

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Noble gases are present in 18 group also called	A. I A B. II B C. VIII A D. VIII B
2	Which hydrocarbon has a higher boiling point due to longer chain length.	A. C <sub>6</sub> H <sub>14</sub> B. C <sub>2</sub> H <sub>6</sub> C. C <sub>2</sub> H <sub>8</sub> D. CH <sub>3</sub>
3	d-block elements are also called.	A. Inner transition B. Outer transition C. Typical transition D. None
4	Which of the following is always negative in an exothermic reaction.	A. Delta H B. Activation energy C. Entropy D. Delta S
5	Recognize the element if it has 3 electron shells, belongs to 's' block and has 2 electrons in its outer most shell	A. Calcium B. Sodium C. Magnesium D. Potassium
6	K <sub>c</sub> = 0.040 at 450 °C for the given reaction, evaluate K <sub>p</sub> for the reaction. PCl <sub>5</sub> ⇌ PCl <sub>3</sub> + Cl <sub>2</sub>	A. 0.40 B. 2.4 C. 0.64 D. 0.052
7	Unit of half life is	A. mol B. s C. mol dm <sup>-3</sup> D. J
8	Which law states that enthalpy change is independent of the path taken.	A. Hess's law B. Boyle's law C. Avogadro's law D. Dalton's law
9	Which of the following causes entropy to increase	A. Condensation B. Freezing C. Evaporation D. Crystallization
10	Unit of rate constant for zero order reaction is	A. mol dm <sup>-3</sup> s <sup>-1</sup> B. s <sup>-1</sup> C. mol <sup>-1</sup> dm <sup>3</sup> s <sup>-1</sup> D. mol <sup>-2</sup> dm <sup>6</sup> s <sup>-1</sup>
11	Why is the compression of solids not possible	A. The particles are fixed in place and cannot move closer B. Their particles are widely spaced C. The particles are charged D. Solids are made of gases
12	The test for unsaturation in a hydrocarbon that disappears the colour is.	A. Reaction with KMnO <sub>4</sub> B. Reaction with NaOH C. Decolorization of Br <sub>2</sub> Water D. Precipitate with AgNO <sub>3</sub>
13	According to law of mass action rate of reaction is proportional to.	A. Temperature B. Pressure C. Product of active masses D. Atomic mass
14	Electronic cloud around the nucleus is uniformly distributed in all directions in.....subshell	A. f B. s C. d D. p

15	4.0 g of NaOH (molar mass 40 g mol <sup>-1</sup> ) contains same number of sodium ions as are present in	<p>A. 58.5 g of NaCl</p> <p>B. 58.5 g of NaCl</p> <p>C. 76 g Na<sub>2</sub>SO<sub>4</sub></p> <p>D. 85 g of NaNO<sub>3</sub></p>
16	One mole of SO <sub>2</sub> contains.	<p>A. 6.02 × 10<sup>23</sup> atoms of oxygen</p> <p>B. 6.022 × 10<sup>23</sup> atoms of sulphur</p> <p>C. 4 g atoms of SO<sub>2</sub></p> <p>D. 18.1 × 10<sup>23</sup> moles of SO<sub>2</sub></p>
17	Electrolysis of CuSO <sub>4</sub> using copper electrodes results in	<p>A. Increase in electrolyte concentration</p> <p>B. Decrease in Cu<sup>2+</sup> concentration</p> <p>C. No change in electrolyte composition</p> <p>D. Formation of new compound</p>
18	Orbitals having same energy are called.	<p>A. Hybrid orbitals</p> <p>B. Valence orbitals</p> <p>C. Degenerate orbitals</p> <p>D. d-orbital</p>
19	The number of possible isomers of C <sub>4</sub> H <sub>10</sub> is	<p>A. 1</p> <p>B. 2</p> <p>C. 3</p> <p>D. 4</p>
20	Which condition is essential for halogenation of methane.	<p>A. UV light</p> <p>B. Acid catalyst</p> <p>C. Heat only</p> <p>D. Presence of water</p>