

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Two electrons in the same orbital have.	A. <p>Same Spin</p> B. <p>Differece spin</p> C. <p>May same or different</p> D. <p>Parallel spin</p>
2	Why does acetone feel colder than water when poured on skin.	A. <p>It reacts with skin</p> B. <p>It has lower boiling point and evaporate faster</p> C. <p>It is more viscous</p> D. <p>It absorbs moisture</p>
3	Overall order is sum of	A. Coefficients B. Exponents in rate law C. Moles D. Products
4	A reaction mechanism describes	A. Experimental conditions B. Overall stoichiometry C. Stepwise molecular events D. Heat of reaction
5	An orbital can accommodate how many electrons.	A. <p>1</p> B. <p>2</p> C. <p>3</p> D. <p>4</p>
6	The mass of one mole of electron is	A. <p>0.55 mg</p> B. <p>1.088 mg</p> C. <p>0.184 mg</p> D. <p>1.67 mg</p>
7	Which step in the Born-Haber cycle is always endothermic	A. Sublimation B. Electron gain enthalpy C. Hydration D. Lattice formation
8	At equilibrium	A. Products dominate B. Reactants dominate C. Forward and reverse rates are zero D. Forward rate = reverse rate
9	Formula for calculating the value of (Azimuthal quantum number)	A. <p>$n+1$</p> B. <p>$n-1$</p> C. <p>$n+2$</p> D. <p>$2n^2$</p>
10	The termination step in a radical chain reaction involves.	A. Chain breaking B. Light absorption C. Chain initiation D. Hydrogen abstraction
11	Liquid are approximately how many times less compressible than gases.	A. <p>10</p> B. <p>105</p> C. <p>100</p> D. <p>1000</p>
12	Which of the followign best describes the molecular arrangemtn in liquid crystals.	A. Random and disordered B. Fixed in all direcions C. Parallel and partially ordered D. Completely rigid and fixed
13	Which of the following statemetns correctly describes the effect of temperature on the equilibrium constant.	A. K_c is directly proportional to temperature B. K_c is inversely proportional to temperature. C. K_c depends on the enthalpy change of the reaction D. Temperature has no effect on the value of K_c
		A. Temperature and pressure B. Concentration of ions

14	Cell potential is the difference between	C. Electrode potentials of cathode and anode D. Mass of electrodes
15	Unit of rate constant for zero order reaction is	A. mol dm ⁻³ s ⁻¹ B. s ⁻¹ C. mol ⁻¹ dm ³ s ⁻¹ D. mol ⁻² dm ⁶ s ⁻¹
16	Oxidation number of oxygen in H ₂ O	A. 1 B. -2 C. 0 D. +1
17	A polymer used for non-stick cookware is.	A. Teflon B. Nylon C. Polyvinyl chloride D. Polythene
18	Sp ³ Hybridization is associated with structure.	A. Linear B. Tetrahedral C. Trigonal D. Octahedra
19	Which of the following contains same number of particles as present in 12 g of carbon	A. 28 fcc iron (Atomic mass of Fe = 56) B. 48 g of magnesium (Atomic mass of Mg = 24) C. 32 g of S molecules (Atomic mass S = 32) D. 44 g of carbon dioxide (molar mass of CO ₂ = 44)
20	Law of mass action was proposed by	A. Le Chatelier B. Arrhenius C. Guldberg and Waage D. Dalton