

Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	The ratio of close packed atoms to tetrahedral holes in cubic close packing is	A. 1 : 1 B. 1 : 2 C. 1 : 3 D. 2 : 1
2	Which one of the following elements occurs free in nature?	A. N_2 B. P C. As D. Sb
3	Octane number can be changed by	A. Isomerisation B. Alkylation C. Cyclisation D. All of these
4	Which of the substances Na, Hg, S Pt and graphic can be used as electrodes in electrolytic cells having aqueous solution?	A. Na, Pt and graphite B. Na and Hg C. Pt and graphite only D. Na and S only
5	The heats evolved in combustion of rhombic and monoclinic sulphur are - 70960 and - 71030 cal mol ⁻¹ respectively what will be heat of conversion of rhombic sulphur to monoclinic?	A. 70960 calories B. 71030 calories C. -70 calories D. +70 calories
6	The disaccharide present in milk is	A. Sucrose B. Maltose C. Lactose D. Cellobiose
7	At 500 K the equilibrium constant for reaction $\text{cis-C}_2\text{H}_2\text{Cl}_2 \rightleftharpoons \text{trans-C}_2\text{H}_2\text{Cl}_2$ is 0.6. At the same temperature the equilibrium constant for the reaction $\text{trans-C}_2\text{H}_2\text{Cl}_2 \rightleftharpoons \text{cis-C}_2\text{H}_2\text{Cl}_2$ will be	A. 0.60 B. 1.67 C. 0.66 D. 2.6
8	The bond angle H - O - H in ice is closest to	A. 120° , 28° B. 60° C. 90° D. 109°
9	The mass of the neutron is of the order of	A. 10 ⁻²³ kg B. 1.44444465637207 × 10 ⁻²⁴ kg C. 1.44444465637207 × 10 ⁻²⁶ kg D. 1.44444465637207 × 10 ⁻²⁷ kg
10	The credit of discovering neutron goes to	A. Rutherford B. Langmuir C. Chadwick D. Austen
11	Red lead is	A. PbO B. Pb ₃ O ₄ C. Pb ₂ O ₃ D. Pb ₄ O ₃

12	Which one is primary alcohol?	<p>B. Propan-2-ol</p> <p>C. Butane-1-ol</p> <p>D. 2,3-Dimethylhexane-4-ol</p>
13	Which of the following will not change the concentration of ammonia in the equilibrium $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$: $\Delta H = -kJ$	<p>A. Increase of pressure</p> <p>B. Increase of temperature</p> <p>C. Decrease of volume</p> <p>D. Addition of catalyst</p>
14	Calcium acetate when dry distilled gives	<p>A. Formaldehyde</p> <p>B. Acetaldehyde</p> <p>C. Acetone</p> <p>D. Acetic anhydride</p>
15	Which of the following statements is most appropriate about effective nuclear charge? It depends upon	<p>A. The shielding constant</p> <p>B. The atomic number</p> <p>C. The charge on the nucleus</p> <p>D. Both the nuclear charge and the shielding constant</p>
16	When electricity is passed through molten $Al_2O_3 + Na_3AlF_6$ and 13.5 gms Al are deposited, the number of faraday must be	<p>A. 0.5</p> <p>B. 1.0</p> <p>C. 1.5</p> <p>D. 2.0</p>
17	The structure of XeF_6	<p>A. Distorted octahedral</p> <p>B. Pyramidal</p> <p>C. Tetrahedral</p> <p>D. None of the above</p>
18	Which of the following process is used to separate insoluble particles from liquids?	<p>A. Separation</p> <p>B. Filtration</p> <p>C. Crystallization</p> <p>D. Condensation</p>
19	Which of the following fluorides does not exist?	<p>A. NF_5</p> <p>B. PF_5</p> <p>C. AsF_5</p> <p>D. SbF_5</p>
20	Inter molecular forces in solid hydrogen are	<p>A. Covalent forces</p> <p>B. Van der Waal forces or London dispersion force</p> <p>C. Hydrogen bonds</p> <p>D. All of these</p>