

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which one of the following compound is added in purification of NaCl in common ion effect.	A. HCl B. H ₂ SO ₄ C. HNO ₃ D. HF
2	Which factor affects lattice energy	A. Ion size B. Ion charge C. Crystal structure D. None of these
3	In electrolysis, reduction always occurs at	A. Anode B. Cathode C. Salt bridge D. Electrolyte
4	Liquid crystals are used in thermometers because.	A. They are highly reflective B. They glow at high temperatures C. Their color changes with temperature D. They conduct electricity
5	Acidic chlorides have pH?	A. More than 7 B. Less than 7 C. Equal to 7 D. May be any
6	An increase in temperature generally	A. No Effect B. Increases cell voltage C. Decreases cell voltage D. Depends on pressure
7	K _c is expressed in terms of.	A. Pressure B. Mole fraction C. Concentration D. Volume
8	What type of reaction has a negative ΔH and ΔS?	A. Always spontaneous B. Never spontaneous C. Spontaneous at low temp D. Spontaneous at high temp
9	The order of a chemical reaction, that is dependent of concentration is.	A. Second order reaction B. First order reaction C. Zero order reaction D. Pseudo first order reaction
10	What happens to the shape of a NaCl crystal when 10% urea is present in its solution.	A. It becomes needle like B. It becomes larger C. It remains the same D. It becomes cubic
11	Increase in concentration of reactants	A. Increase K B. Decrease K C. Shifts equilibrium forward D. Stops reverse reaction
12	Alkenes undergo which type of reaction.	A. Nucleophilic addition B. Electrophilic addition C. Free radical substitution D. Nucleophilic substitution
13	A solution contains 4.0 g of sodium hydroxide in 250 cm ³ of solution. What is the molar concentration of this solution	A. 0.10 mol dm ⁻³ B. 0.20 mol dm ⁻³ C. 0.40 mol dm ⁻³ D. 0.80 mol dm ⁻³
14	Which of the following is a weak base.	A. KOH B. NaOH C. NH ₃ D. Ca(OH) ₂
15	Which one of the following effects explains the cooling during the sudden expansion of a gas?	A. Boyle's Effect B. Joule Thomson Effect C. Charles Effect

a gas.

C. $p \propto \frac{1}{V}$
D. Avogadro's law

16 Geometrical isomerism is possible in.

A. Alkanes
B. Alkene
C. Cycloalkanes
D. Both b and c

17 Lysosomes are formed by budding from which cellular organelle.

A. Smooth endoplasmic reticulum
B. Golgi apparatus
C. Rough endoplasmic reticulum
D. Nucleus

18 A solution with equal $[H^+]$ and $[OH^-]$ is

A. Basic
B. Acidic
C. Neutral
D. Salt

19 Which of the following atoms show more than one oxidation states.

A. Phosphorous
B. Magnesium
C. Sodium
D. Aluminium

20 Who is considered as the father of the periodic table

A. Moseley
B. Mendeleev
C. Newland
D. Dobereiner