

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	What is the first step in the free radical mechanism.	A. Termination B. Propagation C. Initiation D. Elimination
2	The actual bond length in a polar covalent compound is.	A. <p>Lesser than expected</p> B. <p>Greater than expected</p> C. <p>Equal to the expected</p> D. <p>Exactly half of expected</p>
3	Which acid is present in vinegar	A. CH <sub>3</sub> COOH B. HNO <sub>3</sub> C. H <sub>2</sub> SO <sub>4</sub> D. HCl
4	Law of mass action was proposed by	A. Le chatelier B. Arrhenius C. Guldberg and Waage D. Dalton
5	The Bronsted Lowry definition identifies acid as.	A. Electron pair acceptors B. Proton donors C. Proton acceptors D. Electron pair donors
6	Which of the following reactions reaches equilibrium	A. Reversible B. Irreversible C. Combustion D. Neutralization
7	Subshells in an atom are filled with electrons in an increasing order of their energy values called.	A. <p>aufbau principle</p> B. <p>Hund's rule</p> C. <p>Pauli's Exclusion Principle</p> D. <p>None</p>
8	Which metal is the best reducing agent.	A. Cu B. K C. Zn D. Fe
9	Which one is a polar molecule	A. <p>HCl</p> B. <p>CO <sub>2</sub> </p> C. <p>BF <sub>3</sub> </p> D. <p>CCl <sub>4</sub> </p>
10	Born Haber cycle is used to calculate	A. Bond energy B. Heat of hydration C. Lattice energy D. Ionization energy
11	The system that exchanges heat but not mass	A. Closed B. Rigid C. Open D. Isolated
12	Electrochemical series helps in predicting.	A. Rate of reaction B. Type of bond C. Direction of redox reactions D. Heat released
13	Which of the following has a branched chain structure.	A. n - butane B. 2- methylpropane C. Ethane D. Propene
14	k has unit s <sup>-1</sup> for	A. First order B. Second order C. Third order D. zero order
15	What are the number of moles of hydrogen atom in 3.2 g of CH <sub>4</sub> ?	A. <p>0.2</p> B. <p>0.4</p> C. <p>0.6</p> D. <p>0.8</p>

		D. $<p>0.8</p>$
16	First order reactions are best studied using.	A. $\log [A]$ vs time B. Pressure vs time C. $[A]$ vs time D. Titration
17	Which type of intermolecular force is present in all types of molecules regardless of their polarity.	A. $<p>\text{Dipole-dipole forces}</p>$ B. $<p>\text{London dispersion forces}</p>$ C. $<p>\text{Hydrogen bonds}</p>$ D. $<p>\text{Ion-dipole forces}</p>$
18	Highest bond order is found in	A. $<p>\text{O}_2</p>$ B. $<p>\text{O}_2^{-1}</p>$ C. $<p>\text{O}_2^{+2}</p>$ D. $<p>\text{O}_2^{+1}</p>$
19	The product formed when ethene reacts with $\text{H}_2$ in the presence of Ni is.	A. Ethanol B. Ethane C. Acetylene D. Butane
20	The value of 'm' cannot be	A. $<p>\text{Fractional}</p>$ B. $<p>\text{Zero}</p>$ C. $<p>\text{Positive}</p>$ D. $<p>\text{Negative}</p>$