

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	The structure of methane contains how many sigma bonds.	A. 1 B. 3 C. 4 D. 2
2	K _p is used for.	A. Solid state reactions B. Gaseous reactions C. Aqueous reactions D. Liquid reactions
3	Number of moles in 100 g of KClO ₃	A. 0.76 B. 0.56 C. 0.816 D. 0.014
4	A zero order reaction has rate	A. Proportional to [A] B. Independent of [A] C. Proportional to time D. Zero
5	Which one is a monoprotic acid	A. H ₃ PO ₄ B. H ₂ SO ₄ C. HCl D. H ₂ CO ₃
6	Which one of the following compound is added in purification of NaCl in common ion effect.	A. HCl B. H ₂ SO ₄ C. HNO ₃ D. HF
7	Which factor affects rate of reaction.	A. Concentration B. Temperature C. Surface Area D. All of these
8	Which of the following has the lowest boiling point among pentane isomers.	A. 2,2-Dimethylpropane B. n-Pentane C. 2-Methylbutane D. Isohexane
9	The value of 'm' cannot be	A. Fractional B. Zero C. Positive D. Negative
10	Enthalpy change of a process is measured under	A. Constant volume B. Constant pressure C. Constant temperature D. Constant Energy
11	The shape of the ethane molecule is.	A. Planar B. Tetrahedral C. Linear D. Trigonal planar
12	Litmus turns which color in a basic solution.	A. Red B. Orange C. Blue D. Colorless
13	In an electrochemical cell, which ion migrates to cathode through the salt bridge.	A. Anions B. Cations C. Electrons D. Protons
14	How many electrons present in Cl are.	A. 17 B. 18 C. 19 D. 35
15	Modern periodic table provides information about.....elements.	A. 100 B. 118 C. 110 D. 130

16	What type of structure do water molecules form in ice.	<p>A. Linear</p> <p>B. Hexagonal close packed</p> <p>C. Regular tetrahedral</p> <p>D. Irregular amorphous</p>
17	Aromatic hydrocarbons must contain.	<p>A. Straight chains</p> <p>B. Double bonds only</p> <p>C. A benzene ring</p> <p>D. A triple bond</p>
18	Na burns with oxygen to give.	<p>A. Golden yellow flame</p> <p>B. Golden brown flame</p> <p>C. Crimson red flame</p> <p>D. Intense white flame</p>
19	Which increases with rise in temperature?	<p>A. Activation energy</p> <p>B. Enthalpy</p> <p>C. Collision frequency</p> <p>D. Molecular weight</p>
20	pH of 1×10^{-4} M HCl is	<p>A. 4</p> <p>B. 1</p> <p>C. 3.5</p> <p>D. 7</p>