

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Faradya's First Law states tass deposited is direlty proportional tohat m	A. Time B. Voltage C. Resistance D. Quantity of electricity passed
2	Element that can show variable oxidation states.	A. Mg B. Ca C. S D. Mg
3	Which has the highest pH	A. 0.01 M NaOH B. 0.1 M NaOH C. 0.1 M HCl D. 0.01 M HCl
4	Unit of rate for gaseous reaction are.	A. $\text{mol dm}^{-3} \text{ s}^{-1}$ B. atm C. mol D. kg s^{-1}
5	Which oxidation state will not shown by S	A. -2 B. $+4$ C. $+8$ D. $+2$
6	Hydrocarbon are compounds made of.	A. Hydrogen and oxygen B. Carbon and hydrogen C. Carbon , hydrogen, and oxygen. D. Carbon and oxygen
7	Which of the following will show paramagnetic property.	A. O_2 B. N_2 C. O^{2-} D. O_2^{+2}
8	If a chemicial reactio has $\Delta H = -100 \text{ kJ/mol}$, it is	A. Exothermic B. Endothermic C. Isothermal D. Isobaric
9	Which is the correct trend in variation of electronegativity along a period of the periodic table.	A. It decreases from left to right across a period B. It increases from left to right across a period C. It remains constant D. It has no definite trend
10	Half life of zero order reaction is	A. Constant B. Proportional to $[A]$ C. Inversely proportional with $[A]$ D. Independent to rate constant
11	Which gas has highest molar enthalpy of combustion	A. C_2H_2 B. CH_4 C. H_2 D. CO
12	Which is amphoteric	A. H_2O B. HCl C. NaOH D. NH_3
13	If E° cell is $+1.10 \text{ V}$, the reactio is	A. Spontaneous B. Non spontaneous C. At equilibrium D. Endothermic only
14	The metal deposited at cathode during electrolysis of NaCl is.	A. Cl_2 B. Na C. H_2 D. Cu
15	The oxidation number of sulfur in H_2SO_4 is	A. +2 B. +4 C. +6

16 The conjugate base fo H_2CO_3 is

A. HCO_3
B. CO_2
C. CO_3
D. H_2CO_3

17 A salt bridge is usually made from

A. Metallic wires
B. Porous glass
C. Inert salt solution in agar agar
D. Hydrochloric acid

18 Peroxides are those in whcih oxidation state of oxygen is

A. $<\text{p}>-1</\text{p}>$
B. $<\text{p}>+2</\text{p}>$
D. $<\text{p}>-2</\text{p}>$
E. $<\text{p}>-1/2</\text{p}>$

19 No two electrons in an atom have same all four quantum numbers, this is called.

A. $<\text{p}>\text{Aufbau Principle}</\text{p}>$
B. $<\text{p}>\text{Hund's rule}</\text{p}>$
C. $<\text{p}>\text{Pauli's excluding principle}</\text{p}>$
D. $<\text{p}>\text{None}</\text{p}>$

20 Salts of weak acid + strong base are.

A. Basic
B. Acidic
C. Amphoteric
D. Neutral