

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Two electrons in the same orbital have.	A. Same Spin B. Different spin C. May same or different D. Parallel spin
2	Why does acetone feel colder than water when poured on skin.	A. It reacts with skin B. It has lower boiling point and evaporate faster C. It is more viscous D. It absorbs moisture
3	Overall order is sum of	A. Coefficients B. Exponents in rate law C. Moles D. Products
4	A reaction mechanism describes	A. Experimental conditions B. Overall stoichiometry C. Stepwise molecular events D. Heat of reaction
5	An orbital can accommodate how many electrons.	A. 1 B. 2 C. 3 D. 4
6	The mass of one mole of electron is	A. 0.55 mg B. 1.088 mg C. 0.184 mg D. 1.67 mg
7	Which step in the Born-Haber cycle is always endothermic	A. Sublimation B. Electron gain enthalpy C. Hydration D. Lattice formation
8	At equilibrium	A. Products dominate B. Reactants dominate C. Forward and reverse rates are zero D. Forward rate = reverse rate
9	Formula for calculating the value of (Azimuthal quantum number)	A. $n+1$ B. $n-1$ C. $n+2$ D. $2n^2$
10	The termination step in a radical chain reaction involves.	A. Chain breaking B. Light absorption C. Chain initiation D. Hydrogen abstraction
11	Liquids are approximately how many times less compressible than gases.	A. 10 B. 105 C. 100 D. 1000
12	Which of the following best describes the molecular arrangement in liquid crystals.	A. Random and disordered B. Fixed in all directions C. Parallel and partially ordered D. Completely rigid and fixed
13	Which of the following statements correctly describes the effect of temperature on the equilibrium constant.	A. $K_c$ is directly proportional to temperature B. $K_c$ is inversely proportional to temperature. C. $K_c$ depends on the enthalpy change of the reaction D. Temperature has no effect on the value of $K_c$
14	Which of the following is not a factor affecting the rate of a reaction?	A. Temperature and pressure B. Concentration of ions C. Surface area of solid reactants D. Nature of the reactants

14	Cell potential is the difference between	<p>C. Electrode potentials of cathode and anode</p> <p>D. Mass of electrodes</p>
15	Unit of rate constant for zero order reaction is	<p>A. <math>\text{mol dm}^{-3} \text{s}^{-1}</math></p> <p>B. <math>\text{s}^{-1}</math></p> <p>C. <math>\text{mol}^{-1} \text{dm}^3 \text{s}^{-1}</math></p> <p>D. <math>\text{mol}^{-2} \text{dm}^6 \text{s}^{-1}</math></p>
16	Oxidation number of oxygen in $\text{H}_2\text{O}$	<p>A. 1</p> <p>B. -2</p> <p>C. 0</p> <p>D. +1</p>
17	A polymer used for non-stick cookware is.	<p>A. Teflon</p> <p>B. Nylon</p> <p>C. Polyvinyl chloride</p> <p>D. Polythene</p>
18	$\text{sp}^3$ Hybridization is associated with structure.	<p>A. Linear</p> <p>B. Tetrahedral</p> <p>C. Trigonal</p> <p>D. Octahedral</p>
19	Which of the following contains same number of particles as present in 12 g of carbon	<p>A. 28 g of iron (Atomic mass of Fe = 56)</p> <p>B. 48 g of magnesium (Atomic mass of Mg = 24)</p> <p>C. 32 g of S Molecules (Atomic mass S = 32)</p> <p>D. 44 g of carbon dioxide (molar mass of <math>\text{CO}_2</math> = 44)</p>
20	Law of mass action was proposed by	<p>A. Le Chatelier</p> <p>B. Arrhenius</p> <p>C. Guldberg and Waage</p> <p>D. Dalton</p>