

## Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	If the electronegativity difference between bonded atoms is more than 0.5 but less than 1.7, the bond will be.	A. <p>Ionic</p> B. <p>Polar Covalent</p> C. <p>Non polar covalent</p> D. <p>Coordinate covalent</p>
2	CH <sub>3</sub> COONa is a salt of.	A. Strong acid+ Strong base B. Weak acid+Strong base C. None D. Weak base+Weak acid
3	Group 16 elements are also called.	A. <p>Chalcogens</p> B. <p>Pnictogens</p> C. <p>Alkali metals</p> D. <p>Halogens</p>
4	Which of these monomers forms a condensation polymer.	A. Ethane B. Styrene C. Adipic acid+ hexamethylenediamine D. Vinyl chloride
5	For gaseous reactions, rate is often monitored by	A. Mass change B. pH change C. Pressure change D. Temperature change
6	What is the IUPAC name of CH <sub>3</sub> CH <sub>2</sub> CH <sub>3</sub> ?	A. Methane B. Propane C. Butane D. Ethane
7	Atoms having same atomic mass but different atomic number are called.	A. <p>Isobars</p> B. <p>Isotopes</p> C. <p>Isotones</p> D. <p>Isoelectrons</p>
8	The activity series of metals arranges metals in order of their.	A. Atomic mass B. Density C. Ease of oxidation D. Ease of reduction
9	Which of the following is a copolymer.	A. SBR Rubber B. Polyethene C. PVC D. Polypropylene
10	What happens to the volume of water when it freezes.	A. <p>Because F is less electronegative</p> B. <p>Because it is non polar</p> C. <p>Because hydrogen bonding in HF traps the H <sup>+</sup> ion</p> D. <p>Because HF is a gas</p>
11	A sample of nitrogen gas has a mass of 14.0 g. How many nitrogen atoms are present in this sample?	A. <p>3.01 x 10 <sup>25</sup> atoms</p> B. <p>6.02 x 10 <sup>23</sup> atoms</p> C. <p>12.01 x 10 <sup>25</sup> atoms</p> D. <p>2.40 x 10 <sup>25</sup> atoms</p>
12	Electrolysis requires.	A. Redox reaction B. Spontaneous cell C. Non spontaneous reaction D. Combustion
13	Which element has E <sub>0</sub> = 0.00 V?	A. H <sup>+</sup> B. H <sub>2</sub> C. SHE D. All of the above
14	A zero order reaction has rate	A. Proportional to [A] B. Independent of [A] C. Proportional to time

D. Zero

15 Which compound is a cycloalkane.

A. Butene  
B. Cyclopentane  
C. Benzene  
D. Cyclopentane

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16 Electronic cloud around teh nucleus is uniformly distributed in all directions in.....subshell

A.  $f$   
B.  $s$   
C.  $d$   
D.  $p$

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17 If  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ , IT IS

A. Strong acid  
B. Base  
C. Weak acid  
D. Neutral

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18 What is the role of a pressur ecooker in boiling water.

A.  $It$  lowers atmospheric pressure  
B.  $It$  allows water to evaporate quickly  
C.  $It$  increases external pressure raising boiling point  
D.  $It$  cools the steam

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19 A solution contains 4.0 g of sodium hydroxide in 250 cm<sup>3</sup> of solution. What is the molar concentration of this solution

A. 0.10 mol dm<sup>-3</sup>  
B. 0.20 mol dm<sup>-3</sup>  
C. 0.40 mol dm<sup>-3</sup>  
D. 0.80 mol dm<sup>-3</sup>

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20 Which is NOT a feature of dynamic equilibrium.

A. Closed system  
B. Constant Temperature  
C. Unequal reaction rates  
D. No net change