

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Addition of inert gas at constant volume	A. Affects equilibrium B. Shifts reation left C. Shifts reaction right D. No effect
2	Which oxidation state will not shown by S	A2 B. +4 C. +8 D. +2
3	A flat curve in rate vs time gaph indicate.	A. Reaction completed B. High rate C. Constant rate D. Temperature increase
4	The Ka of a weak acid is 10-5 its pKa is	A. 2 B. 3 C. 5 D. 10
5	Which indicator is typically used for titration involving strong acids and strong bases.	A. Methyl red B. Phenolphthalein C. Bromothymol blud D. Litmus solution
6	Which of these is NOT a property of acids.	A. Conduct electricity B. Turn litmus red C. Sour taste D. Feel slippery
7	Unit of rate for gaseous reaction are.	A. mol dm-3 s-1 B. atm C. mol D. kg s-1
8	The term "lattice energy" is applicable to	A. lonic compunds B. Covalent compunds C. Gases D. Metals
9	What happens to the volue of water when it freeze	A. It contracts by 9% B. It remains the same C. It expends by 9% D. It evaporates completely
10	Unit of Kc depend on	A. Change in concentration B. Change in nuebr of moles of gas C. Chage in pressure D. Change in Entropy
11	Rate law is an expression involving.	A. Concentration B. Pressure C. Temperature D. Catalyst
12	What is the role of a pressur ecooker in boiling water.	A. It lowers atmospheric pressure B. It allows water to evaporate quickly C. It increases external pressure raising boiling point D. It cools the steam
13	Rate determining step controls	A. Overall rate B. Temperature C. Volume D. Final yield
14	Which of the following statemetns correctly describes the effect of temperature on the equilirium constant.	A. Kc is directly proportional to temperature B. Kc is inversely proportional to temperature. C. Kc depends on the enthalpy change of the reaction

		D. I emperature has no effect on the value of Kc
15	Why is the compression of solids not possible	A. The particles are fixed in place and cannot move closer B. Their particles are widely spaced C. The partiles are charged D. Solids are made of gases
16	KOH is an example of	A. Strong Base B. Weak acid C. Neutral salt D. Weak base
17	The change in internal energy is zero is.	A. Isothermal process B. Adiabatic process C. Exothermic process D. Endothermic process
18	The total number of oxygen atom is 22 g of CO2 gas is	A. 6.02 x 10 ²³ molecules B. 12.4 x 10 ²³ C. 12.04 x 10 ²³ D. 20 x 10 ²³
19	When 3d subshell is completely filed, the next entering electron goes into.	A. 4f B. 4 p C. 4s D. 4 d
20	The mass of one mole of electron is	A. 1.088 mg B. 0.55 mg C. 0.184 mg D. 1.67 mg