

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Why is water's surface tension so high.	<p>A. Due to its ionic nature</p> <p>B. Because of hydrogen bonding pulling surface molecules downward</p> <p>C. Because it has low boiling point</p> <p>D. Because water molecules are very large</p>
2	If a chemical reaction has $\Delta H = -100 \text{ kJ/mol}$, it is	<p>A. Exothermic</p> <p>B. Endothermic</p> <p>C. Isothermal</p> <p>D. Isobaric</p>
3	What is the value of water's enthalpy of vaporization.	<p>A. 21 kJ/mol</p> <p>B. 41 kJ/mol</p> <p>C. 30 kJ/mol</p> <p>D. 60 kJ/mol</p>
4	How many extra electrons than a normal octet are there in the valence shell of I in ICl_5 ?	<p>A. 2</p> <p>B. 3</p> <p>C. 4</p> <p>D. 5</p>
5	Which one of the following statements is incorrect	<p>A. One mole of nitrogen gas contains Avogadro's number of molecules</p> <p>B. One mole of ozone gas contains Avogadro's number of molecules</p> <p>C. One mole of ozone contains Avogadro's number of O atoms</p> <p>D. One mole of hydrogen gas contains Avogadro's number of molecules</p>
6	Aromatic hydrocarbons must contain.	<p>A. Straight chains</p> <p>B. Double bonds only</p> <p>C. A benzene ring</p> <p>D. A triple bond</p>
7	The mass of one mole of electron is	<p>A. 1.088 mg</p> <p>B. 0.55 mg</p> <p>C. 0.184 mg</p> <p>D. 1.67 mg</p>
8	What is the ratio of volumes of 2 g of O_2 to the volume of 16 g CH_4 both at STP	<p>A. 1:1</p> <p>B. 1:8</p> <p>C. 1:2</p> <p>D. 2:1</p>
9	In the Hall-Héroult process, aluminum is obtained by	<p>A. Chemical reaction</p> <p>B. Thermal decomposition</p> <p>C. Electrolysis of alumina</p> <p>D. Roasting and leaching</p>
10	Na burns with oxygen to give.	<p>A. Golden yellow flame</p> <p>B. Golden brown flame</p> <p>C. Crimson red flame</p> <p>D. Intense white flame</p>
11	The SI unit of enthalpy	<p>A. Calorie</p> <p>B. eV</p> <p>C. kJ mol⁻¹</p> <p>D. J mol⁻¹</p>
12	sp^3 Hybridization is associated with structure.	<p>A. Linear</p> <p>B. Tetrahedral</p> <p>C. Trigonal</p> <p>D. Octahedral</p>
13	At equilibrium ΔG is	<p>A. Zero</p> <p>B. Positive</p> <p>C. Negative</p>

		C. Negative D. Maximum
14	What is the SI unit of viscosity.	A. Pascal B. $\text{kg m}^{-1} \text{s}^{-1}$ C. Joule D. Nm^{-2}
15	A process with increase in entropy and enthalpy is spontaneous at.	A. High temperature B. Low temperature C. All temperature D. never spontaneous
16	The number of atoms present in 6 g of Mg metals.	A. 12.4×10^{23} B. 6.02×10^{23} C. 1.5×10^{23} D. 3.01×10^{23}
17	Which of the following species has unpaired electrons in antibonding molecular orbitals.	A. O_2^{2+} B. N_2^{2-} C. B_2 D. F_2
18	Enthalpy change in hydration depends on.	A. Charge B. Ion size C. Solvent nature D. All of these
19	For a specific reaction the value of the equilibrium constant, K_c ?	A. Always remains the same at different reaction conditions B. Increases if the concentration of one of the products is increased C. Changes with changes in the temperature D. Increases if the concentration of one of the reactants is increased
20	Rate law is an expression involving.	A. Concentration B. Pressure C. Temperature D. Catalyst