Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which one is a polar molecule	A. HCl B. CO2 C. BF3 D. CCl4
2	Rate constant can be determined by	A. Titration B. spectroscopy C. conductometry D. all of these
3	Highest bond order is found in	A. O2 B. O2 ⁻¹ C. O2 ⁺² D. O2 ⁺¹
4	Which order has constant half life	A. First B. Zero C. Second D. Third
5	The rate determinig step in a multistep reaction is	A. The slowest step B. Always the first step C. Always the last step D. The fastest step
6	How many unparied electrons are present in an atom of cobalt.	A. Two B. Three C. Four D. Five
7	What information does the pricnipal quantum numerb (n) gives us about orbitals	A. size B. Shape C. Spin D. Size and shape
8	The anode is electrolysis of molten NaCl is	A. Na+ B. Cl- C. Na D. H-
9	The oxidation number of sulfur in H2SO4 is	A. +2 B. +4 C. +6 D. +8
10	Enthalpy of fusion is the heat required to.	A. Melt a solid B. Boil a liquid C. Freez a liquid D. Vaporize a solid
11	s-subshell has.....orbital	A. 1 B. 2 C. 3 D. 4
12	CH3COONa in water forms	A. Acidic solution B. Basic Solution C. Neutral Solution D. Salt bridge
13	A saturted solution represents a dynamic equilibrium Macroscopically, the concentration of dissolved solute is constant, Microscopically this occurs because.	A. No more solute particles are dissolving B. The rate of dissolution of solute is zero C. Solute particles are dissolving and precipitating at the same rate D. All solute particles have dissolved
14	The most electronegative atom is	A. F B. Cl C. N D. O
		A. Bond breaking

15	A negative lattice eneryg temperature	B. Energy relresed C. Formation of covalent bond D. Spontaneity
16	Which species is both a Bronsted acid and base.	A. H2O B. Na+ C. OH- D. Cl-
17	A process with increase in entropy and enthalpy is spontaneous at.	A. High temperature B. Low temperature C. All temperature D. never spontaneous
18	Which of the following species has unparied electonrs in antibonding molecular orbitals.	A. O2 ²⁺ B. N 2 ²⁻ C. B D. F2
19	Catalyst used in Haber process is	A. Fe B. Pt C. Cu D. Zn
20	If a chemcial reactio has Delta H = -100 kj/mol, it is	A. Exothermic B. Endothermic C. Isothermal D. Isobaric