Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Electrochemical equivalent is.	A. Mass per mole B. Mass per coulomb C. Charge per second D. Current per mass
2	What are teh numebr f moles of hydrogen atom in 3.2 g of CH4?	A. 0.2 B. 0.4 C. 0.6 D. 0.8
3	Rate of reaction is measured as.	A. Increase Temperature B. Change is concentration per unit time C. Total eergy of reactants D. Change in volume per mole
4	Increase in pressur shifts equilibrium to.	A. Side with more moles of gas B. Sides with fewer moles of gas C. Liquid phase D. Solid phase
5	Which of these monomers forms a condensation polymer.	A. Ethane B. Styrene C. Adipic acid+ hexamethylenediamine D. Vinyl chloride
6	Half life of zero order reaction is	A. Constant B. Proportional to [A] C. Inversely proportional with [A] D. Independent to rate constant
7	The enthalpy of solution is	A. Heat evolved /absorbed when 1 mole of solute dissolves B. Solute dissolves C. Always exothermic D. always endothermic
		A. One <m:omathpara><m:omath> <m:r>σ </m:r> <m:r>and one </m:r> </m:omath> </m:omathpara> > > <m:omathpara><m:omath> <m:r> </m:r> </m:omath> </m:omathpara> > > <m:omathpara><m:omath> <span style="font-family:"Cambria

8	The numbers of σ *and* π bonds in the N2 molecule are.	style="font-family:"Cambria Math",serif"><m:r>π </m:r> <m:r>bonds </m:r> </m:omath></m:omathpara> B. The numbers of <m:omathpara><m:omath> <m:r>σ and two</m:r> </m:omath> </m:omathpara> <m:omathpara><m:omath> <m:r>π </m:r> <m:r>bonds </m:r> </m:omath></m:omathpara> C. Three <m:omathpara><m:omath> <m:r>σ </m:r> </m:omath> </m:omathpara> <m:omathpara><m:omath> <m:r> </m:r> <m:r><span style="font-size:11px;">bonds ionly</m:r> </m:omath></m:omathpara> D. Two <m:omathpara><m:omath> <m:r>σ and one</m:r> </m:omath>

		</m:omathpara> <m:omathpara><m:omath> <m:r>π </m:r> </m:omath> </m:omathpara>
9	The pH scale ranges typically from	A. 1 -10 B. 0 - 14 C. -1 to 1 D. 7 - 14
10	The shape of d-orbitals are like	A. Complicated B. Polar C. Spherical D. Cloverleaf
11	Electroyte in electroplating of silver	A. AgNo3 B. NaCl C. CuSO4 D. HNO3
12	What information does the pricnipal quantum numerb (n) gives us about orbitals	A. size B. Shape C. Spin D. Size and shape
13	The quantum number 'm' of a free gascous atom is associated with	A. The effective volume of the orbital B. The shape of the orbital C. The spatial orientation of the orbital D. The energy of the orbital in the absence of a magnetic field
14	Liquid crystals exhibit properties.	A. Only like solids B. Only like liquid C. Between solids and liquids D. Unliek solids or liquids
15	The SI unit of enthalpy	A. Calorie B. eV C. kj mol-1 D. J mol-1
16	Oxidation state fo P in P4 O10 is.	A. +2 B. +3 C. +5 D. +4
17	The number of moles of CH4 in 4 g of gas is	A. 0.2 moles B. 0.25 mole C. 0.3 mol D. 0.4 mole
18	What causes the lower compressibility of liquids compared to gaes.	A. Absence of intermolecular forces B. Larger mollecular size C. Stonger intermolecualr fores and less empty space D. High kinetic energy of molecules
19	A reversibel reaction is one which	A. Proceeds to completion B. Occurs only in one direction C. Proceeds in both directions D. Has no products
20	The quantum numebr values for the fourth electron of Be atom are.	A. 1,0,0 B. 2,0,0 C. 2,1,0 D. 1,1,1