

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	John Newlands gave the idea of.	A. Law of triads B. Law of Octaves C. Modern Periodic law D. Curves between weight and volume
2	Which set of the following conditions results in higher ionization energy.	A. Smaller atom and greater nuclear charge B. Smaller atom and smaller nuclear charge C. Larger atom and greater nuclear charge D. Large atom and the smaller nuclear charge
3	Which dialyzing solution has high pH	A. Vinegar B. Lemon juice C. Ammonia solution D. Coffee
4	Which of the following carbocations would be the least stable.	A. $(\text{CH}_3)_2\text{CH}^+$ B. $\text{CH}_3\text{CH}_2^+$ C. $\text{CH}_3^+$ D. $(\text{CH}_3)_3\text{C}^+$
5	An acid with low $K_a$ value is.	A. Strong B. Weak C. Base D. Neutral
6	When a bond is formed	A. Energy is absorbed B. Energy is released C. $\Delta H$ is always zero D. No energy change
7	Which metal will displace Cu from $\text{CuSO}_4$ Solution.	A. Zn B. Ag C. Hg D. Pb
8	If 1 Faraday of electricity is passed the mass deposited equals.	A. 1 gram equivalent B. 1 gram C. 1 mole D. 1 atm
9	The enthalpy change for a reaction depends on.	A. Pathway taken from reactants to products B. Presence of a catalyst C. Initial and final states of the reactants and products D. Rate of the reaction
10	Structural isomers differ in	A. Only the boiling point B. Only the melting point C. Arrangement of atoms D. Number of atoms
11	Which hydrocarbon has only single bonds.	A. Ethene B. Ethyne C. Propane D. Benzene
12	When water freezes at $0^\circ\text{C}$ its density decreases due to.	A. Cubic structure of ice B. Empty spaces C. Decrease in volume D. Decrease in viscosity
13	What intermediate is formed during the electrophilic addition of $\text{HBr}$ to an alkene.	A. Carbonation B. Carbanion C. Radical D. Epoxide
		A. Resonance

14	What else is required on the electronegative atom for hydrogen bonding besides bonding to H.	<p>B. Lone pair of electrons</p> <p>C. Double bond</p> <p>D. Positive charge</p>
15	Which one is the correct statement among the following	<p>A. Anionic radius is generally smaller than atomic radius</p> <p>B. Cationic radius is generally bigger than atomic radius</p> <p>C. Cationic radius is generally smaller than atomic radius</p> <p>D. Both anionic and cationic radii are smaller than atomic radius</p>
16	The $K_a$ of a weak acid is $10^{-5}$ its $pK_a$ is	<p>A. 2</p> <p>B. 3</p> <p>C. 5</p> <p>D. 10</p>
17	The quantum number 'm' of a free gaseous atom is associated with	<p>A. The effective volume of the orbital</p> <p>B. The shape of the orbital</p> <p>C. The spatial orientation of the orbital</p> <p>D. The energy of the orbital in the absence of a magnetic field</p>
18	The strength of base is measured by its.	<p>A. <math>K_a</math></p> <p>B. <math>K_b</math></p> <p>C. <math>pK_w</math></p> <p>D. <math>K_w</math></p>
19	What is the prefix for a three-carbon chain in IUPAC nomenclature.	<p>A. Eth-</p> <p>B. Prop-</p> <p>C. But-</p> <p>D. Pent-</p>
20	The reducing agent in a redox reaction	<p>A. Gains electrons</p> <p>B. Loss electrons</p> <p>C. Accepts protons</p> <p>D. Donates protons</p>