

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Enthalpy of fusion is the heat required to.	<p>A. Melt a solid                      B. Boil a liquid                      C. Freez a liquid                      D. Vaporize a solid</p>
2	Which of the following is an extensive property.	<p>A. Enthalpy                      B. Temperature                      C. Density                      D. Pressure</p>
3	When does a liquidi start boiling.	<p>A. &lt;p&gt;When temperatur reahes 100 oC&lt;/p&gt;                      B. &lt;p&gt;When its vaporu pressur eequals atmospheriic pressure&lt;/p&gt;                      C. &lt;p&gt;When all molecules turn into vapour&lt;/p&gt;                      D. &lt;p&gt;When surface tension becomes zero&lt;/p&gt;</p>
4	The most electronegative atom is	<p>A. &lt;p&gt;F&lt;/p&gt;                      B. &lt;p&gt;Cl&lt;/p&gt;                      C. &lt;p&gt;N&lt;/p&gt;                      D. &lt;p&gt;O&lt;/p&gt;</p>
5	Which theory explain the paramagnetic behavior exhibited by O <sub>2</sub> molecule.	<p>A. &lt;p&gt;Band Theory&lt;/p&gt;                      B. &lt;p&gt;VBT&lt;/p&gt;                      C. &lt;p&gt;MOT&lt;/p&gt;                      D. &lt;p&gt;VSEPR Theory&lt;/p&gt;</p>
6	The enthalpy change when one mole of ionic compound is dissolved in water is	<p>A. Het of hydration                      B. Heat of solution                      C. Heat of combustion                      D. Heat of atomization</p>
7	Which of the followng has the lowest molar hat of vaporization.	<p>A. &lt;p&gt;HCl&lt;/p&gt;                      B. &lt;p&gt;Water&lt;/p&gt;                      C. &lt;p&gt;NH<sub>3</sub>&lt;/p&gt;                      D. &lt;p&gt;Ethanol&lt;/p&gt;</p>
8	Which component maintains elecrical neutrality in an electronchemical cell	<p>A. Salt bridge                      B. Voltmeter                      C. Electrolyte                      D. Electrodes</p>
9	The metal deposited at cathode during electolysis of NaCl is.	<p>A. Cl<sub>2</sub>                      B. Na                      C. H<sub>2</sub>                      D. Cu</p>
10	The molecule with linear structure is	<p>A. &lt;p&gt;H<sub>2</sub>O&lt;/p&gt;                      B. &lt;p&gt;CO<sub>2</sub>&lt;/p&gt;                      C. &lt;p&gt;H<sub>2</sub>S&lt;/p&gt;                      D. &lt;p&gt;BF<sub>3</sub>&lt;/p&gt;</p>
11	Subshells in an atom are filled with electrons in an increasing order of their energy values called.	<p>A. &lt;p&gt;aufbau priciple&lt;/p&gt;                      B. &lt;p&gt;Hund's rule&lt;/p&gt;                      C. &lt;p&gt;Pauli's Exclusion Principle&lt;/p&gt;                      D. &lt;p&gt;None&lt;/p&gt;</p>
12	What is the ratio of volumes of 2 g of O <sub>2</sub> to the volum of 16 g CH <sub>4</sub> both volume areat STP	<p>A. &lt;p&gt;1:1&lt;/p&gt;                      B. &lt;p&gt;1:8&lt;/p&gt;                      C. &lt;p&gt;1:2&lt;/p&gt;                      D. &lt;p&gt;2:1&lt;/p&gt;</p>
13	The SI unit of enthalpy	<p>A. Calorie                      B. eV                      C. kJ mol<sup>-1</sup>                      D. J mol<sup>-1</sup></p>
14	The numebr of coulombs required to deposit 1 mole of Ag	<p>A. 96500 C                      B. 193000 C                      C. 1 C                      D. 241250 C</p>

15	Reaction of ethene with cold, dilute alkaline $\text{KMnO}_4$ gives.	A. $\text{CO}_2$ B. Ethanoic acid C. Ethylene glycol D. Ethanol
16	What is molecular geometry of $\text{SO}_4^{2-}$ ion.	A. Tetrahedral B. Trigonal Planar C. Trigonal pyramidal D. Linear
17	Structural isomers differ in	A. Only the boiling point B. Only the melting point C. Arrangement of atoms D. Number of atoms
18	If we remove electrons from an atom until the nucleus is left, then it is called.	A. Ionization energy B. Ionization potential C. Successive ionization energy D. All
19	Which one is the correct statement among the following	A. Anionic radius is generally smaller than atomic radius B. Cationic radius is generally bigger than atomic radius C. Cationic radius is generally smaller than atomic radius D. Both anionic and cationic radii are smaller than atomic radius
20	Which of the following has maximum mass	A. 2 moles of P B. 5 moles of $\text{H}_2\text{O}$ C. 2 moles of $\text{Na}_2\text{CO}_3$ D. 1 mole of glucose