

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which of the following is an aromatic hydrocarbon	A. Benzene B. Ethene C. Methane D. Cyclohexane
2	Why does acetone feel colder than water when poured on skin.	A. <p>It reacts with skin</p> B. <p>It has lower boiling point and evaporate faster</p> C. <p>It is more viscous</p> D. <p>It absorbs moisture</p>
3	Which one of the following gases will have maximum volume at STP.	A. <p>88 g of N₂O</p> B. <p>22 g of CO₂</p> C. <p>28 g of CO</p> D. <p>28 g of N₂</p>
4	A base that ionizes completely in aqueous solution is called.	A. Strong base B. Weak base C. Conjugate acid D. Neutral
5	The shape in which a crystal usually grows is called its.	A. <p>Habit</p> B. <p>Crystal lattice</p> C. <p>Melting Point</p> D. <p>Cleavage plane</p>
6	What is the first step in the electrophilic addition reaction of alkenes.	A. Formation of carbocation B. Attack by nucleophile C. <p>Attack by an electrophile on the double bond</p> D. Formation of a free radical
7	Which dialyzing solution has highest pH	A. Vinegar B. Lemon juice C. <p>Ammonia solution</p> D. Coffee
8	Which of the following is a characteristic of crystalline solids.	A. <p>They do not have a definite geometrical shape</p> B. <p>They melt over a wide temperature range</p> C. <p>They have a sharp melting point</p> D. <p>Their properties do not depend on the direction of measurement</p>
9	Consider a reaction with $\Delta H > 0$ and $\Delta S < 0$ this reaction will be	A. Spontaneous at all temperatures B. <p>Non spontaneous at all temperature</p> C. Spontaneous only at high temperatures D. Spontaneous only at low temperatures
10	Why is water's surface tension so high.	A. <p>Due to its ionic nature</p> B. <p>Because of hydrogen bonding pulling surface molecules downward</p> C. <p>Because it has low boiling point</p> D. <p>Because water molecules are very large</p>
11	Rate determining step controls	A. Overall rate B. Temperature C. Volume D. Final yield
12	Which is the correct trend in variation of electronegativity along a period of the periodic table.	A. <p>It decreases from left to right across a period</p> B. <p>It increases from left to right across a period</p> C. <p>It remains constant</p> D. <p>It has no definite trend</p>

13	The test for unsaturation in a hydrocarbon that disappears the colour is.	A. Reaction with KMnO4 B. Reaction with NaOH C. Decolorization of Br ₂ Water D. Precipitate with Ag NO ₃
14	The number of possible isomers of C ₄ H ₁₀ is	A. 1 B. 2 C. 3 D. 4
15	Which one of the following compound is added in purification of NaCl in common ion effect.	A. HCl B. H ₂ SO ₄ C. HNO ₃ D. HF
16	A positive value for the standard electrode potential of a metal ion metal half cell indicates that	A. The metal is a strong reducing agent B. The metal ion is readily oxidized C. The metal ion is readily reduced D. The metal will readily displace hydrogen from dilute acids.
17	Covalent bonds are	A. <p>Right and directional</p> B. <p>Non rigid and directional</p> C. <p>Rigid and non directional</p> D. <p>Non rigid and non directional</p>
18	The number of lone pairs of electrons in ammonium ion is	A. <p>One</p> B. <p>Two</p> C. <p>Three</p> D. <p>Zero</p>
19	The reversible reaction cannot be achieved in	A. Open system B. Closed system C. Both a and b D. None of these
20	Increasing temperature favors	A. Exothermic reaction B. Endothermic reaction C. Formation of solid D. Reverse in all cases
