Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which metal is the best reducing agent.	A. Cu B. K C. Zn D. Fe
2	Law of mass action was proposed by	A. Le chatelier B. Arrheneius C. Guldberg and Waage D. Dalton
3	The actual bond length in a polar covalent compund is.	A. Lesser thanexpected B. Greeater ahtn expected C. Equal to the expected D. Exactly half of expected
4	The solubility product of AgCl is 2.0 x 10^{-10} mol2 dm-6. The maximum concentration of Ag+ ions in the solution is.	A. 2.0 x 10⁻¹⁰ mol dm⁻³ B. 1.41 x 10⁻⁵ mol dm⁻³ C. 1.0 x 10⁻¹⁰ mol dm⁻³ D. 4.0 x 10⁻²⁰ mol dm⁻³
5	Which is a pair of geometrical isomers.	A. Ethane and propane B. But-1- ene and but-2- ene C. eis-but-2- ene and trans -but-2-ene D. Benzene and toluene
6	If 16 g of O2 react with excess C2H6,how many gamof CO2 will be formed. 2C2H6+7O2----------4CO2 +5H2O	A. 22 g B. 13 g C. 9 g D. 7g
7	Free radical reactions typically occur in	A. Aromatic hydrocarbons B. Saturated alkanes C. Unsaturated alkenes D. Alcohols
8	Electroysis of brine produces.	A. Hydrogen, oxygen , NaOH B. Sodium, Chlorine , H2O C. Hydrogen, Chlorine, NaOH D. NaOH , Cl2, CO2
9	A sample of nitrogen gas has a mass of 14.0 g,How many nitrogen atoms are present in the this sample	A. 3.01 x 10 ²⁵ atoms B. 6.02 x 10 ²³ atoms C. 12.01 x 10 ²⁵ atoms D. 2.40 x 10 ²⁵ atoms
10	Catalyst affects	A. Value of K B. Equilibrium position C. Activation energy D. Final concentrations
11	Consider a reaction with Delta H > 0 and Delta S<0 this reaction will be	A. Spontaneous at all temperaturres B. Non spontaneous at all temperature C. Spontaneous only at high temperatures D. Spontaneous only atlow temperature
12	Which of the following has maximum mass	A. 2 moles of P B. 5 moles of H2O C. 2 mole of Na2CO3 D. 1 mole of glucose
13	Chlorinatin of propane gives how many thpes of monochlorinated products.	A. 1 B. 2 C. 4

		D. 3
14	Which one of the following is not an example of reversible reaction.	A. Formation of ammonia B. Foramtion ow water C. Decomposition of PCl5 D. Decomposition of NO2
15	A solution contains 4.0 g of sodium hydroxide in 250 cm3 of solution. What is the molar concentration of this solution	A. 0.10 mol dm-3 B. 0.20 mol dm-3 C. 0.40 mol dm-3 D. 0.80 mol dm-3
16	The most stable carbonium ion among the following is.	A. (CH3)3C+ B. CH3+ C. CH3 CH2+ D. (CH3)2CH-
17	Alkenes undergo which type of reaction.	A. Nucloeophilic addition B. Electrophilic addition C. Free radical substitution D. Nucleophilic substitution
18	Which one is a monoprotic acid	A. H3PO4 B. H2SO4 C. HCl D. H2CO3
19	Consider the gas phase equilibrium sysem represented by teh equation 2 H2O..... 2H2+O2 Givent that the forward reactio is endothermic, which of the following chagnes will decrease the equilibrium amount of H2O.	A. Adding more oxygen B. Adding a solid phase catalyst C. Decreasing the volume of the container D. Increaseing the temperature at constant pressure
20	John Newlands gavt he idea of.	A. Law of triads B. Law of Octaves C. Modern Periodic law D. Curves between weight and volume