Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	Half life is time is which	A. Rate becomes zero B. Concentration reduces to 1/4 C. Concentratio reduces to half D. Product is half formed
2	On an energy profile diagram the presence of a catalyst is represented by	A. A highe rpeak representing the activation energy B. A lower peak representing the activation energy C. A change in the energy level of the reactants or products D. A shift in the equilibrium position
3	If Delta n is positive, then Kp is	A. Greater than Kc B. Less than Kc C. equal Kc D. Zero
4	The bond formation energy of a compund is.	A. Equal to bond idssociation energy B. Less thanbond idssociation energy C. Greater than bond dissociaton energy D. Inverselyproportional to bond dissociation energy
5	Element that can show variable oxidation states.	A. Mg B. Ca C. S D. Mg
6	Enthalpy change in hydration depends on.	A. Charge B. Ion size C. Solvent nature D. All of these
7	Catalyst used in contact process	A. Fe B. V2O5 C. Ni D. Al2O3
8	What is the SI unit of viscosity.	A. Pascal B. kg m-1 s-1 C. Joule D. Nm-2
9	In which subshell electrons are filled first according to n+1 rule.	A. 2p B. 3p C. 4p D. 4s
10	In which of the following molecules ionic bond is found.	A. CsI B. NH3 C. HF D. CaC2
11	The number of atoms present in 6 g of Mg metalis.	A. 12.4 x 10²³ B. 6.02 x 10²³ C. 1.5 x 10²³ D. 3.01 x 10²³
12	Which of the followng has the lowest molar hat of vaporization.	A. HCl B. Water C. NH3 D. Ethanol
13	The enthalpy change for a reaction depends on.	A. Pathway taken from reactants to products B. Presence of a catalyst C. Intitial and final states of the

		reactants and products D. Rate of the reaction
14	A zero order reaction has rate	A. Proportional to [A] B. Independent of [A] C. Proportional to time D. Zero
15	The sum of atomic masses of all the elements present in a oles in called.	A. Atomic mass B. Ionic mass C. Molecular mass D. Emperical formula mass
16	Which is more acidic pH3 or pH 5?	A. Both same B. pH 3 C. pH5 D. Cannot say
17	Rate = k [A] ; defines	A. Stoichiometry B. Reaction order C. Reaction mechanism D. Titration
18	For a specific reaction the valu eof the equilibrium constant, Kc?	A. Always remains the same at differrent reaction conditions B. Increases if the concentration of one ofthe product is increased C. Changes with changes in the temperature D. Increasees if the concentration of one of the ratants is increased
19	What i the rearranged form of the idela gas equation to calculate molar mass of a gas.	A. M=PVn/RT B. M =PV /MRT C. M = m RT/PV D. M = RT/PV m
20	The term "lattice energy" is applicable to	A. Ionic compunds B. Covalent compunds C. Gases D. Metals