodium
2	The rate of a reaction can be increased in general by all the factors except by	A. Using a catalyst B. Increasing temperature C. Increasing the activation energy D. Increasing the conc. of reactants
3	What quantity of limestone (CaCO_3) on heating will give 56 kg of CaO ?	A. 1000 kg B. 56 kg C. 44 kg D. 100kg
4	A chemical reaction A B is said to be in equilibrium when	A. Complete conversion of A to B has taken place B. Conversion of A to B is only 50% complete C. Only 10% conversion of A to B has taken place D. The rate of transformation of A to B is just equal to rate of transformation of B to A in the system
5	The relative rates of diffusion of a gas (Mol. wt. - 98) as compared to hydrogen will be	A. $\frac{1}{7}$ B. $\frac{1}{5}$ C. $\frac{1}{4}$ D. 1
6	Which of the following statement is correct if the intermolecular forces in liquids A, B and C are in the order $A < B < C$?	A. B evaporates more readily than A B. B evaporates less readily than C C. A and B evaporates at the same rate D. A evaporates more readily than C
7	Which of the following transition metal ions will have definite value of magnetic moment?	A. Sc^{3+} B. Ti^{3+} C. Cu^{+} D. Zn^{2+}
8	Among alkali metal salts, the lithium salts are the poorest conductors of electricity in aqueous solution because of	A. Easy diffusion of Li^{+} ions B. Lower ability of Li^{+} ions to polarize water molecules C. Lowest charge to radius ratio D. Higher degree of hydration of Li^{+} ions.
9	The disaccharide present in milk is	A. Sucrose B. Maltose C. Lactose D. Cellobiose
10	Inter molecular forces in solid hydrogen are	A. Covalent forces B. Van der Waal forces or London dispersion force C. Hydrogen bonds D. All of these
11	Maximum number of active hydrogens are present in	A. Acetic-acid B. Glycerol C. Methane D. Methanol
12	SO_2 and NO_2 pollution by increasing	A. Alkalinity B. Acidity C. ...

		C. Neutrality D. Buffer action
13	The alkali metal which is liquid at 15°C is	A. K B. Cs C. Na D. None
14	In crystal structure of sodium chloride the arrangement of Cl ⁻ ions is	A. Fee B. Both fee and bcc C. Bee D. None of these
15	When electrons revolve in stationary orbits	A. There is no change in energy level B. They become stationary C. They are gaining kinetic energy D. There is increase in energy
16	Magnesium keeps on burning in	A. N ₂ B. CO ₂ C. N ₂ O D. N ₂ as well as CO ₂
17	Mark the smallest atom	A. F B. Cl C. Br D. I
18	Which of the following process is used to separate insoluble particles from liquids?	A. Separation B. Filtration C. Crystallization D. Condensation
19	Which of the following statements is most appropriate about effective nuclear charge? It depends upon	A. The shielding constant B. The atomic number C. The charge on the nucleus D. Both the nuclear charge and the shielding constant
20	Toluene can be oxidized to benzoic acid by	A. KMnO ₄ (alk) B. K ₂ Cr ₂ O ₇ (acidic) C. Both D. None