

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Electroysis of brine produces.	A. Hydrogen, oxygen, NaOH B. Sodium, Chlorine, H2O C. Hydrogen, Chlorine, NaOH D. NaOH, Cl2, CO2
2	Mass of neutron in amu is.	A. 1.0073 B. 1.0067 C. 1.0087 D. 1.0000
3	The numer of possible isomers of C4H10 is	A. 1 B. 2 C. 3 D. 4
4	Equilibrium constant depends on	A. Pressure B. Temperature C. Volume D. Concentration
5	Ethene belongs to which class of hydrocarbons.	A. Alkane B. Alkene C. Alkyne D. Aromatic
6	Pressure used in Haber Process.	A. 2 atm B. 300 atm C. 200 atm D. 5 atm
7	Which is a pair of geometrical isomers.	A. Ethane and propane B. But-1- ene and but-2- ene C. eis-but-2- ene and trans -but-2- ene D. Benzene and toluene
8	pH of 10 ⁻⁹ M HCl is	A. 5 B. 9 C. Slightly below 7 D. Exactly -9
9	Increase in surface area	A. Increases rate B. Decreases rate C. No effect D. Stops reaction
10	If the concentration of CI- ion in a solution is increaed, the soubility of silver chloride will	A. Decreasse B. Increase C. Remain unchanged D. Become zero
11	Which of the following is not a free radical	A. CH3 B. CI+ C. O2 D. NO+
12	Delta H for an endothermic reaction is	A. Positive B. Negative C. Zero D. Depends on temperature
13	Which of the following is essential for hydrogen bond formation.	A. Hydrogen bonded to a metal B. Hydrogen bonded to high electongtive atoms F,O, or N C. Hydogen boned to non pola ratom D. Present of pi bond
14	The actual bond length in a polar covalent compund is.	A. Lesser than expected B. Greeater ahtn expected C. Equal to the expected D. Exactly half of expected

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15	For a redox reaction to be spontaneous, Delga G should be	A. Positive B. Negative C. Zero D. Constant
16	What happens to the particles in a solid when the temperature increases.	A. They move closer to each other B. Their vibrational motion increases C. They start moving freely D. The solid melts
17	Oxidation state fo P in P4 O10 is.	A. +2 B. +3 C. +5 D. +4
18	The pH of a buffer doesn't chagne much when	A. Diluted with water B. Acid or base is added C. Heated D. Temp changes
19	When a bond is formed	A. Energy is absorbed B. Energy is released C. Delta H is always zero D. No energy change
20	Which one of the following is not the applicatio of solubility products.	A. Determination of solubility from Ksp B. Common ion effect C. Predicting precipitation D. Dissociation of bonds