

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which one of the following gases will have maximum volume at STP.	A. <p>88 g of N ₂ O</p> B. <p>22 g of CO ₂ </p> C. <p>28 g of CO</p> D. <p>28 g of N ₂ </p>
2	The pK _a value for HCOOH is.	A. 4.74 B. 3.78 C. 4.78 D. 4.24
3	Liquid crystals exhibit properties.	A. <p>Only like solids</p> B. <p>Only like liquid</p> C. <p>Between solids and liquids</p> D. <p>Unlike solids or liquids</p>
4	Electrochemical series helps in predicting.	A. Rate of reaction B. Type of bond C. Direction of redox reactions D. Heat released
5	In a titration, the equivalence point is when	A. No acid left B. pH = 7 C. Moles of acid = moles of base D. Buffer forms
6	A zero order reaction has rate	A. Proportional to [A] B. Independent of [A] C. Proportional to time D. Zero
7	The enthalpy change for a reaction depends on.	A. Pathway taken from reactants to products B. Presence of a catalyst C. Initial and final states of the reactants and products D. Rate of the reaction
8	Consider a reaction with $\Delta H > 0$ and $\Delta S < 0$ this reaction will be	A. Spontaneous at all temperatures B. Non spontaneous at all temperature C. Spontaneous only at high temperatures D. Spontaneous only at low temperature
9	Removing a reactant shifts equilibrium to	A. Left B. Right C. No shift D. Depends on temperature
10	If E _o cell is +1.10 V , the reaction is	A. Spontaneous B. Non spontaneous C. At equilibrium D. Endothermic only
11	Which of the following shows position isomerism.	A. But-1-ene and but-2-ene B. Ethene and ethyne C. Propane and propene D. Cyclohexane and benzene
12	A p orbital has a characteristic shape with how many lobes.	A. <p>1</p> B. <p>2</p> C. <p>3</p> D. <p>4</p>
13	Octet rule is not obeyed by which of the following compounds during its formation	A. <p>H ₂ O</p> B. <p>NaCl</p> C. <p>NH ₃ </p> D. <p>PF ₅ </p>
14	Which step in the Born-Haber cycle is always endothermic	A. Sublimation B. Electron gain enthalpy C. Hydration D. Lattice formation

15 The ratio of number of molecules in 71 g of gaseous chlorine to the number of molecules in 2 g of gaseous hydrogen is

A. $1:1$
B. $71:2$
C. $1:2$
D. $71:1$

16 Electrochemical equivalent is.

A. Mass per mole
B. Mass per coulomb
C. Charge per second
D. Current per mass

17 On a Boltzmann distribution curve, the activation energy is represented by

A. The height of the peak
B. The area under the entire curve
C. A vertical line drawn at a specific kinetic energy value
D. The difference between the peak and the X axis

18 Enthalpy change in hydration depends on.

A. Charge
B. Ion size
C. Solvent nature
D. All of these

19 When 3d subshell is completely filled, the next entering electron goes into.

A. $4f$
B. $4p$
C. $4s$
D. $4d$

20 The mass of 0.5 mole of Al is

A. 13.5g
B. 12g
C. 14 g
D. 2.7 g
