

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	The pH scale ranges typically from	A. 1 - 10 B. 0 - 14 C. -1 to 1 D. 7 - 14
2	Which of the following is most stable.	A. Tertiary radical B. Primary radical C. Secondary radical D. Methyl radical
3	Reaction in Haber process is	A. Endothermic B. Exothermic C. Irreversible D. Neutral
4	Dipole moment of CO ₂ is	A. >1.8 D B. >1.94 D C. >1.0 D D. Zero
5	The sign of Delta H during melting is.	A. Negative B. Positive C. Zero D. Can't be predicted
6	Position of equilibrium is affected by	A. Temperature B. Catalyst C. Inert gas D. Surface area
7	The K _a of a weak acid is 10 ⁻⁵ its pK _a is	A. 2 B. 3 C. 5 D. 10
8	Which of the following statements is true about diffusion in liquids?	A. Liquids do not diffuse at all B. Liquids diffuse faster than gases C. Liquids diffuse slower than gases due to longer intermolecular forces D. Liquids diffuse only when heated
9	Which one is a polar molecule	A. HCl B. CO ₂ C. BF ₃ D. CCl ₄
10	Electrochemical cells convert chemical energy into	A. Electrical energy B. Heat C. Light D. Nuclear energy
11	Mass of moles of CO ₂ is.	A. 44 g B. 88 g C. 40 g D. 50 g
12	The solubility product of AgCl is 2.0 x 10 ⁻¹⁰ mol ² dm ⁻⁶ . The maximum concentration of Ag ⁺ ions in the solution is.	A. 2.0 x 10 ⁻¹⁰ mol dm ⁻³ B. 1.41 x 10 ⁻⁵ mol dm ⁻³ C. 1.0 x 10 ⁻¹⁰ mol dm ⁻³ D. 4.0 x 10 ⁻²⁰ mol dm ⁻³
13	Aromatic hydrocarbons must contain.	A. Straight chains B. Double bonds only C. A benzene ring D. A triple bond
		A. Ash former

14	The term "Halogen" mean	<p>B. <p>Salt former</p> C. <p>Coper giver</p> D. <p>Sulpher giver</p></p>
15	The quantum numebr values for the fourth electron of Be atom are.	<p>A. <p>1,0,0</p> B. <p>2,0,0</p> C. <p>2,1,0</p> D. <p>1,1,1</p></p>
16	Decreasing pressure shifts equilibrium towards.	<p>A. Sides with mroe gas moleules B. Side with fewer gas molecules C. Liquids D. Solids</p>
17	John Newlands gavt he idea of.	<p>A. <p>Law of triads</p> B. <p>Law of Octaves</p> C. <p>Modern Periodic law</p> D. <p>Curves between weight and volume</p></p>
18	Overall order is sum of	<p>A. Coefficients B. Exponents in rate law C. Moles D. Products</p>
19	1 Mole of OH-1 ion is equal to	<p>A. <p>18 g</p> B. <p>17 g</p> C. <p>16 g</p> D. <p>10 g</p></p>
20	The activity series of meals arranges metals in order of their.	<p>A. Atomic mass B. Density C. Ease of oxidation D. Ease of reduction</p>