

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	The quantum number 'm' of a free gascous atom is associated with	<p>A. <p>The effective volume of the orbital</p></p> <p>B. <p>The shape of the orbital</p></p> <p>C. <p>The spatial orientation of the orbital</p></p> <p>D. <p>The energy of the orbital in the absence of a magnetic field</p></p>
2	Markovnikov's rule is used to predict	<p>A. Major product i addition of HX to alkene</p> <p>B. Stability of alkenes</p> <p>C. Ptovuvy og gtrr tsfivsl hslohrnsyion</p> <p>D. Hydrolysis of esters</p>
3	Homogenous catalysts are in	<p>A. Same phase as reactions</p> <p>B. Different phase</p> <p>C. Solid only</p> <p>D. Gaseous only</p>
4	Vulcanization of rubber increases its.	<p>A. Conductivity</p> <p>B. Reactivity</p> <p>C. Hardness and elasticity</p> <p>D. Transparency</p>
5	The conjugate acid of NH ₃ is	<p>A. NH₄</p> <p>B. NH₂</p> <p>C. NO₃</p> <p>D. N₂H₄</p>
6	Redox reactions always involve.	<p>A. Gain of protons</p> <p>B. Transfer of electrons</p> <p>C. Loss of neutrons</p> <p>D. Nuclear change</p>
7	According to Avodgaro's Law, volume is directly proportional to.	<p>A. <p>Pressure</p></p> <p>B. <p>Temperature</p></p> <p>C. <p>Number of moles</p></p> <p>D. <p>Density</p></p>
8	Chemical bond formation taken pace when	<p>A. <p>Force of attraction are equal to the force of repulsion</p></p> <p>B. <p>Force of repulsion is greater than force of attraction</p></p> <p>C. <p>Force of attraction overcomes force of repulsion</p></p> <p>D. <p>None of these</p></p>
9	Which component maintains eleclrical neutrality in an electronchemical cell	<p>A. Salt bridge</p> <p>B. Voltmeter</p> <p>C. Electrolyte</p> <p>D. Electrodes</p>
10	Hydrocarbon are compounds made of.	<p>A. Hydrogen and oxygen</p> <p>B. Carbon and hydrogen</p> <p>C. Carbon , hydrogen, and oxygen.</p> <p>D. Carbon and oxygen</p>
11	Catalyst used in contact process	<p>A. Fe</p> <p>B. V₂O₅</p> <p>C. Ni</p> <p>D. Al₂O₃</p>
12	Which of the following is a polar moleuce taht shows permanent dipole interactions.	<p>A. <p>CHCl</p></p> <p>B. <p>CH₄</p></p> <p>C. <p>Cl₂</p></p> <p>D. <p>H₂</p></p>
13	Which of the following is an irreversible reaction.	<p>A. Haber process</p> <p>B. Precipitation of Ag Cl</p> <p>C. Synthesis of ammonia</p> <p>D. Esterification</p>
14	The numebr of molesof CO ₂ which contain 8 g of oxygen	<p>A. <p>0.25</p></p> <p>B. <p>0.5</p></p> <p>C. <p>1.0</p></p>

D. $$1.50$$

15 Increase in concentration of reactants

- A. Increase K
- B. Decrease K
- C. Shifts equilibrium forward
- D. Stops reverse reaction

16 Highest electronegative element is.

- A. F
- B. Cl
- C. I
- D. Br

17 The number of moles of CH₄ in 4 g of gas is

- A. 0.2 moles
- B. 0.25 mole
- C. 0.3 mol
- D. 0.4 mole

18 CH₃COONa in water forms

- A. Acidic solution
- B. Basic Solution
- C. Neutral Solution
- D. Salt bridge

19 Liquid crystals exhibit properties of

- A. Only solid
- B. Only liquid
- C. Both solids and liquids
- D. Gases and solids

20 d-block elements are also called.

- A. Inner transition
- B. Outer transition
- C. Typical transition
- D. None