

## Fsc Part 1 Chemistry MCQ's Test For Full Book

| Sr | Questions  | Answers Choice   |
|----|--|--|
| 1  | How many subshells are present in M-Shell  | A. <p>1</p><br>B. <p>2</p><br>C. <p>3</p><br>D. <p>4</p>   |
| 2  | Optimum temperatur ein Haber process is  | A. 50 oC<br>B. 450 oC<br>C. 200 oC<br>D. 1000 oC   |
| 3  | Which statement is true at dynamic equilibrium.                                      | A. No reactio is occurring<br>B. Concentrations are changing<br>C. Rates of forward and reverse reactions are equal<br>D. Rate forward reaction < reverse  |
| 4  | A process with increase in entropy and enthalpy is spontaneous at.                   | A. High temperature<br>B. Low temperature<br>C. All temperature<br>D. never spontaneous  |
| 5  | The Ka of a weak acid is 10-5 its pKa is   | A. 2<br>B. 3<br>C. 5<br>D. 10  |
| 6  | Whcih of the following reactions reaches equilibrium                                 | A. Reversible<br>B. Irreversible<br>C. Combustion<br>D. Neutralization   |
| 7  | Addition of inert gas at constant volume   | A. Affects equilibrium<br>B. Shifts reation left<br>C. Shifts reaction right<br>D. No effect   |
| 8  | The metal deposited at cathode during electrolysis of NaCl is.                       | A. Cl2<br>B. Na<br>C. H2<br>D. Cu  |
| 9  | Which of the following factors would lead to a greater enthalpy chage of hydration . | A. A larger ionic reduis and a smaller charge<br>B. A smaller ionic radius and a smaller charge<br>C. A larger ionic radius and a larger charge<br>D. A smaller ionic radius and a larger charge |
| 10 | How many electrons are presente in the valence shell of P in PO4 3                   | A. <p>8</p><br>B. <p>10</p><br>C. <p>12</p><br>D. <p>14</p>  |
| 11 | What types of force is hydrogen bonding.   | A. <p>Ionic bond</p><br>B. <p>A special type of dipole dipole force</p><br>C. <p>Metalic Bond</p><br>D. <p>London Dispersion force</p>   |
| 12 | Which has the highest pH   | A. 0.01 M NaOH<br>B. 0.1 M NaOH<br>C. 0.1 M HCl<br>D. 0.01 M HCl   |
| 13 | The general formula for alkanes is.  | A. CnH2n+2<br>B. CnH2n-2<br>C. CnH2n<br>D. CnH2n-1   |
| 14 | Increase in pressur shifts equilibrium to.   | A. Side with more moles of gas<br>B. Sides with fewer moles of gas<br>C. Liquid phase<br>D. Gas phase  |

15 How many d orbitals are there in a given energy level.

A.  $1</p>$   
B.  $5</p>$   
C.  $3</p>$   
D.  $7</p>$

16 Which of the following hybridizations is found in ethyne.

A.  $Sp</p>$   
B.  $sp2</p>$   
C.  $sp3</p>$   
D.  $No\ hybridization</p>$

17 A p orbital has a characteristic shape with how many lobes.

A.  $1</p>$   
B.  $2</p>$   
C.  $3</p>$   
D.  $4</p>$

18 Which component maintains electrical neutrality in an electrochemical cell

A. Salt bridge  
B. Voltmeter  
C. Electrolyte  
D. Electrodes

19 In the electrophilic addition of HBr to propene, the major product is.

A. 1- bromopropane  
B. 2- bromopropane  
C. 3- bromopropane  
D. Isopropanol

20 The solubility product of AgCl is  $2.0 \times 10^{-10}$  mol<sup>2</sup> dm<sup>-6</sup>. The maximum concentration of Ag<sup>+</sup> ions in the solution is.

A.  $2.0 \times 10^{-10}$  mol dm<sup>-3</sup>  
B.  $1.41 \times 10^{-5}$  mol dm<sup>-3</sup>  
C.  $1.0 \times 10^{-10}$  mol dm<sup>-3</sup>  
D.  $4.0 \times 10^{-20}$  mol dm<sup>-3</sup>