

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	A substance that changes color with pH is a	A. Orange B. Red C. Yellow D. Blue
2	Which of the following is an extensive property.	A. Enthalpy B. Temperature C. Density D. Pressure
3	Electrolysis of brine produces.	A. Hydrogen, oxygen , NaOH B. Sodium, Chlorine , H ₂ O C. Hydrogen, Chlorine, NaOH D. NaOH , Cl ₂ , CO ₂
4	Standard electrode potential is measured under	A. 100 oC and 1 atm B. 0 oC and 1 M concentration C. 25 oC and 1 M concentration D. 100 oC and 1 M concentration
5	Which compound is a cycloalkane.	A. Butene B. Cyclopentane C. Benzene D. Cyclopentane
6	Catalyst usually act in	A. Initiation step B. Rate determining step C. Fast step D. Intermediate formation
7	Salts of weak acid + strong base are.	A. Basic B. Acidic C. Amphoteric D. Neutral
8	Consider the gas phase equilibrium system represented by the equation $2\text{H}_2 + \text{O}_2 \rightleftharpoons 2\text{H}_2\text{O}$. Given that the forward reaction is endothermic, which of the following changes will decrease the equilibrium amount of H ₂ O.	A. Adding more oxygen B. Adding a solid phase catalyst C. Decreasing the volume of the container D. Increasing the temperature at constant pressure
9	Reaction quotient Q is calculated using	A. Initial concentration B. Equilibrium concentrations C. Temperature D. Catalyst
10	The Bronsted Lowry definition identifies acid as.	A. Electron pair acceptors B. Proton donors C. Proton acceptors D. Electron pair donors
11	Which factor affects lattice energy	A. Ion size B. Ion charge C. Crystal structure D. None of these
12	When does a liquid start boiling.	A. When temperature reaches 100 oC B. When its vapor pressure equals atmospheric pressure C. When all molecules turn into vapour D. When surface tension becomes zero
13	Which of the following has all the three characters i.e. ionic, covalent and coordinate covalent.	A. H ₂ O B. NH ₃ C. KBr D. NH ₄ Cl
14	What types of force is hydrogen bonding.	A. Ionic bond B. A special type of dipole-dipole force C. Metallic Bond

C. ~~p>metallic bond</p>~~
D. ~~p>London Dispersion force</p>~~

15 Removing product from equilibrium

- A. Shifts equilibrium left
- B. Stops reaction
- C. Shifts equilibrium right
- D. Has no effect

16 Square root of the frequency of X-rays is.....atomic number

- A. ~~p>Directly proportional to</p>~~
- B. ~~p>Inversely Proportional to</p>~~
- C. ~~p>Equal to</p>~~
- D. ~~p>Half of</p>~~

17 Effective collisions are those that.

- A. Have enough energy and proper orientation
- B. Are elastic
- C. Are in vacuum
- D. Produce no change

18 Atoms having same atomic mass but different atomic number are called.

- A. ~~p>Isobars</p>~~
- B. ~~p>Isotopes</p>~~
- C. ~~p>Isotones</p>~~
- D. ~~p>Isoelectrons</p>~~

19 Which of the following is a characteristic of crystalline solids.

- A. ~~p>They do not have a definite geometrical shape</p>~~
- B. ~~p>They melt over a wide temperature range</p>~~
- C. ~~p>They have a sharp melting point</p>~~
- D. ~~p>Their properties do not depend on the direction of measurement</p>~~

20 John Newlands gave the idea of.

- A. ~~p>Law of triads</p>~~
- B. ~~p>Law of Octaves</p>~~
- C. ~~p>Modern Periodic law</p>~~
- D. ~~p>Curves between weight and volume</p>~~