

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Two electrons in the same orbital have.	<p>A. Same Spin</p> <p>B. Different spin</p> <p>C. May same or different</p> <p>D. Parallel spin</p>
2	The mass of molecules of CO	<p>A. 0.5 moles</p> <p>B. 2.0 moles</p> <p>C. 1.0 moles</p> <p>D. 1.06 moles</p>
3	In which subshell electrons are filled first according to n+1 rule.	<p>A. 2p</p> <p>B. 3p</p> <p>C. 4p</p> <p>D. 4s</p>
4	A negative lattice energy temperature	<p>A. Bond breaking</p> <p>B. Energy released</p> <p>C. Formation of covalent bond</p> <p>D. Spontaneity</p>
5	A process with increase in entropy and enthalpy is spontaneous at.	<p>A. High temperature</p> <p>B. Low temperature</p> <p>C. All temperature</p> <p>D. never spontaneous</p>
6	In redox reactions, oxidizing agents are	<p>A. reduced</p> <p>B. Oxidized</p> <p>C. Always metals</p> <p>D. Always gases</p>
7	Collision frequency depends on	<p>A. Pressure</p> <p>B. Concentration</p> <p>C. Temperature</p> <p>D. All of these</p>
8	Which of the following molecules has zero dipole moment.	<p>A. <math>\text{BF}_3</math></p> <p>B. <math>\text{NH}_3</math></p> <p>C. <math>\text{CHCl}_3</math></p> <p>D. <math>\text{H}_2\text{O}</math></p>
9	Rate of reaction is measured as.	<p>A. Increase Temperature</p> <p>B. Change in concentration per unit time</p> <p>C. Total energy of reactants</p> <p>D. Change in volume per mole</p>
10	If Zn -Cu galvanic cell works ideally after complete discharge, both compartments will have	<p>A. <math>\text{CuSO}_4</math> Solution</p> <p>B. <math>\text{ZnSO}_4</math> Solution</p> <p>C. Cu ions</p> <p>D. Zn Solid</p>
11	Number of lone pair of electrons in ammonium ion is	<p>A. One</p> <p>B. Two</p> <p>C. Three</p> <p>D. Zero</p>
12	Geometrical isomerism is possible in.	<p>A. Alkanes</p> <p>B. Alkene</p> <p>C. Cycloalkanes</p> <p>D. Both b and c</p>
13	Which of the following will show paramagnetic property.	<p>A. <math>\text{O}_2</math></p> <p>B. <math>\text{N}_2</math></p> <p>C. <math>\text{O}_2^{2-}</math></p> <p>D. <math>\text{O}_2^{+2}</math></p>
14	$K_p$ is used for.	<p>A. Solid state reactions</p> <p>B. Gaseous reactions</p> <p>C. Aqueous reactions</p> <p>D. Liquid reactions</p>
15	Which of the following is the conjugate base of water.	<p>A. <math>\text{OH}^-</math></p> <p>B. <math>\text{OH}^+</math></p> <p>C. <math>\text{H}_2\text{O}</math></p>

		D. $\text{H}_2\text{O}^+$
16	Effect of temperature on rate is	A. Negligible B. Linear C. Exponential D. Logarithmic
17	The units of the rate constant (k) for a reaction depend on the	A. Activation energy of the reaction B. Temperature of the reaction C. Overall order of the reaction D. Stoichiometry of the balanced chemical equation
18	Which dialyzing solution has highest pH	A. Vinegar B. Lemon juice C. Ammonia solution D. Coffee
19	Born Haber cycle is used to calculate	A. Bond energy B. Heat of hydration C. Lattice energy D. Ionization energy
20	What is the correct IUPAC name of $\text{CH}_3\text{CH}=\text{CH}_2$ ?	A. Propene B. Propane C. Ethene D. Butene