

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	In a galvanic cell, the anode is	A. Negative and site of oxidation B. Positive and site of oxidation C. Negative and site of reduction D. Positive and site of reduction
2	Which of the following species has unpaired electrons in antibonding molecular orbitals.	A. <chem>O2</chem> <sup>2+</sup> B. <chem>N2</chem> <sup>2-</sup> C. <chem>B</chem> D. <chem>F2</chem>
3	Which of the following statements about ideal gases is true.	A. <p>they have strong intermolecular forces</p> B. <p>Their particles have significant volume</p> C. <p>Their volume is mainly due to particle size</p> D. <p>They have negligible intermolecular forces</p>
4	For a specific reaction the value of the equilibrium constant, $K_c$ ?	A. Always remains the same at different reaction conditions B. Increases if the concentration of one of the product is increased C. Changes with changes in the temperature D. Increases if the concentration of one of the reactants is increased
5	Why is water's surface tension so high.	A. Due to its ionic nature B. Because of hydrogen bonding pulling surface molecules downward C. Because it has low boiling point D. Because water molecules are very large
6	Geometrical isomerism is due to.	A. rotation around single bond B. Restricted rotation around double bond C. Electronegativity difference D. Difference in functional groups
7	The pH of distilled water at 25 °C is	A. 6 B. 7 C. 5 D. 4
8	Faraday's First Law states that deposited is directly proportional to that m	A. Time B. Voltage C. Resistance D. Quantity of electricity passed
9	The shape in which a crystal usually grows is called its.	A. Habit B. Crystal lattice C. Melting Point D. Cleavage plane
10	Which of the following causes entropy to increase	A. Condensation B. Freezing C. Evaporation D. Crystallization
11	The energy required to break a chemical bond is called.	A. Ionization energy B. Bond energy C. Enthalpy D. Activation energy
12	In an acid base titration, the equivalence point is reached when.	A. pH of the solution is 7.0 B. The indicator changes color C. Equal volumes of acid and base have been added D. The reaction stops

13 Polymerization of ethene gives.  
A. Polyvinyl chloride  
B. Polyethylene  
C. Polystyrene  
D. Teflon

---

14 The volume of occupied by 1.4 g of N1 at  
A.  $1.12 \text{ dm}^3$   
B.  $2.24 \text{ dm}^3$   
C.  $22.4 \text{ dm}^3$   
D.  $112 \text{ dm}^3$

---

15 One mole of carbon -12 has mass  
A.  $0.012 \text{ kg}$   
B.  $1 \text{ kg}$   
C.  $0.0224 \text{ kg}$   
D.  $12 \text{ kg}$

---

16 Which of the following affects bond energy.  
A. Bond length  
B. Bond length  
C. Atomic size  
D. All of these

---

17 Which of the following molecules has a dipole moment.  
A.  $\text{CO}_2$   
B.  $\text{CS}_2$   
C.  $\text{SO}_2$   
D.  $\text{CCl}_4$

---

18 If 1 Faraday of electricity is passed the mass deposited equals.  
A. 1 gram equivalent  
B. 1 gram  
C. 1 mole  
D. 1 atm

---

19 Electrochemical cells convert chemical energy into  
A. Electrical energy  
B. Heat  
C. Light  
D. Nuclear energy

---

20 Which of the following represents a position isomer of  $\text{CH}_3\text{CH}_2\text{CH}=\text{CH}_2$   
A.  $\text{C}_2\text{H}_2$   
B.  $\text{CH}_3\text{-CH}=\text{CH-CH}_3$   
C.  $\text{CH}_4$   
D.  $\text{C}_2\text{H}_5\text{OH}$

---