

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	How many subshells are present in M-Shell	A. 1 B. 2 C. 3 D. 4
2	Increase in pressur shifts equilibrium to.	A. Side with more moles of gas B. Sides with fewer moles of gas C. Liquid phase D. Solid phase
3	Catalyst used in Haber process is	A. Fe B. Pt C. Cu D. Zn
4	What typ of molecular shape in communly found in liquid crystals.	A. Spherical B. Rod like and elongated C. Spiral D. Cubic
5	Element that can show variable oxidation states.	A. Mg B. Ca C. S D. Mg
6	At equilibrium in a clsoed system, whoih two processes occur at the same rate.	A. Melting and freezing B. Evaporation and boiling C. Evporatin and condensation D. Sublimation and condensation
7	The reversibel reation cannot be achieved in	A. Open system B. Closed system C. Both a and b D. None of these
8	Which of the following materials is an example of an amorphous solid	A. Glass B. lce C. Diamon D. Sodium Chaloride
9	Which of the following atoms show anomalous configuration.	A. Cu B. CO C. Fe D. Ni
10	Which one is a non polar molecule.	A. CS2 B. CaC2 C. H2S D. CaC12
11	Liquid crystals exhibit properties.	A. Only like solids B. Only like liquid C. Between solids and liquids D. Unliek solids or liquids
12	Which one is a polar molecule	A. HCI B. CO2 C. BF3 D. CCI4
13	4.0 g of NaOH (molar mass 40 g mol-1) contains same numebr of sodium ions as are present in	A. 5.3 g of Na2CO 3 B. 58.5 g of NaCl C. 76 g Na2 SO4 D. 8 of NaNO3
14	An increase in entropy favors	A. Non spontaneity B. Disorder C. Order D. Equilibrium
15	What property of wate allows quals life to survive under frozen surface	A. High surface tension B. High heat capacity C. C. location surfaces and insulates water

	>	before D. Low boiling point
16	Born Haber cycle is used to calculate	A. Bond energy B. Heat of hydration C. Lattice energy D. lonization energy
17	No two electrons in an atom have same all four quantum numbers, this is called.	A. Aufbau Principle B. Hund's rule C. Pauli's exclusing principle D. None
18	An ionic compound A+B is most likely to be formed when	A. The ionization eenrgy of A is high and electron affinity of B is low B. The ionization energy of A is low and electron affinity of B is high C. Both the ionization energy of A and electron affinity of B are high D. Both the ionization energy of A and electron affinity of B are low.
19	Which one of the following Statments is incorrect	A. One mole of nitrogen gas contains Avogadro's numebr of moleclues B. One mole of ozone gas contains avogadr's number of molecules C. One mole of ozone contains avogadro's number of O atom D. One mole of ozone contains avogadro's number of O atom D. One moleof hydrogen gas contains Avogadro's numebrof molecules
20	Which of the following has maximum mass	A. 2 moles of P B. 5 moles of H2O C. 2 mole of Na2CO3 D. 1 mole of glucose