

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	If salt bridge is not used, between two half cells in a Galvanic cell, then the voltage.	A. Decrease slowly B. Decrease rapidly C. Does not change D. Drope to zero
2	A gas has a density of 1.43 g dm ³ at STP. What is the molar mass of the gas?	A. <p>14.3 g mol ⁻¹ </p> B. <p>32.0 g mol ⁻¹ </p> C. <p>22.4 g mol ⁻¹ </p> D. <p>64.0 g mol ⁻¹ </p>
3	Which of the following best describes the molecular arrangement in liquid crystals.	A. <p>Random and disordered</p> B. <p>Fixed in all directions</p> C. <p>Parallel and partially ordered</p> D. <p>Completely rigid and fixed</p>
4	What is the first step in the free radical mechanism.	A. Termination B. Propagation C. Initiation D. Elimination
5	What is bond angle in NF ₃ ?	A. <p>102^o</p> B. <p>109.5^o</p> C. <p>104.5^o</p> D. <p>107.5^o</p>
6	In an acid base titration, the equivalence point is reached when.	A. pH of the solution is 7.0 B. The indicator changes color C. Equal volumes of acid and base have been added D. The reaction stops
7	If Zn -Cu galvanic cell works ideally after complete discharge, both compartments will have	A. CuSO ₄ Solution B. Zn SO ₄ Solution C. Cu ions D. Zn Solid
8	The molecule of ethyne contains how many pi bonds.	A. 2 B. 1 C. 3 D. 0
9	The enthalpy of solution is	A. Heat evolved /absorbed when 1 mole of solute dissolves B. Solute dissolves C. Always exothermic D. always endothermic
10	Which of the following hybridizations is found in ethyne.	A. <p>Sp</p> B. <p>sp ² </p> C. <p>sp ³ </p> D. <p>No hybridization </p>
11	Which of the following is always negative in an exothermic reaction.	A. Delta H B. Activation energy C. Entropy D. Delta S
12	Law of mass action applies to	A. Closed systems only B. Open systems C. Irreversible systems D. Combustion reactions
13	A saturated solution represents a dynamic equilibrium Macroscopically, the concentration of dissolved solute is constant, Microscopically this occurs because.	A. No more solute particles are dissolving B. The rate of dissolution of solute is zero C. Solute particles are dissolving and precipitating at the same rate D. All solute particles have dissolved
14	Why does acetone feel colder than water when poured on skin.	A. <p>It reacts with skin</p> B. <p>It has lower boiling point and evaporate faster</p>

		<p>C. ρ is more viscous</p> <p>D. ρ absorbs moisture</p>
15	Aromatic hydrocarbons must contain.	<p>A. Straight chains</p> <p>B. Double bonds only</p> <p>C. A benzene ring</p> <p>D. A triple bond</p>
16	The pH of a buffer doesn't change much when	<p>A. Diluted with water</p> <p>B. Acid or base is added</p> <p>C. Heated</p> <p>D. Temp changes</p>
17	If $Q=K$, then	<p>A. Reaction is irreversible</p> <p>B. System is at equilibrium</p> <p>C. Forward reaction dominates</p> <p>D. Reverse reaction dominates</p>
18	s-subshell has.....orbital	<p>A. 1</p> <p>B. 2</p> <p>C. 3</p> <p>D. 4</p>
19	Unit of k for second order reaction is	<p>A. mol dm^{-3}</p> <p>B. s^{-1}</p> <p>C. $\text{mol}^{-1} \text{dm}^3 \text{s}^{-1}$</p> <p>D. $\text{mol}^{-2} \text{dm}^6 \text{s}^{-1}$</p>
20	What happens to the volume of water when it freezes	<p>A. It contracts by 9%</p> <p>B. It remains the same</p> <p>C. It expands by 9%</p> <p>D. It evaporates completely</p>