

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Electrolysis of $\text{CuSO}_4$ using copper electrodes results in	A. Increase in eletolyte concentration B. Decrease in $\text{Cu}^{2+}$ concentration C. No change in electrolyte composition D. Formation of new compound
2	At equilibrium in a closed system, which two processes occur at the same rate.	A. $\text{Mellting and Freezing}$ B. $\text{Evaporation and condensation}$ C. $\text{Sublimation and condensation}$ D. $\text{Evaporation and boiling}$
3	Standard electrode potential is measured under	A. 100 oC and 1 atm B. 0 oC and 1 M concentration C. 25 oC and 1 M concentration D. 100 oC and 1 M concentration
4	The molecuel with greatest dipole moment is.	A. $\text{H}_2\text{O}$ B. $\text{H}_2\text{S}$ C. $\text{NH}_3$ D. $\text{HF}$
5	Which of the following molecules has zero dipole moment.	A. $\text{BF}_3$ B. $\text{NH}_3$ C. $\text{H}_2\text{O}$ D. $\text{CHCl}_3$
6	Markovnikov's rule is used to predict	A. Major product in addition of $\text{HX}$ to alkene B. Stability of alkenes C. Ptofuvy og gtrr tsfivsl hslohrnsyion D. Hydrolysis of esters
7	For gaseous reactions, rate is often monitored by	A. Mass change B. pH change C. Pressure change D. Temperature change
8	The number of moles of $\text{CO}_2$ which contain 8 g of oxygen.	A. $0.25$ B. $0.5$ C. $1.0$ D. $1.50$
9	Dipole moment of $\text{CO}_2$ is	A. $1.8$ B. $1.94$ C. $1.0$ D. $Zero$
10	$Q < K$ implies.	A. Reaction proceeds forward B. equilibrium is established C. System stops D. Reaction shifts in reverse
11	Unit of $k$ for second order reaction is	A. $\text{mol dm}^{-3}$ B. $\text{s}^{-1}$ C. $\text{mol}^{-1} \text{dm}^3 \text{s}^{-1}$ D. $\text{mol}^{-2} \text{dm}^6 \text{s}^{-1}$
12	Who consider as the father of the periodic table	A. Moseley B. Mendeleev C. Newland D. Dobereiner
13	Modern periodic table provides information about.....elements.	A. $100$ B. $118$ C. $110$ D. $130$
14	When a crystalline solid is broken, it does so along specific planes. These planes are known as.	A. Cleavage planes B. Crystal faces C. Surface planes

D. <p>Growth planes</p>

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15 What happens to the shape of a NaCl crystal when 10% urea is present in its solution.

A. <p>It becomes needle like</p>  
B. <p>It becomes larger</p>  
C. <p>It remains the same</p>  
D. <p>It becomes cubic</p>

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16 What is molecular geometry of  $\text{SO}_4^{2-}$  ion.

A. <p>Tetrahedra</p>  
B. <p>Trigonal Planar</p>  
C. <p>Trigonal pyramidal</p>  
D. <p>Linear</p>

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17 In an acid base titration, the equivalence point is reached when.

A. pH of the solution is 7.0  
B. The indicator changes color  
C. Equal volumes of acid and base have been added  
D. The reaction stops

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18 When the vapour pressure of a liquid equals the external pressure, what phenomenon occurs.

A. <p>Sublimation</p>  
B. <p>Boiling</p>  
C. <p>Freezing</p>  
D. <p>condensation</p>

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19 What are the number of moles of hydrogen atom in 3.2 g of  $\text{CH}_4$ ?

A. <p>0.2</p>  
B. <p>0.4</p>  
C. <p>0.6</p>  
D. <p>0.8</p>

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20 The molecule of ethyne contains how many pi bonds.

A. 2  
B. 1  
C. 3  
D. 0