

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which of the following has minimum mass	<p>A. 3×10^{23} atom of C</p> <p>B. 1 mole of S</p> <p>C. 79 gram of Ag</p> <p>D. 2 gram atomsof N</p>
2	A mixture of 8 g of H ₂ with 8 g of O ₂ is ignited $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ What is the mass of water formed.	<p>A. 9 g</p> <p>B. 36 g</p> <p>C. 16 g</p> <p>D. 72 g</p>
3	Dipole moment is the product of.	<p>A. Charge x displacement</p> <p>B. Charge x distance</p> <p>C. Newton x displacement</p> <p>D. Charge x mass</p>
4	A buffer solution is.	<p>A. Stron acid+ strong base</p> <p>B. Weak Acid + Its salt</p> <p>C. Weak base + Its salt</p> <p>D. Weak base+ salt of strong acid</p>
5	Which property increases as you go down a group in the periodic table?	<p>A. Atomic radiium</p> <p>B. Electron affinity</p> <p>C. Electronegativity</p> <p>D. Ionization Energy</p>
6	Aromatic hydrocarbons must contain.	<p>A. Straight chains</p> <p>B. Double bonds only</p> <p>C. A benzene ring</p> <p>D. A riple bond</p>
7	Which one of the following effects explains the cooling during the sudden expansion of a gas.	<p>A. Boyle's Effect</p> <p>B. Joule Thomson Effect</p> <p>C. Charles Effect</p> <p>D. Avogadr's law</p>
8	A solution contains 4.0 g of sodium hydroxide in 250 cm ³ of solution. What is the molar concentration of this solution	<p>A. 0.10 mol dm⁻³</p> <p>B. 0.20 mol dm⁻³</p> <p>C. 0.40 mol dm⁻³</p> <p>D. 0.80 mol dm⁻³</p>
9	Which acid is present in vinegar	<p>A. CH₃COOH</p> <p>B. HNO₃</p> <p>C. H₂SO₄</p> <p>D. HCl</p>
10	The rate determinig step in a multistep reaction is	<p>A. The slowest step</p> <p>B. Always the first step</p> <p>C. Always the last step</p> <p>D. The fastest step</p>
11	What is teh oxidation state of sulfur in the sulfate ion SO ₄ ²⁻	<p>A. +4</p> <p>B. +6</p> <p>C. +2</p> <p>D. 0</p>
12	Heat of combustion of H ₂ is used to determine	<p>A. Calorific value</p> <p>B. Enthalpy</p> <p>C. Ionization energy</p> <p>D. Lattice energy</p>
13	Unit of K _c depend on	<p>A. Change in concentration</p> <p>B. Change in nuebr of moles of gas</p> <p>C. Chage in pressure</p> <p>D. Change in Entropy</p>
14	Na burns with oxygen to give.	<p>A. Golden yellow flame</p> <p>B. Golden brown flame</p> <p>C. Crimson red flame</p> <p>D. Intense white flame</p>
15	Ion an exothermic reaction , the nergy of products is	<p>A. Greater than reactants</p> <p>B. Less than reactants</p> <p>C. Equal to reactants</p>

		D. Zero
16	The intermediate formed during electrophilic addition	A. Free radical B. Carbocation C. Carbene D. Carbanion
17	Electrochemical cells convert chemical energy into	A. Electrical energy B. Heat C. Light D. Nuclear energy
18	For a redox reaction to be spontaneous, ΔG should be	A. Positive B. Negative C. Zero D. Constant
19	The sum of atomic masses of all the elements present in a molecule is called.	A. Atomic mass B. Ionic mass C. Molecular mass D. Empirical formula mass
20	Rate of fastest	A. At start B. In middle C. At end D. After completion