

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which species is both a Bronsted acid and base.	A. H ₂ O B. Na ⁺ C. OH ⁻ D. Cl ⁻
2	For gaseous reactions, rate is often monitored by	A. Mass change B. pH change C. Pressure change D. Temperature change
3	K _c = 0.040 at 450 °C for the following reaction, evaluate K _p for the reaction. PCl ₅ ⇌ PCl ₃ + Cl ₂	A. 0.40 B. 2.4 C. 0.64 D. 0.052
4	Anti-Markovnikov's addition occurs in the presence of	A. HBr only B. Organic peroxide C. Water D. KMnO ₄
5	If the electronegativity difference between bonded atoms is more than 0.5 but less than 1.7, the bond will be.	A. Ionic B. Polar Covalent C. Non polar covalent D. Coordinate covalent
6	Unit of K _c depends on	A. Change in concentration B. Change in number of moles of gas C. Change in pressure D. Change in Entropy
7	Effect of temperature on rate is	A. Negligible B. Linear C. Exponential D. Logarithmic
8	Addition of inert gas at constant volume	A. Affects equilibrium B. Shifts reaction left C. Shifts reaction right D. No effect
9	pH of 10 ⁻⁹ M HCl is	A. 5 B. 9 C. Slightly below 7 D. Exactly -9
10	What information does the principal quantum number (n) give us about orbitals	A. Size B. Shape C. Spin D. Size and shape
11	At equilibrium in a closed system, which two processes occur at the same rate.	A. Melting and Freezing B. Evaporation and condensation C. Sublimation and condensation D. Evaporation and boiling
12	Surface tension of liquid is due to.	A. Inward pull of surface molecules B. Upward pull from the surface C. Collision of molecules D. Repulsive forces
13	Why are cholesteric liquid crystals useful in medical diagnostics.	A. They kill bacteria B. They react with tumors chemically C. They detect temperature differences D. They absorb X-rays
14	The bond angle in ethene molecule is.	A. 120° B. 180°

		<p>C. 109°</p> <p>D. 107.5°</p>
15	A negative lattice energy temperature	<p>A. Bond breaking</p> <p>B. Energy released</p> <p>C. Formation of covalent bond</p> <p>D. Spontaneity</p>
16	Subshells in an atom are filled with electrons in an increasing order of their energy values called.	<p>A. Aufbau principle</p> <p>B. Hund's rule</p> <p>C. Pauli's Exclusion Principle</p> <p>D. None</p>
17	Which one is a polar molecule	<p>A. HCl</p> <p>B. CO₂</p> <p>C. BF₃</p> <p>D. CCl₄</p>
18	A p orbital has a characteristic shape with how many lobes.	<p>A. 1</p> <p>B. 2</p> <p>C. 3</p> <p>D. 4</p>
19	Electroplating involves.	<p>A. Electrolysis for coating</p> <p>B. Spontaneous oxidation</p> <p>C. Redox without electrodes</p> <p>D. Thermolysis</p>
20	Which element has $E^{\circ} = 0.00$ V?	<p>A. H⁺</p> <p>B. H₂</p> <p>C. SHE</p> <p>D. All of the above</p>