

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	The units of the rate constant (k) for a reaction depend on the	A. Activation energy of the reaction B. Temperature of the reaction C. Overall order of the reaction D. Stoichiometry of the balanced chemical equation
2	For a redox reaction to be spontaneous, ΔG should be	A. Positive B. Negative C. Zero D. Constant
3	The pH of distilled water at 25 °C is	A. 6 B. 7 C. 5 D. 4
4	Earthenware pots keep water cool because.	A. They are made of clay B. They reflect sunlight C. They absorb water D. They allow water to evaporate through pores
5	Why is the compression of solids not possible	A. The particles are fixed in place and cannot move closer B. Their particles are widely spaced C. The particles are charged D. Solids are made of gases
6	$\text{HCl} + \text{NH}_3 \rightarrow \text{NH}_4\text{Cl}$ shows	A. Redox reaction B. Acid base reaction C. Bronsted acid base D. Precipitation
7	Why does acetone feel colder than water when poured on skin.	A. It reacts with skin B. It has lower boiling point and evaporate faster C. It is more viscous D. It absorbs moisture
8	Dipole moment of CO_2 is	A. 1.8 D B. 1.94 D C. 1.0 D D. Zero
9	The shape in which a crystal usually grows is called its.	A. Habit B. Crystal lattice C. Melting Point D. Cleavage plane
10	The mass of one mole of electron is	A. 1.088 mg B. 0.55 mg C. 0.184 mg D. 1.67 mg
11	Unit of rate for gaseous reaction are.	A. $\text{mol dm}^{-3} \text{ s}^{-1}$ B. atm C. mol D. kg s^{-1}
12	The term "Halogen" mean	A. Ash former B. Salt Former C. Copper giver D. Sulphur giver
13	Which of the following is most stable.	A. Tertiary radical B. Primary radical C. Secondary radical D. Methyl radical
14	The pH scale ranges typically from	A. 1 - 10 B. 0 - 14 C. -1 to 1 D. 7 - 14

15	An Increase in temperature generally	A. No Effect B. Increases cell voltage C. Decreases cell voltage D. Depends on pressure
16	The metal deposited at cathode during electrolysis of NaCl is.	A. Cl ₂ B. Na C. H ₂ D. Cu
17	specific heat capacity is amount of heat needed to raise temp of	A. 1 mole B. 1 gram C. 1 kg D. 100 grams
18	Which of the following is essential for hydrogen bond formation.	A. Hydrogen bonded to a metal B. Hydrogen bonded to highly electronegative atoms F, O, or N C. Hydrogen bonded to non polar atom D. Present of pi bond
19	A p orbital has a characteristic shape with how many lobes.	A. 1 B. 2 C. 3 D. 4
20	Delta G tells about	A. Speed B. Mechanism C. Feasibility D. All of these