

Chemistry Fsc Part 1 Online Test

Sr	Questions	Answers Choice
1	If Kc> 1, the reaction.	A. Favors products B. Favors reactants C. Is at equilirbium D. Does not occur
2	Which of the following is NOT a property of liqud crystals.	A. Fludity B. Rigidity like crystals C. Complete isotropy D. Viscosity
3	Why does ice float on water.	A. It has lower density than water B. It is heavier than water C. It forms ionic bonds D. It denser than water
4	The quantum numebr values for the fourth electron of Be atom are.	A. 1,0,0 B. 2,0,0 C. 2,1,0 D. 1,1,1
5	Which of the following is essential for hydrogen bond formation.	A. Hydrogen bonded to a metal B. Hydrogen bonded to high electongtive atoms F,O, or N C. Hydogen boned to non pola ratom D. Present of pi bond
6	Contact process is used for	A. Sulfuric acid productionB. Ammonia SynthesisC. Nitric Acid productionD. Hydrogenation
7	An acid and base react to give aacid and base	A. Amphoteric B. Salt C. Conjugate D. Neutral
8	In a reversible reaction.	A. Products do not reform reactants B. Rate of forward reaction is always higher C. Both forward and reverse reactions occur D. Products are in excess
9	Paramagnetic specie is.	A. N2 B. N2 ⁻¹ C. O2 ⁺² D. O2 ⁻²
10	Enthalpy change of a process is measured under	A. Constant volume B. Constant pressure C. Constant temperature D. Constant Energy
11	Kp is used for.	A. Solid state reactions B. Gaseous reactions C. Aqueous reactions D. Liquid reactions
12	The ratio of numebr of molecules in 71 g of gaseous chlroine to the numebr of molecules in 2 g of gaseous hydrogen is	A. 1:1 B. 71:2 C. 1:2 D. 71:1
13	Liquid are approximately how many times less compressible than gases.	A. 10 B. 105 C. 100 D. 1000
14	Which of the following moleules has zero dipole moment.	A. BF3 B. NH3 C. CHCI3

		υ. H2U
15	In a chemical reaction, a catalyst.	A. Alters Delta H B. Lowers activation energy C. Is consumed D. Forms product
16	An ionic compound A+B is most likely to be formed when	A. The ionization eenrgy of A is high and electron affinity of B is low B. The ionization energy of A is low and electron affinity of B is high C. Both the ionization energy of A and electron affinity of B are high D. Both the ionization energy of A and electron affinity of B are low.
17	L- Pauling developed a scale for	A. lonization energy B. Electron affinity C. Electronegativity D. None
18	On Boltzmann distribution curve, the area under the curve represents.	A. Activation energy of the rreaction B. Total numebr f moleclues in the sample C. Average kinetic energy of the molecules D. Rate constant of the reaction
19	A reaction mechanism describes	A. Experimental conditions B. Overall stoichiometry C. Stepwise molecular events D. Heat of reaction
20	Which one is a Lewis base.	A. H+ B. NH3 C. HCI D. BF3