

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	The solubility product of AgCl is $2.0 \times 10^{-10} \text{ mol}^2 \text{ dm}^{-6}$. The maximum concentration of Ag ⁺ ions in the solution is.	<p>A. $2.0 \times 10^{-10} \text{ mol dm}^{-3}$</p> <p>B. $1.41 \times 10^{-5} \text{ mol dm}^{-3}$</p> <p>C. $1.0 \times 10^{-10} \text{ mol dm}^{-3}$</p> <p>D. $4.0 \times 10^{-20} \text{ mol dm}^{-3}$</p>
2	Down cell is used for extractio of.	<p>A. Zinc</p> <p>B. Iron</p> <p>C. Sodium</p> <p>D. Aluminum</p>
3	The reversibel reation cannot be achieved in	<p>A. Open system</p> <p>B. Closed system</p> <p>C. Both a and b</p> <p>D. None of these</p>
4	Which of the following is a tribasic acid.	<p>A. H₃PO₄</p> <p>B. HNO₃</p> <p>C. HCl</p> <p>D. H₂SO₄</p>
5	On an energy profile diagram the presence of a catalyst is represented by	<p>A. A high peak representing the activation energy</p> <p>B. A lower peak representing the activation energy</p> <p>C. A change in the energy level of the reactants or products</p> <p>D. A shift in the equilibrium position</p>
6	The calorie content of food, often expressed in Calories (keal), is fundamentally related to which thermodynamic quantity during its metabolism or combustion.	<p>A. Enthalpy change</p> <p>B. Entropy change</p> <p>C. Gibbs free energy change</p> <p>D. Specific heat capacity</p>
7	Half life formula for 1st order reaction is	<p>A. $0.693 / k$</p> <p>B. $k \times t$</p> <p>C. $1 / k$</p> <p>D. $2 k$</p>
8	Alkanes react with halogens via.	<p>A. Electrophilic substitution</p> <p>B. Free radical substution</p> <p>C. Nucleophilic addition</p> <p>D. Electrophilic addition</p>
9	What is the SI unit of viscosity.	<p>A. Pascal</p> <p>B. $\text{kg m}^{-1} \text{ s}^{-1}$</p> <p>C. Joule</p> <p>D. Nm^{-2}</p>
10	$K_c = 0.040$ at 450°C for the teh given reaction, evaluate K_p for the reaction. $\text{PCl}_5 \rightleftharpoons \text{PCl}_2 + \text{Cl}_2$	<p>A. 0.40</p> <p>B. 2.4</p> <p>C. 0.64</p> <p>D. 0.052</p>
11	The order of a chemicla reaction, that is dependent of concentration is.	<p>A. Second order reaction</p> <p>B. First order reaction</p> <p>C. Zeroorder reaction</p> <p>D. Pseudo first order reaction</p>
12	Electrolysis is used in the extration of	<p>A. Silver</p> <p>B. Gold</p> <p>C. Aluminum</p> <p>D. Mercury</p>
13	The numebr of molesof CO ₂ which contain 8 g of oxygen	<p>A. 0.25</p> <p>B. 0.5</p> <p>C. 1.0</p> <p>D. 1.50</p>
		<p>A. 81 g water</p> <p>B. 6.02×10^{23}</p>

14	One mole of water contains.	<p>C. 6.02×10^{23} ions</p> <p>D. 6.02×10^{23} molecule</p> <p>E. 6.02×10^{23} atom</p>
15	Standard hydrogen electrode is assigned potential of.	<p>A. 1.0 V</p> <p>B. 0.5 V</p> <p>C. 0.0 V</p> <p>D. -1.0 V</p>
16	When a crystalline solid is broken, it does so along specific planes, These planes are known as.	<p>A. Cleavage planes</p> <p>B. Crystal faces</p> <p>C. Surface planes</p> <p>D. Growth planes</p>
17	A mixture of 8 g of H ₂ with 8 g of O ₂ is ignited $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ What is the mass of water formed.	<p>A. 9 g</p> <p>B. 36 g</p> <p>C. 16 g</p> <p>D. 72 g</p>
18	CH ₃ COONa is a salt of.	<p>A. Strong acid+ Strong base</p> <p>B. Weak acid+Strong base</p> <p>C. None</p> <p>D. Weak base+Weak acid</p>
19	Homogenous catalysts are in	<p>A. Same phase as reactions</p> <p>B. Different phase</p> <p>C. Solid only</p> <p>D. Gaseous only</p>
20	Unit of rate for gaseous reaction are.	<p>A. mol dm⁻³ s⁻¹</p> <p>B. atm</p> <p>C. mol</p> <p>D. kg s⁻¹</p>