

Fsc Part 1 Chemistry MCQ's Test For Full Book

| Sr | Questions | Answers Choice |
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| 1 | The space orientation for f-subshell is. | A. 3 B. 7 C. 1 D. 5 |
| 2 | What intermediate is formed during the electrophilic addition of HBr to an alkene. | A. Carbonation B. Carbanion C. Radical D. Epoxide |
| 3 | Which component maintains electrical neutrality in an electrochemical cell? | A. Salt bridge B. Voltmeter C. Electrolyte D. Electrodes |
| 4 | The number of lone pairs of electrons in an ammonium ion is | A. One B. Two C. Three D. Zero |
| 5 | The change in internal energy is zero in | A. Isothermal process B. Adiabatic process C. Exothermic process D. Endothermic process |
| 6 | Enthalpy of neutralization of a strong acid and a strong base is approximately | A. $+57.3 \text{ kJ/mol}$ B. -57.3 kJ/mol C. 0 D. $+5.73 \text{ kJ/mol}$ |
| 7 | ΔH for an endothermic reaction is | A. Positive B. Negative C. Zero D. Depends on temperature |
| 8 | John Newlands gave the idea of | A. Law of triads B. Law of Octaves C. Modern Periodic law D. Curves between weight and volume |
| 9 | If the electronegativity difference between bonded atoms is more than 0.5 but less than 1.7, the bond will be | A. Ionic B. Polar Covalent C. Non polar covalent D. Coordinate covalent |
| 10 | When water freezes at 0°C its density decreases due to | A. Cubic structure of ice B. Empty spaces in the structure of ice C. Decrease in volume D. Decrease in viscosity |
| 11 | K_c is expressed in terms of | A. Pressure B. Mole fraction C. Concentration D. Volume |
| 12 | At equilibrium | A. Products dominate B. Reactants dominate C. Forward and reverse rates are zero D. Forward rate = reverse rate |
| 13 | The number of moles of CH_4 in 4 g of gas is | A. 0.2 moles B. 0.25 mole C. 0.3 mol D. 0.4 mole |
| 14 | Enthalpy of fusion is the heat required to | A. Melt a solid B. Boil a liquid C. Freeze a liquid D. Vaporize a solid |

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| 15 | Why does acetone feel colder than water when poured on skin. | <p>A. It reacts with skin</p> <p>B. It has lower boiling point and evaporate faster</p> <p>C. It is more viscous</p> <p>D. It absorbs moisture</p> |
| 16 | On Boltzmann distribution curve, the area under the curve represents. | <p>A. Activation energy of the reaction</p> <p>B. Total number of molecules in the sample</p> <p>C. Average kinetic energy of the molecules</p> <p>D. Rate constant of the reaction</p> |
| 17 | Highest bond order is found in | <p>A. O_2</p> <p>B. O_2^{-1}</p> <p>C. O_2^{+2}</p> <p>D. O_2^{+1}</p> |
| 18 | The number of moles of CO_2 which contain 8 g of oxygen. | <p>A. 0.25</p> <p>B. 0.5</p> <p>C. 1.0</p> <p>D. 1.50</p> |
| 19 | The reaction at the cathode is always | <p>A. Oxidation</p> <p>B. Neutral</p> <p>C. Reduction</p> <p>D. Ionization</p> |
| 20 | L- Pauling developed a scale for | <p>A. Ionization energy</p> <p>B. Electron affinity</p> <p>C. Electronegativity</p> <p>D. None</p> |