

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which one of the following gases will have maximum volume at STP.	A. $<p>88\text{ g of N}_2\text{O}</p>$ B. $<p>22\text{ g of CO}_2</p>$ C. $<p>28\text{ g of CO}</p>$ D. $<p>28\text{ g of N}_2</p>$
2	The pKa value for HCOOH is.	A. 4.74 B. 3.78 C. 4.78 D. 4.24
3	Liquid crystals exhibit properties.	A. $<p>\text{Only like solids}</p>$ B. $<p>\text{Only like liquid}</p>$ C. $<p>\text{Between solids and liquids}</p>$ D. $<p>\text{Unlielik solids or liquids}</p>$
4	Electrochemical series helps in predicting.	A. Rate of reaction B. Type of bond C. Direction of redox reactions D. Heat released
5	In a titration, the equivalence point is when	A. No acid left B. pH = 7 C. Moles of acid = moles of base D. Buffer forms
6	A zero order reaction has rate	A. Proportional to [A] B. Independent of [A] C. Proportional to time D. Zero
7	The enthalpy change for a reaction depends on.	A. Pathway taken from reactants to products B. Presence of a catalyst C. Initial and final states of the reactants and products D. Rate of the reaction
8	Consider a reaction with $\Delta H > 0$ and $\Delta S < 0$ this reaction will be	A. Spontaneous at all temperatures B. Non spontaneous at all temperature C. Spontaneous only at high temperatures D. Spontaneous only at low temperature
9	Removing a reactant shifts equilibrium to	A. Left B. Right C. No shift D. Depends on temperature
10	If E°_{cell} is +1.10 V, the reaction is	A. Spontaneous B. Non spontaneous C. At equilibrium D. Endothermic only
11	Which of the following shows position isomerism.	A. But-1-ene and but-2-ene B. Ethene and ethyne C. Propane and propene D. Cyclohexane and benzene
12	A p orbital has a characteristic shape with how many lobes.	A. $<p>1</p>$ B. $<p>2</p>$ C. $<p>3</p>$ D. $<p>4</p>$
13	Octet rule is not obeyed by which of the following compounds during its formation	A. $<p>\text{H}_2\text{O}</p>$ B. $<p>\text{NaCl}</p>$ C. $<p>\text{NH}_3</p>$ D. $<p>\text{PF}_5</p>$
14	Which step in the Born-Haber cycle is always endothermic	A. Sublimation B. Electron gain enthalpy C. Hydration D. Lattice formation

15	The ratio of numebr of molecules in 71 g of gaseous chloirine to the numebr of molecules in 2 g of gaseous hydrogen is	<p>A. $1:1$</p> <p>B. $71:2$</p> <p>C. $1:2$</p> <p>D. $71:1$</p>
16	Electrochemical equivalent is.	<p>A. Mass per mole</p> <p>B. Mass per coulomb</p> <p>C. Charge per second</p> <p>D. Current per mass</p>
17	On a Boltzmann distribution curve, the activation energy is represented by	<p>A. The height of the peak</p> <p>B. The area under the entire curve</p> <p>C. A vertical line drawn at a specific kinetic energy value</p> <p>D. The difference between the peak and the X axis</p>
18	Enthalpy change in hydration depends on.	<p>A. Charge</p> <p>B. Ion size</p> <p>C. Solvent nature</p> <p>D. All of these</p>
19	When 3d subshell is completely filed, the next entering electron goes into.	<p>A. $4f$</p> <p>B. $4p$</p> <p>C. $4s$</p> <p>D. $4d$</p>
20	The mass of 0.5 mole of Al is	<p>A. $13.5g$</p> <p>B. $12g$</p> <p>C. $14g$</p> <p>D. $2.7g$</p>