

Chemistry Fsc Part 1 Online Test

| Sr | Questions | Answers Choice |
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| 1 | The atomic radius geneallyacross a period in the periodic table. | A. Increase B. Decreases C. Remains constant D. First increasses then decreases |
| 2 | Number of molesin 100 g of KClO3 | A. 0.76 B. 0.56 C. 0.816 D. 0.014 |
| 3 | Oxidation state fo P in P4 O10 is. | A. +2 B. +3 C. +5 D. +4 |
| 4 | Reaction in Haber process is | A. Endothermic B. Exothermic C. Irrversible D. Neutral |
| 5 | Why can gases be compressed easily. | A. Because they have strong intermolecular forces B. Becasue their particles are lightly packed C. Because most of their volume is empty space D. Because they are denser than solids |
| 6 | One moleof crbon -12 has mass | A. 0.012 kg B. 1 kg C. 0.0224 kg D. 12 kg |
| 7 | John Newlands gavt he idea of. | A. Law of triads B. Law of Octaves C. Modern Periodic law D. Curves between weight and volume |
| 8 | Which of the following is an irreversible reaction. | A. Haber process B. Precipitation of Ag Cl C. Synthesis of ammonia D. Esterification |
| 9 | Which of the following has minimum mass | A. 3 x 10 ²³ atom of C B. 1 mole of S C. 79 gram of Ag D. 2 gram atomsof N |
| 10 | Which is more acidic pH3 or pH 5? | A. Both same B. pH 3 C. pH5 D. Cannot say |
| 11 | Which one of the following Statments is incorrect | A. One mole of nitrogen gas contains Avogadro's numebr of moleclues B. One mole of ozone gas contains avogadr's number of molecules C. One mole of ozone contains avogadro's number of O atom D. One mole of hydrogen gas contains Avogadro's numebrof molecules |
| 12 | Which of the following compounds exhibites permanent dipole permanent dipole intersctions. | A. HCI B. CI2 C. CH4 D. O2 |

| 14 What is the number of moless of oxygen in 11 g of CO2 S. cp-0.05c/p> B. cp-0.05c/p> C. cp-0.75c/p> D. cp-1.0c/p> A Initial conceration B. Equilibrium concentrations C. Temperature D. Catalyst | 13 | The standard enthalpy of atomization of an element is alwyas. | A. Negative B. Positive C. Zero D. Depend on element |
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| B. Equlibrium concentrations C. Temperature D. Catalyst | 14 | What is the number of moless of oxygen in 11 g of CO2 | B. 0.50 C. 0.75 |
| 16 | 15 | Reaction quotient Q is calculated using | B. Equlibrium concentrations C. Temperature |
| The number of moles of CO2 whoi contain 8 g of oxygen. B. C. < | 16 | If the pH of solution is 11, what is the [OH-]concentratio in the solution. | B. 1 x 10 ⁻¹¹ M C. 1 x 10 ⁻² M |
| Rate determining step controls B. Temperature C. Volume D. Final yield A. Second order reaction B. First order reaction C. Zeroorder reaction C. Zeroorder reaction D. Pseudo first order reaction A. Increase Temperature B. Change is concentration per unit time C. Total eergy of reactants | 17 | The number of moles of CO2 whoi contain 8 g of oxygen. | B. 0.5 C. 1.0 |
| The order of a chemicla reaction, that is dependent of concentration is. B. First order reaction C. Zeroorder reaction D. Pseudo first order reaction A. Increase Temperature B. Change is concentration per unit time C. Total eergy of reactants | 18 | Rate determining step controls | B. Temperature C. Volume |
| 20 Rate of reaction is measured as. B. Change is concentration per unit time C. Total eergy of reactants | 19 | The order of a chemicla reaction, that is dependent of concentration is. | B. First order reactionC. Zeroorder reaction |
| | 20 | Rate of reaction is measured as. | B. Change is concentration per unit time C. Total eergy of reactants |