

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which is a pair of geometrical isomers.	A. Ethane and propane B. But-1- ene and but-2- ene C. <b>eis-but-2- ene and trans -but-2- ene</b> D. Benzene and toluene
2	In a chemical reaction, a catalyst.	A. Alters Delta H B. <b>Lowers activation energy</b> C. Is consumed D. Forms product
3	pH of blood is around.	A. 6.0 B. <b>7.4</b> C. 5.4 D. 8.0
4	If we double the concentration of a reactant, the rate increases by four times, the reaction is.	A. First Order B. <b>Second Order</b> C. Third Order D. Zero order
5	Reaction rate by conductometry depends on	A. <b>Ionic conductivity</b> B. Pressure C. colour D. Temperature
6	Which of the following is NOT a property of liquid crystals.	A. <b>Fluidity</b> B. <b>Rigidity like crystals</b> C. <b>Complete isotropy</b> D. <b>Viscosity</b>
7	$K_p = K_c(RT)^{\Delta n}$ is used to relate.	A. <b><math>K_p</math> and <math>K_c</math></b> B. Temperature and pressure C. Energy and volume D. Gibbs free energy
8	Decreasing pressure shifts equilibrium towards.	A. <b>Sides with more gas molecules</b> B. Side with fewer gas molecules C. Liquids D. Solids
9	Which has the lowest pH?	A. 0.1 M $\text{CH}_3\text{COOH}$ B. <b>1 M HCl</b> C. 1 M $\text{NH}_3$ D. 0.1 M HCl
10	Which of the following has greater viscosity than water	A. <b>Glycerine</b> B. <b>Hexane</b> C. <b>Acetone</b> D. <b>Methanol</b>
11	The enhanced stability of conjugated dienes compared to isolated dienes is primarily attributed to	A. Inductive effects the double bonds B. Increases s-character of the hybridized orbitals C. <b>Delocalisation of electrons across the conjugated system</b> D. Steric hindrance between the double bonds
12	By 1700 A.D only----- elements were recognized.	A. <b>3</b> B. <b>6</b> C. <b>12</b> D. <b>10</b>
13	The pH of distilled water at 25 °C is	A. 6 B. <b>7</b> C. 5 D. 4
14	The most stable carbonium ion among the following is.	A. <b><math>(\text{CH}_3)_3\text{C}^+</math></b> B. $\text{CH}_3^+$ C. $\text{CH}_3\text{CH}_2^+$ D. $(\text{CH}_3)_2\text{CH}^+$

15	Which species is both a Bronsted acid and base.	<p>A. <math>\text{H}_2\text{O}</math></p> <p>B. <math>\text{Na}^+</math></p> <p>C. <math>\text{OH}^-</math></p> <p>D. <math>\text{Cl}^-</math></p>
16	Which one of the following effects explains the cooling during the sudden expansion of a gas.	<p>A. Boyle's Effect</p> <p>B. Joule Thomson Effect</p> <p>C. Charles Effect</p> <p>D. Avogadro's law</p>
17	The oxidation number of sulfur in $\text{H}_2\text{SO}_4$ is	<p>A. +2</p> <p>B. +4</p> <p>C. +6</p> <p>D. +8</p>
18	Which of the following is a Lewis acid but not a Brontsted Lowry acid	<p>A. <math>\text{HCl}</math></p> <p>B. <math>\text{NH}_3</math></p> <p>C. <math>\text{AlCl}_3</math></p> <p>D. <math>\text{H}_2\text{O}</math></p>
19	Which of the following is an aromatic hydrocarbon	<p>A. Ethene</p> <p>B. Benzene</p> <p>C. Methane</p> <p>D. Cyclohexane</p>
20	The mass of $1.55 \times 10^{23}$ atoms of Si	<p>A. 0.5 g</p> <p>B. 8 g</p> <p>C. 0.7 g</p> <p>D. 0.6 g</p>