

## Chemistry General Science Test Hard Mode

Sr	Questions	Answers Choice
1	Among the alkaline earth metals the element forming predominantly covalent compounds is	A. Be B. Mg C. Sr D. Calcium
2	Setting of plaster of paris involves	A. Oxidation with atmospheric oxygen B. Combination with atmosphere CO <sub>2</sub> C. Dehydration D. Hydration to yield another hydrate.
3	Which of the following elements is most electronegative?	A. Oxygen B. Chlorine C. Nitrogen D. Fluorine
4	By using the fluted filter paper rate of filtration is	A. Increased B. Decreased C. Filtration is constant D. Having no effect
5	Which one of the following has the lowest boiling point?	A. <span style="font-size: 14.4444465637207px;">B</span> B. <span style="font-size: 14.4444465637207px;">Al</span> C. <span style="font-size: 14.4444465637207px;">Ga</span> D. <span style="font-size: 14.4444465637207px;">Ti</span>
6	For the carbylamine reaction we need hot alc.KOH and	A. Any amine and chloroform B. Chloroform and Ag powder C. A primary amine and chloroform D. A mono alkyl amine and trichloromethane
7	Which one of the following elements occurs free in nature?	A. N B. P C. As D. Sb
8	The maximum number of electrons in a subshell for which $l = 3$ is	A. 14 B. 10 C. 8 D. 4
9	If $N_A$ is Avogadro's number then number of valence electrons in 4.2 g of nitride ions $N^{3-}$ is	A. 2.4 $N_A$ B. 4.2&nbsp; <span style="font-size: 14.4444465637207px;">N</span> <sub>A</sub> C. 1.6&nbsp; <span style="font-size: 14.4444465637207px;">N</span> <sub>A</sub> D. 3.2&nbsp; <span style="font-size: 14.4444465637207px;">N</span> <sub>A</sub>
10	Benzene + Ozone $\rightarrow$ Y. in this sequence Y is	A. Benzene monoozonide B. Benzene diozonide C. Benzene triozonide D. Succinic acid
11	Fertilizer are made by	A. Nature only B. Artificial methods only C. Both artificial and natural methods D. None of the above
12	For the reaction 2A(g) + B(g) 3C(g) + D(g) two moles each of A and B were taken into a flask. The following must always be true when the system attained equilibrium	A. [A] = [B] B. [A] &lt; [B] C. [B] = [C] D. [A] &gt; [B]
		A. F B. Cl

13	Mark the smallest atom	C. Br D. I
14	Which one is not usually used for the crystallization	A. Acetone B. Acetic acid C. <b>Sulphuric acid</b> D. Chloroform
15	Acetic anhydride is obtained from acetyl chloride by the reaction of	A. $\text{CH}_3\text{COCl} + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{COOH} + \text{HCl}$ B. $\text{CH}_3\text{COCl} + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{COOH} + \text{HCl}$ C. $\text{CH}_3\text{COCl} + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{COOH} + \text{HCl}$ D. $\text{CH}_3\text{COCl} + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{COOH} + \text{HCl}$
16	Which of the following is acidic?	A. $\text{SO}_3$ B. $\text{N}_2\text{O}$ C. $\text{BeO}$ D. $\text{HgO}$
17	The treatment of benzene with isobutene in the presence of sulphuric acid give	A. Isobutyl benzene B. Tert-Butyl benzene C. n-Butyl benzene D. no reaction
18	Which of the following halogens does not form its oxyacids?	A. Fluorine B. Chlorine C. Bromine D. Iodine
19	Octane number can be changed by	A. Isomerisation B. Alkylation C. Cyclisation D. All of these
20	With increasing principle quantum number the energy difference between adjacent energy levels in H atom	A. Decreases B. Increases C. Remains constant D. Decreases for low value of Z and increases for higher value of Z.