# Fsc Part 1 Chemistry MCQ's Test For Full Book

| Sr | Questions | Answers Choice |
|---|---|---|
| 1 | A Daniell cell produces electricity through | A. Electrolysis<br>B. Radioactivity<br>C. Spontaneous redox reaction<br>D. Endothermic reaction |
| 2 | The numer of possible isomers of C4H10 is | A. 1<br>B. 2<br>C. 3<br>D. 4 |
| 3 | Which oxidation state will not shown by S | A. -2 <br>B. +4 <br>C. +8 <br>D. +2 |
| 4 | The solubility prodct of Pb SO4 is | A. 1.6 x 10<sup>-8</sup><br>B. 1.4 x 10<sup>-4</sup><br>C. 1.6 x 10<sup>-6</sup><br>D. 1.6 x 10<sup>-2</sup> |
| 5 | Cycloalkanes differ from alkanes ny | A. Being cyclic<br>B. Having hygher rectivity<br>C. Having sigma bonds<br>D. Having delocalized electrons |
| 6 | What is molecular geopetry of $SO_4^{-2}$ ion. | A. Tetrahedra <br>B. Trigonal Planar <br>C. Trigonal pyramidal <br>D. Linear |
| 7 | Octet rule is not obeyed b which of the following compounds during its formation | A. H2O <br>B. NaCl <br>C. NH3 <br>D. PF5 |
| 8 | One mole of SO2 conatains. | A. 6.02 x 10 <sup>23</sup> atoms of oxygen <br>B. 6.022 x 10 <sup>23</sup> atoms of sulphur <br>C. 4 g atoms of SO2 <br>D. 18.1  X 10<sup> 23</sup> moles of  SO2 |
| 9 | Which scientist first time observed eh periodiciy in the elements. | A. Dobereiner <br>B. D. Mendeleev <br>C. j. Newlands  <br>D. L. Meyer |
| 10 | The reaction at the cathode is alwyas | A. Oxidation<br>B. Neutral<br>C. Reduction<br>D. Ionization |
| 11 | Kc is expressed in terms of. | A. Pressure<br>B. Mole fraction<br>C. Concentration<br>D. Volume |
| 12 | Removing a reactant shifts equilibrium to | A. Left<br>B. Right<br>C. No shift<br>D. Depends on temperture |
| 13 | Which of the followng molecules has zero dipole moment. | A. BF3 <br>B. NH3 <br>C. H2O <br>D. CHCl3 |
| 14 | NaOH is not considered an Bronsted Lowry acid because. | A. It's neutral<br>B. It's not soluble<br>C. Doesn't donate H+<br>D. Doesn't produce H+ |

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| 15 | In an electrochecmial cell, which ion migrates to cathode through the salt bridge. | A. Anions<br>B. Cations<br>C. Eletrons<br>D. Protons |
| 16 | The reducing agent in a redox reaction | A. Gains electrons<br>B. Loss electrons<br>C. Acceptw protons<br>D. Donates protons |
| 17 | Which of the following contribtes to water's high viscosity | A. Covalet bonds <br>B. Metallic interactions <br>C. Strong  hydrogen bonding <br>D. Low surface area |
| 18 | Acid rain is mainly due to | A. SO3 and NO3<br>B. CO2<br>C. H2<br>D. CH4 |
| 19 | Active mass means | A. Moles<br>B. Volume<br>C. Mass<br>D. Molar concentration |
| 20 | When a bond is formed | A. Energy is absorbed<br>B. Energy is released<br>C. Delta H is always zero<br>D. No energy change |