

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Which is not a path function	A. Work B. Heat C. Entropy D. Temperature
2	Select metalloid from the following.	A. Po B. Cl C. Fe D. P
3	The term "lattice energy" is applicable to	A. Ionic compounds B. Covalent compounds C. Gases D. Metals
4	What is molecular geometry of SO_4^{2-} ion.	A. Tetrahedral B. Trigonal Planar C. Trigonal pyramidal D. Linear
5	Standard conditions include all except	A. 298 K B. 0 °C C. 1 atm D. 1 M concentration
6	A large K value indicates.	A. Products are favored B. Reactants are favored C. No reaction D. Slow reaction
7	An increase in entropy favors	A. Non spontaneity B. Disorder C. Order D. Equilibrium
8	Which of the following is a polar molecule that shows permanent dipole interactions.	A. CHCl_3 B. CH_4 C. Cl_2 D. H_2
9	Which of the following will show paramagnetic property.	A. O_2 B. N_2 C. O_2^{2-} D. O_2^{+2}
10	Which of the following has highest ionization energy.	A. P B. S C. Ar D. Cl
11	When a bond is formed	A. Energy is absorbed B. Energy is released C. ΔH is always zero D. No energy change
12	Why does ice float on water.	A. It has lower density than water B. It is heavier than water C. It forms ionic bonds D. It is denser than water
13	Which law states that enthalpy change is independent of the path taken.	A. Hess's law B. Boyle's law C. Avogadro's law D. Dalton's law
14	Earthenware pots keep water cool because.	A. They are made of clay B. They reflect sunlight C. They absorb water D. They allow water to evaporate through pores
15	Which is a weak acid	A. HNO_3 B. H_2SO_4 C. H_2CO_3 D. HCl

		<p>C. HF D. HCl</p>
16	If a chemical reaction has $\Delta H = -100 \text{ kJ/mol}$, it is	<p>A. Exothermic B. Endothermic C. Isothermal D. Isobaric</p>
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18	The SI unit of enthalpy	<p>A. Calorie B. eV C. kJ mol^{-1} D. J mol^{-1}</p>
19	Which species is both a Bronsted acid and base.	<p>A. H_2O B. Na^+ C. OH^- D. Cl^-</p>
20	What cause the polarity in a molecule like CHCl_3	<p>A. $\text{Equal sharing of electrons}$ B. Non Polar Bond C. $\text{Uneven distribution of electrons creating positive and negative centers}$ D. $\text{Absence of hydrogen atoms}$</p>