

Chemistry Fsc Part 1 Online Test

C.	Quantiana	Answers Choice
Sr 1	Questions Active mass means	A. Moles B. Volume
I	Active mass means	C. Mass D. Molar concentration
2	A saturted solution represents a dynamic equilibrium Macroscopically, the concentration of dissolved solute is constant, Microscopically this occurs because.	 A. No more solute particles are dissolving B. The rate of dissolution of solute is zero C. Solute particles are dissolving and precipitating at the same rate D. All solute particles have dissolved
3	Effective collisions are those that.	A. Have enough energy and proper orientation B. Are elastic C. Are in vacuum D. Produce no change
4	The standard enthalpy of atomization of an element is alwyas.	A. Negative B. Positive C. Zero D. Depend on element
5	Liquid are approximately how many times less compressible than gases.	A. 10 B. 105 C. 100 D. 1000
6	The molecuel with greatest dipole moment is.	A. H2O B. H2S C. NH3 D. HF
7	Half life formula for 1st order reaction is	A. 0.693 /k B. k x t C. 1 /k D. 2 k
8	k has unit s-1 for	A. First order B. Second order C. Third order D. zero order
9	The rate of reaction .	A. Increase as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reaction proceeds D. May decrease or increse as the reaction proceeds
10	What happens to the shae of a NaCl crystal when 10% urea is present in its solution.	A. It becomes needle like B. It becomes larger C. It reamains the same D. It becomes cubic
11	Peroxides are those in whoih oxidation state of oxygen is	A1 B. +2 D2 E1/2
12	Heat of combustion of H2 is used to determine	A. Calorific value B. Enthalpy C. lonization energy D. Lattice energy
13	Highest electonegatie element is.	A. F B. Cl C. I D. Br
14	The rate determinig step in a multistep reaction is	A. The slowest step B. Always the first step C. Always the last step

		D. The tastest step
15	At equilibrium in a closed system, which two processes occur at tha same rate.	A. Mellting and Freezing B. Evporation and condensation C. Sublimation and condensation C. Sublimation and condensation D. Evaporation and boiling
16	Consider the gas phase equilibrium sysem represented by teh equation 2 H2O2H2+O2 Givent that the forward reactio is endothermic, which of the following chagnes will decrease the equilibrium amount of H2O.	A. Adding more oxygen B. Adding a solid phase catalyst C. Decreasing the volume of the container D. Increaseing the temperature at constant pressure
17	Rate of fastest	A. At start B. In middle C. At end D. After completion
18	Surface tension usually with interasig temperature.	A. Increase B. Decrease C. Remain constant D. Doubles
19	Litmus turns which color in a basic solution.	A. Red B. Orange C. Blue D. Colorless
20	Which is amphoteric	A. H2O B. HCI C. NaOH D. NH3