

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which of the following represents a position isomer of $\text{CH}_3\text{CH}_2\text{CH}=\text{CH}_2$	<p>A. C_2H_2</p> <p>B. $\text{CH}_3-\text{CH}=\text{CH}-\text{CH}_3$</p> <p>C. CH_4</p> <p>D. $\text{C}_2\text{H}_5\text{OH}$</p>
2	Standard electrode potential is measured under	<p>A. 100 oC and 1 atm</p> <p>B. 0 oC and 1 M concentration</p> <p>C. 25 oC and 1 M concentration</p> <p>D. 100 oC and 1 M concentration</p>
3	Covalent bonds are	<p>A. \rightarrowRight and directional</p> <p>B. \rightarrowNon rigid and directional</p> <p>C. \rightarrowRigid and non directional</p> <p>D. \rightarrowNon rigid and non directional</p>
4	The process of coating a metal with zinc is called.	<p>A. Alloying</p> <p>B. Galvanization</p> <p>C. Electrorefining</p> <p>D. Electroplating</p>
5	Which of these hydrocarbons is saturated	<p>A. Propane</p> <p>B. Ethyne</p> <p>C. Butane</p> <p>D. Cyclopentene</p>
6	Which metal is the best reducing agent.	<p>A. Cu</p> <p>B. K</p> <p>C. Zn</p> <p>D. Fe</p>
7	Which one is a non polar molecule.	<p>A. $\rightarrow \text{CS}_2$</p> <p>B. $\rightarrow \text{CaC}_2$</p> <p>C. $\rightarrow \text{H}_2\text{S}$</p> <p>D. $\rightarrow \text{CaCl}_2$</p>
8	When water freezes at 0 oC its density decrease due to.	<p>A. \rightarrowCubic structure of ice</p> <p>B. \rightarrowEmpty spaces; present; in the structure of ice</p> <p>C. \rightarrowDecrease in volume</p> <p>D. \rightarrowDecrease in viscosity</p>
9	Group 16 elements are also called.	<p>A. \rightarrowChalcogens</p> <p>B. \rightarrowpnictogens</p> <p>C. \rightarrowAlkali metals</p> <p>D. \rightarrowHalogenes</p>
10	For a specific reaction the value of the equilibrium constant, K_c ?	<p>A. Always remains the same at different reaction conditions</p> <p>B. Increases if the concentration of one of the products is increased</p> <p>C. Changes with changes in the temperature</p> <p>D. Increases if the concentration of one of the reactants is increased</p>
11	Which of the following metals would most readily displace hydrogen gas from dilute acids.	<p>A. Copper</p> <p>B. Silver</p> <p>C. Magnesium</p> <p>D. Platinum</p>
12	Why does acetone feel colder than water when poured on skin.	<p>A. \rightarrowIt reacts with skin</p> <p>B. \rightarrowIt has lower boiling point and evaporate faster</p> <p>C. \rightarrowIt is more viscous</p> <p>D. \rightarrowIt absorbs moisture</p>
13	The general formula for alkanes is.	<p>A. $\text{C}_n\text{H}_{2n+2}$</p> <p>B. $\text{C}_n\text{H}_{2n-2}$</p> <p>C. C_nH_{2n}</p> <p>D. $\text{C}_n\text{H}_{2n-1}$</p>
14	Law of mass action was proposed by	<p>A. Le chatelier</p> <p>B. Arrhenius</p>

		C. Guldberg and Waage D. Dalton
15	Suppose there are 100 molecules of a gas initially in jar A, which is connected to an evacuated jar B, When the stopcock is opened, the positive ways of arrangement of molecules will be	A. 100 B. 1/100 C. 2/100 D. 1/2
16	The ratio of numebr of molecules in 71 g of gaseous chloirne to the numebr of molecules in 2 g of gaseous hydrogen is	A. 1:1 B. 71:2 C. 1:2 D. 71:1
17	Active mass means	A. Moles B. Volume C. Mass D. Molar concentration
18	Na burns with oxygen to give.	A. Golden yellow flame B. Golden brown flame C. Crimson red flame D. Intense white flame
19	At equilirbium	A. Products dominate B. Reactants dominate C. Forward and reverse rates are zero D. Forward rate = reverse rate
20	In pressure of catalyst ΔH	A. Increases B. Decreases C. Double D. Unchanged