

## Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	s-subshell hasorbital	A. 1 B. 2 C. 3 D. 4
2	At equilibrium in a clsoed system, which two processes occur at the same rate.	A. Melting and freezing B. Evaporation and boiling C. Evporatin and condensation D. Sublimation and condensation
3	Increase in pressur shifts equilibrium to.	A. Side with more moles of gas B. Sides with fewer moles of gas C. Liquid phase D. Solid phase
4	Dipole moment of CO2 is	A. 1.8 D B. 1.94 D C. 1.0 D D. Zero
5	Which one of the following is not the applicatio of solubility products.	A. Determination of solubility from Ksp     B. Common ion effect     C. Predicting precipitation     D. Dissociation of bonds
6	When the vapour pressure of a liquid equals the external pressure, what phenomenon occurs.	A. Sublimation B. Bioling C. Freezing D. condensation
7	Which one of the following effects explains the cooling during the sudden expasion of a gas.	A. Boyle's Effect B. Joule Thomson Effect C. Charles Effect D. Avogadr's law
8	What happens to the shae of a NaCl crystal when 10% urea is present in its solution.	A. It becomes needle like B. It becomes larger C. It reamains the same D. It becomes cubic
9	pH of 10 <sup>-9</sup> M HCl is	A. 5 B. 9 C. Slightly below 7 D. Exactly -9
10	The sign of Delta H during melting is.	A. Negative B. Positive C. Zero D. Can't be predicted
11	Which of the following atoms show anomalous configuration.	A. Cu B. CO C. Fe D. Ni
12	Tendency to lose electrons called	A. metallic character B. Electron affinity C. loniztion energy D. Electrongativity
13	Overall order is sum of	A. Coefficients B. Exponents in rate law C. Moles D. Products
14	At equilirbium the observabel properties.	A. Keep changing B. Fluctuae randomly C. Remain constant D. Oscillate
15	A negative lattice eneryg temperature	A. Bond breaking B. Energy refresed C. Formation of covalent hand

	D. Spontaneity
Which of these is NOT a property of acids.	<ul><li>A. Conduct electricity</li><li>B. Turn litmus red</li><li>C. Sour taste</li><li>D. Feel slippery</li></ul>
Which of the following is a weak base.	A. KOH B. NaOH C. NH3 D. Ca(OH)2
First order reactions are best studies using.	A. log [A] vs time B. Pressure vs time C. [A] vs time D. Titriation
What is teh oxidation state of sulfur in the sulfate ion SO <sup>-2</sup> 4	A. +4 B. +6 C. +2 D. >0
The ideal gas equation is given as.	A. PV =nRT B. PV=RT C. PV = nRT/V D. None of these
	Which of the following is a weak base.  First order reactions are best studies using.  What is teh oxidation state of sulfur in the sulfate ion SO-24